

# 2-Ethylbutyl methyl phthalate

<b>Other names:</b>	1,2-Benzenedicarboxylic acid, 2-ethylbutyl methyl ester
<b>Inchi:</b>	InChI=1S/C15H20O4/c1-4-11(5-2)10-19-15(17)13-9-7-6-8-12(13)14(16)18-3/h6-9,11H,4-
<b>InchiKey:</b>	UFXXNCPNXIUUKHW-UHFFFAOYSA-N
<b>Formula:</b>	C15H20O4
<b>SMILES:</b>	CCC(CC)COC(=O)c1ccccc1C(=O)OC
<b>Mol. weight [g/mol]:</b>	264.32

## Physical Properties

Property code	Value	Unit	Source
gf	-292.08	kJ/mol	Joback Method
hf	-622.75	kJ/mol	Joback Method
hfus	30.31	kJ/mol	Joback Method
hvap	69.85	kJ/mol	Joback Method
log10ws	-3.81		Crippen Method
logp	3.066		Crippen Method
mvol	213.330	ml/mol	McGowan Method
pc	1989.43	kPa	Joback Method
rinpol	1885.00		NIST Webbook
tb	726.40	K	Joback Method
tc	933.26	K	Joback Method
tf	427.07	K	Joback Method
vc	0.809	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	594.17	J/molxK	726.40	Joback Method
cpg	659.69	J/molxK	898.78	Joback Method
cpg	648.50	J/molxK	864.30	Joback Method
cpg	636.36	J/molxK	829.83	Joback Method
cpg	623.26	J/molxK	795.35	Joback Method
cpg	609.21	J/molxK	760.88	Joback Method
cpg	669.96	J/molxK	933.26	Joback Method
dvisc	0.0000976	Paxs	726.40	Joback Method

dvisc	0.0001255	Paxs	676.51	Joback Method
dvisc	0.0001681	Paxs	626.62	Joback Method
dvisc	0.0002367	Paxs	576.74	Joback Method
dvisc	0.0003556	Paxs	526.85	Joback Method
dvisc	0.0005817	Paxs	476.96	Joback Method
dvisc	0.0010677	Paxs	427.07	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U373896&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U373896&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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