

# 4-Hydroxybenzamide

<b>Other names:</b>	4-Hydroxyphenylacetamide Benzamide, 4-hydroxy- p-Hydroxybenzamide p-Hydroxyphenyl acetamide
<b>Inchi:</b>	InChI=1S/C7H7NO2/c8-7(10)5-1-3-6(9)4-2-5/h1-4,9H,(H2,8,10)
<b>InchiKey:</b>	QXSAKPUBHTZHKW-UHFFFAOYSA-N
<b>Formula:</b>	C7H7NO2
<b>SMILES:</b>	NC(=O)c1ccc(O)cc1
<b>Mol. weight [g/mol]:</b>	137.14
<b>CAS:</b>	619-57-8

## Physical Properties

Property code	Value	Unit	Source
gf	-96.62	kJ/mol	Joback Method
hf	-207.38	kJ/mol	Joback Method
hfus	25.20	kJ/mol	Thermodynamic and structural aspects of hydroxybenzamide molecular crystals study
hvap	63.85	kJ/mol	Joback Method
log10ws	-1.19		Crippen Method
logp	0.491		Crippen Method
mcvol	103.150	ml/mol	McGowan Method
pc	6009.25	kPa	Joback Method
tb	593.26	K	Joback Method
tc	841.84	K	Joback Method
tf	439.98	K	Joback Method
vc	0.321	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	242.37	J/mol×K	593.26	Joback Method
cpg	251.48	J/mol×K	634.69	Joback Method
cpg	259.81	J/mol×K	676.12	Joback Method

cpg	267.46	J/mol×K	717.55	Joback Method
cpg	274.53	J/mol×K	758.98	Joback Method
cpg	281.13	J/mol×K	800.41	Joback Method
cpg	287.34	J/mol×K	841.84	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Partial molar volumes of some drug and pro-drug substances in 1-octanol</b>	<a href="https://www.doi.org/10.1016/j.jct.2009.10.002">https://www.doi.org/10.1016/j.jct.2009.10.002</a>
<b>Thermodynamic and structural aspects of hydroxybenzamide molecular crystals</b>	<a href="https://www.doi.org/10.1016/j.tca.2012.10.013">https://www.doi.org/10.1016/j.tca.2012.10.013</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C619578&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C619578&amp;Units=SI</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mccvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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