

# 2,3-Butanedithiol

<b>Other names:</b>	butane-2,3-dithiol
<b>Inchi:</b>	InChI=1S/C4H10S2/c1-3(5)4(2)6/h3-6H,1-2H3
<b>InchiKey:</b>	TWWSEEHCVDRRRI-UHFFFAOYSA-N
<b>Formula:</b>	C4H10S2
<b>SMILES:</b>	CC(S)C(C)S
<b>Mol. weight [g/mol]:</b>	122.25
<b>CAS:</b>	4532-64-3

## Physical Properties

Property code	Value	Unit	Source
gf	36.70	kJ/mol	Joback Method
hf	-59.49	kJ/mol	Joback Method
hfus	7.15	kJ/mol	Joback Method
hvap	37.20	kJ/mol	Joback Method
log10ws	-1.86		Crippen Method
logp	1.623		Crippen Method
mcvol	99.920	ml/mol	McGowan Method
pc	4704.19	kPa	Joback Method
tb	415.76	K	Joback Method
tc	645.26	K	Joback Method
tf	177.76	K	Joback Method
vc	0.355	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	173.86	J/mol×K	415.76	Joback Method
cpg	183.94	J/mol×K	454.01	Joback Method
cpg	193.50	J/mol×K	492.26	Joback Method
cpg	202.53	J/mol×K	530.51	Joback Method
cpg	211.07	J/mol×K	568.76	Joback Method
cpg	219.12	J/mol×K	607.01	Joback Method
cpg	226.70	J/mol×K	645.26	Joback Method

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.30164e+01
Coeff. B	-3.39747e+03
Coeff. C	-6.63360e+01
Temperature range (K), min.	333.25
Temperature range (K), max.	507.28

## Sources

<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C4532643&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C4532643&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

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