

# 1-Octen-4-ol

<b>Inchi:</b>	InChI=1S/C8H16O/c1-3-5-7-8(9)6-4-2/h4,8-9H,2-3,5-7H2,1H3
<b>InchiKey:</b>	UZGCMRVEDHLBGY-UHFFFAOYSA-N
<b>Formula:</b>	C8H16O
<b>SMILES:</b>	C=CCC(O)CCCC
<b>Mol. weight [g/mol]:</b>	128.21
<b>CAS:</b>	40575-42-6

## Physical Properties

Property code	Value	Unit	Source
gf	-34.94	kJ/mol	Joback Method
hf	-240.53	kJ/mol	Joback Method
hfus	15.76	kJ/mol	Joback Method
hvap	49.02	kJ/mol	Joback Method
log10ws	-2.40		Crippen Method
logp	2.114		Crippen Method
mcvol	125.150	ml/mol	McGowan Method
pc	2940.89	kPa	Joback Method
tb	470.86	K	Joback Method
tc	636.76	K	Joback Method
tf	223.98	K	Joback Method
vc	0.477	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	273.08	J/molxK	470.86	Joback Method
cpg	284.36	J/molxK	498.51	Joback Method
cpg	295.19	J/molxK	526.16	Joback Method
cpg	305.56	J/molxK	553.81	Joback Method
cpg	315.50	J/molxK	581.46	Joback Method
cpg	325.02	J/molxK	609.11	Joback Method
cpg	334.14	J/molxK	636.76	Joback Method
dvisc	0.1097275	Paxs	223.98	Joback Method
dvisc	0.0162547	Paxs	265.13	Joback Method

dvisc	0.0040223	Paxs	306.27	Joback Method
dvisc	0.0013856	Paxs	347.42	Joback Method
dvisc	0.0005982	Paxs	388.57	Joback Method
dvisc	0.0003033	Paxs	429.71	Joback Method
dvisc	0.0001732	Paxs	470.86	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.38593e+01
Coeff. B	-3.77935e+03
Coeff. C	-6.68480e+01
Temperature range (K), min.	345.32
Temperature range (K), max.	508.99

## Sources

<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C40575426&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C40575426&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l

<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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