2-Methyl-1-butanol

Other names: 1-butanol, 2-methyl-
Inchi: InChI=1S/C5H12O/c1-3-5(2)4-6/h5-6H,3-4H2,1-2H3
InchiKey: QPRQEDXDYOZYLA-UHFFFAOYSA-N
Formula: C5H12O
SMILES: CCC(C)CO
Mol. weight [g/mol]: 88.15
CAS: 137-32-6

Physical Properties

Property code	Value	Unit	Source
gf	-148.04	kJ/mol	Joback Method
hf	-304.04	kJ/mol	Joback Method
hfus	9.27	kJ/mol	Joback Method
hvap	43.02	kJ/mol	Joback Method
log10ws	-0.47		Aqueous Solubility Prediction Method
logp	1.025		Crippen Method
mcvol	87.180	ml/mol	McGowan Method
pc	3916.03	kPa	Joback Method
tb	402.45	K	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
tb	401.93	K	Vapor-Liquid Equilibrium of Binary Mixtures Containing Ethyl Acetate + 2-Methyl-1-propanol and Ethyl Acetate + 2-Methyl-1-butanol at 101.3 kPa
tb	401.93	K	Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures Containing 2-Methyl-1-propanol + 2-Methyl-1-butanol, 2-Methyl-1-propanol + 3-Methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol
tb	401.27	K	Isobaric Vapor Liquid Equilibria for the Binary and Ternary Systems of 2-Methyl-1-butanol, 2-Methyl-butanol Acetate, and Dimethylformamide (DMF) at 101.3 kPa

tc	570.76	K	Joback Method
tf	191.93	K	Joback Method
tt	118.00	K	Heat Capacities in the Solid and in the Liquid Phase of Isomeric Pentanols
vc	0.329	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	216.94	J/mol×K	570.76	Joback Method
cpg	186.73	J/mol×K	460.61	Joback Method
cpg	169.82	J/mol×K	405.54	Joback Method
cpg	194.72	J/mol×K	488.15	Joback Method
cpg	178.43	J/mol×K	433.08	Joback Method
cpg	209.82	J/mol×K	543.22	Joback Method
cpg	202.42	J/mol×K	515.68	Joback Method
dvisc	0.3005917	Pa×s	191.93	Joback Method
dvisc	0.0005174	Pa×s	369.94	Joback Method
dvisc	0.0010745	Pa×s	334.34	Joback Method
dvisc	0.0026559	Pa×s	298.74	Joback Method
dvisc	0.0379452	Pa×s	227.53	Joback Method
dvisc	0.0083860	Pa×s	263.13	Joback Method
dvisc	0.0002833	Pa×s	405.54	Joback Method
kvisc	0.0000061	m2/s	293.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K
kvisc	0.0000031	m2/s	318.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K

kvisc	0.0000035	m2/s	313.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000041	m2/s	308.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000047	m2/s	303.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000053	m2/s	298.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
kvisc	0.0000027	m2/s	323.15	Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K	
pvap	50.00	kPa	382.75	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System	

pvap	69.33	kPa	391.40	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	74.66	kPa	393.43	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	79.99	kPa	395.35	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	85.33	kPa	397.17	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	90.66	kPa	398.90	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	97.33	kPa	400.96	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	16.70	kPa	357.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	23.30	kPa	364.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	30.00	kPa	370.45	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	33.30	kPa	372.95	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	36.70	kPa	375.15	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	43.30	kPa	379.05	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System

pvap	63.99	kPa	389.23	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	56.60	kPa	385.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	63.30	kPa	388.95	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	66.60	kPa	390.35	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	70.00	kPa	391.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	76.60	kPa	394.15	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	83.30	kPa	396.65	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	90.00	kPa	398.85	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	96.60	kPa	401.05	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	101.30	kPa	402.45	Vapor-Liquid Equilibrium of the Ethanol + 2-Methyl-1-butanol System
pvap	23.30	kPa	364.85	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	30.00	kPa	370.45	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol

pvap	33.30	kPa	373.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	36.70	kPa	375.15	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	43.30	kPa	379.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	50.00	kPa	382.75	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	56.60	kPa	385.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	63.30	kPa	388.95	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	70.00	kPa	391.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	58.66	kPa	386.92	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	83.30	kPa	396.65	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	90.00	kPa	398.85	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	96.60	kPa	401.05	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol
pvap	53.33	kPa	384.41	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2

pvap	48.66	kPa	382.05	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	44.00	kPa	379.50	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	40.00	kPa	377.13	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	36.00	kPa	374.56	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	32.00	kPa	371.74	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	24.00	kPa	365.12	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	20.66	kPa	361.80	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	17.33	kPa	358.02	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	14.66	kPa	354.51	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	9.33	kPa	345.51	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	12.00	kPa	350.44	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2
pvap	76.60	kPa	394.15	Effect of addition of calcium chloride on vapor pressure of 2-methyl-1-butanol

pvap	28.00	kPa	368.62	Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2	
rfi	1.40860		298.15	Study of the Effects of Temperature and Pressure on the Thermodynamic and Acoustic Properties of 2-Methyl-1-butanol at Temperatures from 293K to 318K and Pressures up to 100MPa	
rfi	1.40872		298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Propanol + 2-Methyl-1-butanol and Propanol + 3-Methyl-1-butanol	
rfi	1.40872		298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Methanol + 2-Methyl-1-butanol and Ethanol + 2-Methyl-1-butanol	
rfi	1.40950		293.15	Isobaric Vapor - Liquid Equilibria for the Ternary System of 2-Methyl-1-butanol, 3-Methyl-1-butanol, and Ethylene Glycol at 101.3 kPa	

rfi	1.40840		298.15	Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa
rhol	823.02	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	815.85	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	812.12	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	808.29	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	804.41	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa

rhol	800.47	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	796.47	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	792.41	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	814.83	kg/m3	298.15	Vapor Liquid Equilibrium for 2-Methyl-1-butanol + Ethylbenzene + Xylene Isomers at 101.33 kPa
rhol	819.64	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	823.35	kg/m3	288.15	Densities and Viscosities for Binary Liquid Mixtures of Pentanol Isomers from (288.15 to 328.15) K at 0.1 MPa
rhol	783.84	kg/m3	338.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa

rhol	788.20	kg/m3	333.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	792.34	kg/m3	328.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	796.39	kg/m3	323.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	814.98	kg/m3	298.15	Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K
rhol	804.33	kg/m3	313.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	808.20	kg/m3	308.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa

rhol	812.03	kg/m3	303.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	815.82	kg/m3	298.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	819.56	kg/m3	293.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	800.39	kg/m3	318.15	Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa
rhol	807.63	kg/m3	308.15	Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K
srf	0.02	N/m	323.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	318.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K

srf	0.02	N/m	313.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	303.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	298.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	293.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K
srf	0.02	N/m	308.15	Surface Tensions of Three Amyl Alcohol + Ethanol Binary Mixtures from (293.15 to 323.15) K

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.43135e+01
Coeff. B	-2.78286e+03
Coeff. C	-1.14962e+02
Temperature range (K), min.	313.37
Temperature range (K), max.	424.10

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	2.88019e+02
Coeff. B	-1.78736e+04
Coeff. C	-4.06446e+01

Coeff. D	2.96397e-05
Temperature range (K), min.	203.00
Temperature range (K), max.	565.00

Sources

Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Alcohols at (94.0 or 95.6) kPa:	https://www.doi.org/10.1021/je030141r
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=828
Vapor Liquid Equilibrium for 2-Methyl-1-butanol + Ethylbenzene + Xylene Isomers at 101.33 kPa:	https://www.doi.org/10.1021/je301220s
Ethanol + 2-Methyl-1-butanol + Calcium Chloride System: Vapor-Liquid Equilibrium Data and Correlation Using the NRTL Electrolyte Model:	https://www.doi.org/10.1021/je060399s
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Temperature Dependence of Limiting Activity Coefficients, Henry s Law Constants, and Related Infinite Dilution Properties of Branched Pentanols in Water. Measurement, Critical Compilation, Correlation, and Recommended Data:	https://www.doi.org/10.1021/je901063s
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Density, Viscosity, Excess Molar Volume, and Viscosity Deviation of Three Amyl Alcohols + Ethanol Binary Mixtures from 293.15 to 323.15 K:	https://www.doi.org/10.1021/je050416y
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Mutual diffusion coefficients of 3-methyl-1-butanol + n-heptane and 2-methyl-1-butanol + n-heptane from 288.15 K to 318.15 K:	https://www.doi.org/10.1016/j.jct.2018.10.023
Liquid Liquid Equilibria for the Ternary Systems of Water + Butane-2,3-diol + 2-Methylbutan-1-ol, 2-Ethylbutan-1-ol, and 4-Methylpentan-2-ol at 298.15 and 308.15 K:	https://www.doi.org/10.1021/acs.jced.6b00945
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Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Propanol + 2-Methyl-1-butanol and Propanol + 3-Methyl-1-butanol:	https://www.doi.org/10.1021/je0502142
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The influence of cation, anion and temperature on the liquid-liquid equilibrium of some pentanols-water system:	https://www.doi.org/10.1016/j.fluid.2011.03.018
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=B6001118&Units=SI
Solubility of Anthracene in Binary Alcohol + Acetonitrile Solvent Mixtures at 298.2 K:	https://www.doi.org/10.1021/je0201727
Henry's law constants and infinite dilution activity coefficients of cis-2-butene, dimethylether, chloroethane, and 1,1-difluoroethane in methanol, 1-propanol, 2-propanol, 1-butanol, 2-butanol, isobutanol, tert-butanol, 1-pentanol, 2-pentanol, 3-pentanol, 2-methyl-1-butanol, 3-methyl-1-butanol, and 2-methyl-2-butanol:	https://www.doi.org/10.1016/j.jct.2005.08.004
Vapor pressure of selected aliphatic alcohols by ebulliometry. Part 2:	https://www.doi.org/10.1016/j.fluid.2010.06.020
Henry's law constants and infinite dilution activity coefficients of propane, propene, butane, isobutane, 1-butene, isobutene, trans-2-butene, and 1,3-butadiene in 2-methyl-1-butanol, 3-methyl-1-butanol, and 2-methyl-2-butanol:	https://www.doi.org/10.1021/je049723e
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=828
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Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Methanol + 2-Methyl-1-butanol and Ethanol + 2-Methyl-1-butanol:	https://www.doi.org/10.1021/je0500330
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Densities and Viscosities for Binary Liquid Mixtures of Biodiesel + 1-Pentanol, 2-Pentanol, or 2-Methyl-1-Butanol from (288.15 to 338.15) K at 0.1 MPa:	https://www.doi.org/10.1021/acs.jced.7b00996
Excess molar enthalpies of the binary systems: (Dibutyl ether + isomers of pentanol) at T = (298.15 and 308.15) K:	https://www.doi.org/10.1016/j.jct.2015.03.021

Isobaric Vapor - Liquid Equilibria for the Ternary System of 2-Methyl-1-butanol, 3-Methyl-1-butanol, and Diethylene Glycol at 101.3 kPa: https://www.doi.org/10.1021/je030216r

Isobaric Vapor-Liquid Equilibria for 2-Methyl-1-butanol + 3-Methyl-1-butanol + CuCl2, ZnCl2, and FeCl3 Systems at 101.3 kPa: https://www.doi.org/10.1021/je9008365

Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures Containing 2-Methyl-1-propanol + 2-Methyl-1-butanol, 2-Methyl-1-propanol + 3-Methyl-1-butanol, and 2-Methyl-1-propanol + 1-Pentanol: https://www.doi.org/10.1021/je060238o

Legend

cpg:	Ideal gas heat capacity
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hvap:	Enthalpy of vaporization at standard conditions
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume