3-Pentanone

Other names:	(C2H5)2CO
	1,3-Dimethylacetone
	DEK
	Diethyl ketone
	Diethylcetone
	Dimethylacetone
	Ethyl ketone
	Ethyl propionyl
	Metacetone
	Methacetone
	NSC 8653
	PROPIONE
	Pentan-3-one
	Pentanone-3
	UN 1156
Inchi:	InChI=1S/C5H10O/c1-3-5(6)4-2/h3-4H2,1-2H3
InchiKey:	FDPIMTJIUBPUKL-UHFFFAOYSA-N
Formula:	C5H10O
SMILES:	CCC(=O)CC
Mol. weight [g/mol]:	86.13
CAS:	96-22-0

Physical Properties

Property code	Value	Unit	Source
af	0.3440		KDB
affp	836.80	kJ/mol	NIST Webbook
basg	807.00	kJ/mol	NIST Webbook
basg	802.60	kJ/mol	NIST Webbook
basg	803.80 ± 0.30	kJ/mol	NIST Webbook
chl	-3104.70 ± 0.90	kJ/mol	NIST Webbook
chl	-3100.20 ± 1.00	kJ/mol	NIST Webbook
dm	2.70	debye	KDB
gf	-135.40	kJ/mol	KDB
hf	-253.40 ± 0.90	kJ/mol	NIST Webbook
hf	-257.95 ± 0.84	kJ/mol	NIST Webbook
hf	-260.50 ± 1.60	kJ/mol	NIST Webbook
hf	-258.80	kJ/mol	KDB

hfl	-296.51 ± 0.83	kJ/mol	NIST Webbook
hfus	10.30	kJ/mol	Joback Method
hvap	38.50	kJ/mol	NIST Webbook
hvap	38.52	kJ/mol	NIST Webbook
hvap	38.68	kJ/mol	NIST Webbook
hvap	38.60	kJ/mol	NIST Webbook
hvap	38.70 ± 0.30	kJ/mol	NIST Webbook
ie	9.31 ± 0.02	eV	NIST Webbook
ie	9.32 ± 0.01	eV	NIST Webbook
ie	9.31 ± 0.01	eV	NIST Webbook
ie	9.31 ± 0.01	eV	NIST Webbook
ie	9.30	eV	NIST Webbook
ie	9.31	eV	NIST Webbook
ie	9.31 ± 0.02	eV	NIST Webbook
ie	9.37 ± 0.03	eV	NIST Webbook
ie	9.22 ± 0.02	eV	NIST Webbook
log10ws	-0.28		Estimated Solubility Method
log10ws	-0.28		Aqueous Solubility Prediction Method
logp	1.375		Crippen Method
mcvol	82.880	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=1)		KDB
pc	3740.00 ± 41.40	kPa	NIST Webbook
pc	3729.00	kPa	KDB
pc	3729.00 ± 10.00	kPa	NIST Webbook
rhoc	255.81 ± 30.15	kg/m3	NIST Webbook
rinpol	675.00		NIST Webbook
rinpol	675.52		NIST Webbook
rinpol	677.74		NIST Webbook
rinpol	677.00		NIST Webbook
rinpol	675.42		NIST Webbook
rinpol	676.00		NIST Webbook
rinpol	676.00		NIST Webbook
rinpol	682.00		NIST Webbook
rinpol	681.00		NIST Webbook
rinpol	682.00		NIST Webbook
rinpol	682.00		NIST Webbook
rinpol	640.00		NIST Webbook
rinpol	658.00		NIST Webbook
rinpol	665.00		NIST Webbook
rinpol	651.00		NIST Webbook
rinpol	675.75		NIST Webbook
rinpol	647.00		NIST Webbook

rinpol	651.00	NIST Webbook
rinpol	647.00	NIST Webbook
rinpol	672.00	NIST Webbook
rinpol	700.00	NIST Webbook
rinpol	651.00	NIST Webbook
rinpol	651.00	NIST Webbook
rinpol	670.00	NIST Webbook
rinpol	659.00	NIST Webbook
rinpol	651.00	NIST Webbook
rinpol	703.00	NIST Webbook
rinpol	707.00	NIST Webbook
rinpol	694.00	NIST Webbook
rinpol	703.00	NIST Webbook
rinpol	675.00	NIST Webbook
rinpol	665.00	NIST Webbook
rinpol	672.00	NIST Webbook
rinpol	689.00	NIST Webbook
rinpol	697.00	NIST Webbook
rinpol	687.00	NIST Webbook
rinpol	697.00	NIST Webbook
rinpol	675.00	NIST Webbook
rinpol	700.00	NIST Webbook
rinpol	701.30	NIST Webbook
rinpol	671.00	NIST Webbook
rinpol	669.00	NIST Webbook
rinpol	669.00	NIST Webbook
rinpol	681.00	NIST Webbook
rinpol	669.92	NIST Webbook
rinpol	701.00	NIST Webbook
rinpol	681.00	NIST Webbook
rinpol	669.00	NIST Webbook
rinpol	688.00	NIST Webbook
rinpol	685.00	NIST Webbook
rinpol	651.00	NIST Webbook
rinpol	676.00	NIST Webbook
rinpol	650.00	NIST Webbook
rinpol	677.00	NIST Webbook
rinpol	705.00	NIST Webbook
rinpol	675.00	NIST Webbook
rinpol	676.91	NIST Webbook
rinpol	688.00	NIST Webbook
rinpol	675.00	NIST Webbook
rinpol	676.00	NIST Webbook
rinpol	678.00	NIST Webbook

rinpol	678.00	NIST Webbook
rinpol	683.00	NIST Webbook
rinpol	666.00	NIST Webbook
rinpol	669.00	NIST Webbook
rinpol	685.00	NIST Webbook
rinpol	683.00	NIST Webbook
rinpol	643.00	NIST Webbook
rinpol	676.40	NIST Webbook
rinpol	638.00	NIST Webbook
rinpol	675.00	NIST Webbook
rinpol	675.84	NIST Webbook
ripol	1011.30	NIST Webbook
ripol	976.00	NIST Webbook
ripol	958.00	NIST Webbook
ripol	983.00	NIST Webbook
ripol	983.00	NIST Webbook
ripol	983.00	NIST Webbook
ripol	978.00	NIST Webbook
ripol	977.00	NIST Webbook
ripol	970.00	NIST Webbook
ripol	956.00	NIST Webbook
ripol	979.00	NIST Webbook
ripol	979.00	NIST Webbook
ripol	970.00	NIST Webbook
ripol	981.00	NIST Webbook
ripol	975.00	NIST Webbook
ripol	964.00	NIST Webbook
ripol	980.00	NIST Webbook
ripol	971.00	NIST Webbook
ripol	969.00	NIST Webbook
ripol	1002.00	NIST Webbook
ripol	983.00	NIST Webbook
ripol	980.00	NIST Webbook
ripol	997.00	NIST Webbook
ripol	980.00	NIST Webbook
ripol	986.00	NIST Webbook
ripol	965.00	NIST Webbook
ripol	977.00	NIST Webbook
ripol	981.00	NIST Webbook
ripol	965.00	NIST Webbook
ripol	977.00	NIST Webbook
ripol	970.00	NIST Webbook
ripol	971.00	NIST Webbook
ripol	974.00	NIST Webbook

ripol	1001.50		NIST Webbook
ripol	996.90		NIST Webbook
ripol	1006.30		NIST Webbook
sl	266.00	J/mol×K	NIST Webbook
tb	375.11	K	KDB
tc	560.90 ± 0.56	K	NIST Webbook
tc	561.50	K	NIST Webbook
tc	561.46 ± 0.20	K	NIST Webbook
tc	561.46	K	KDB
tf	234.00	K	KDB
tf	231.15 ± 0.50	K	NIST Webbook
tf	231.20 ± 1.50	K	NIST Webbook
tf	233.35 ± 0.40	K	NIST Webbook
tf	234.18 ± 0.01	K	NIST Webbook
tf	234.15 ± 0.02	K	NIST Webbook
tf	233.90	K	Aqueous Solubility Prediction Method
tt	234.36	K	Thermodynamics of binary mixtures of N-methyl-2-pyrrolidinone and ketone. Experimental results and modelling of the solid-liquid equilibrium and vapou-liquid equilibrium. The Modified UNIFAC (Do) model characterization
tt	234.16 ± 0.03	K	NIST Webbook
tt	234.16 ± 0.03	K	NIST Webbook
vc	0.336	m3/kmol	KDB
zc	0.2683960		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	162.13	J/mol×K	423.15	NIST Webbook
cpg	156.77	J/mol×K	403.15	NIST Webbook
cpg	151.38	J/mol×K	383.15	NIST Webbook
cpg	146.31	J/mol×K	364.15	NIST Webbook
cpg	175.14	J/mol×K	473.15	NIST Webbook
cpg	168.70	J/mol×K	448.15	NIST Webbook
cpl	200.70	J/mol×K	298.15	NIST Webbook
cpl	195.70	J/mol×K	298.15	NIST Webbook
cpl	190.30	J/mol×K	298.15	NIST Webbook
cpl	190.00	J/mol×K	298.15	NIST Webbook

cpl	190.90	J/mol×K	298.15	NIST Webbook	
cpl	196.40	J/mol×K	298.15	NIST Webbook	
dvisc	0.0003970	Pa×s	308.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of Diethyl Oxalate with Some Ketones at (303.15, 308.15, and 313.15) K	
dvisc	0.0004290	Pa×s	303.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of Diethyl Oxalate with Some Ketones at (303.15, 308.15, and 313.15) K	
dvisc	0.0004714	Pa×s	293.15	Viscosities, Densities, and Ultrasonic Velocities of 3-Pentanone + Ethylbenzene and 3-Pentanone + o-Xylene at (293.15, 303.15, and 313.15) K	
dvisc	0.0003880	Pa×s	313.15	Densities, Viscosities, and Refractive Indices of Binary Mixtures of Diethyl Oxalate with Some Ketones at (303.15, 308.15, and 313.15) K	
dvisc	0.0003760	Pa×s	313.15	Density and Viscosity of Binary Mixtures of n-Butyl Acetate with Ketones at (298.15, 303.15, 308.15, and 313.15) K	
dvisc	0.0003960	Pa×s	308.15	Density and Viscosity of Binary Mixtures of n-Butyl Acetate with Ketones at (298.15, 303.15, 308.15, and 313.15) K	

dvisc	0.0004190	Pa×s	303.15	Density and Viscosity of Binary Mixtures of n-Butyl Acetate with Ketones at (298.15, 303.15, 308.15, and 313.15) K
dvisc	0.0004490	Pa×s	298.15	Density and Viscosity of Binary Mixtures of n-Butyl Acetate with Ketones at (298.15, 303.15, 308.15, and 313.15) K
dvisc	0.0003799	Pa×s	313.15	Viscosities, Densities, and Ultrasonic Velocities of 3-Pentanone + Ethylbenzene and 3-Pentanone + o-Xylene at (293.15, 303.15, and 313.15) K
dvisc	0.0004231	Pa×s	303.15	Viscosities, Densities, and Ultrasonic Velocities of 3-Pentanone + Ethylbenzene and 3-Pentanone + o-Xylene at (293.15, 303.15, and 313.15) K
hfust	11.59	kJ/mol	234.20	NIST Webbook
hfust	0.11	kJ/mol	118.50	NIST Webbook
hfust	0.01	kJ/mol	180.00	NIST Webbook
hfust	11.59	kJ/mol	234.20	NIST Webbook
hvapt	35.90 ± 0.20	kJ/mol	332.50	NIST Webbook
hvapt	36.60	kJ/mol	377.50	NIST Webbook
hvapt	33.70	kJ/mol	461.50	NIST Webbook
hvapt	33.30	kJ/mol	527.50	NIST Webbook
hvapt	36.10 ± 0.10	kJ/mol	335.00	NIST Webbook
hvapt	34.90 ± 0.10	kJ/mol	354.00	NIST Webbook
hvapt	33.50 ± 0.10	kJ/mol	375.00	NIST Webbook
hvapt	36.90	kJ/mol	303.00	NIST Webbook
hvapt	33.45	kJ/mol	375.20	NIST Webbook
hvapt	36.60	kJ/mol	356.50	NIST Webbook

rfi	1.39020		298.15	Excess molar volumes and ultrasonic studies of N-methyl-2-pyrrolidone with ketones at T = 303.15 K
rfi	1.39000		298.15	Phase Equilibria in the Binary and Ternary Systems Composed of Diethyl ketone, 2-Pentanone and 3-Pentanol at 101.3 kPa
rhol	809.38	kg/m3	298.15	Excess molar volumes, viscosity deviations and excess thermal expansion coefficients for binary and ternary mixtures consist of diethylketone + 2-butanol + ethylchloroacetate at (298.15, 308.15 and 318.15) K
rhol	809.16	kg/m3	298.15	Thermal and Volumetric Properties of Some C5 and C6 Alkanones at Infinite Dilution in Water
rhol	818.90	kg/m3	288.15	Thermal and Volumetric Properties of Some C5 and C6 Alkanones at Infinite Dilution in Water
rhol	804.62	kg/m3	303.15	Excess molar volumes and ultrasonic studies of dimethylsulphoxide with ketones at T = 303.15 K
rhol	799.32	kg/m3	308.15	Thermal and Volumetric Properties of Some C5 and C6 Alkanones at Infinite Dilution in Water

rhol	789.63	kg/m3	318.15	Excess molar volumes, viscosity deviations and excess thermal expansion coefficients for binary and ternary mixtures consist of diethylketone + 2-butanol + ethylchloroacetate at (298.15, 308.15 and 318.15) K
rhol	814.00	kg/m3	293.00	KDB
rhol	809.09	kg/m3	298.15	Excess molar enthalpies and volumes of binary mixtures of nonafluorobutylmethylether with ketones at T = 298.15 K
rhol	799.55	kg/m3	308.15	Excess molar volumes, viscosity deviations and excess thermal expansion coefficients for binary and ternary mixtures consist of diethylketone + 2-butanol + ethylchloroacetate at (298.15, 308.15 and 318.15) K
rhol	789.41	kg/m3	318.15	Thermal and Volumetric Properties of Some C5 and C6 Alkanones at Infinite Dilution in Water
sfust	0.96	J/mol×K	118.50	NIST Webbook
sfust	0.04	J/mol×K	180.00	NIST Webbook
sfust	49.50	J/mol×K	234.20	NIST Webbook
speedsl	1179.00	m/s	308.15	Densities and Speeds of Sound for Binary Liquid Mixtures of Thiolane-l,l-dioxide with Butanone, Pentan-2-one, Pentan-3-one, and 4-Methyl-pentan-2-one at T = (303.15 or 308.15 or 313.15) K

speedsl	1160.00	m/s	313.15	Densities and Speeds of Sound for Binary Liquid Mixtures of Thiolane-l,l-dioxide with Butanone, Pentan-2-one, Pentan-3-one, and 4-Methyl-pentan-2-one at T = (303.15 or 308.15 or 313.15) K
speedsl	1197.00	m/s	303.15	Densities and Speeds of Sound for Binary Liquid Mixtures of Thiolane-l,l-dioxide with Butanone, Pentan-2-one, Pentan-3-one, and 4-Methyl-pentan-2-one at T = (303.15 or 308.15 or 313.15) K
srf	0.02	N/m	319.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	329.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	315.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	317.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K

srf	0.02	N/m	321.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	313.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	311.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	309.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	307.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	305.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	303.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K

srf	0.02	N/m	323.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	325.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	283.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.02	N/m	301.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	299.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	297.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	295.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K

srf	0.02	N/m	327.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	291.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	289.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	287.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	285.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K
srf	0.03	N/m	293.15	Thermodynamic investigation of methyl salicylate/1-pentanol binary system in the temperature range from 278.15 K to 303.15 K

Correlations

Information	Value

Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.47173e+01
Coeff. B	-3.36681e+03
Coeff. C	-4.16190e+01
Temperature range (K), min.	274.94
Temperature range (K), max.	399.57

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	7.00329e+01
Coeff. B	-6.68960e+03
Coeff. C	-8.16282e+00
Coeff. D	5.65048e-06
Temperature range (K), min.	234.18
Temperature range (K), max.	561.00

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
308.15	101.30	0.0003710

Reference *https://www.doi.org/10.1021/je8003723*

Sources

Measurement of vapor.liquid equilibria (VLE) and excess enthalpies (HE) of binary systems with https://www.doi.org/10.1016/j.fluid.2007.06.001

Activity coefficients at infinite dilution and physicochemical properties for https://www.doi.org/10.1016/j.jct.2015.02.024

Excess molar enthalpies and volumes of binary mixtures of https://www.doi.org/10.1016/j.jct.2006.03.003

Prediction of these properties and A https://www.doi.org/10.1016/j.fluid.2014.11.020

https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure

Activity coefficients at infinite dilution of organic solvents and water in 1-butyl-3-methylimidazolium dicyanamide. A literature review of hexane/hex-1-ene separation: https://www.doi.org/10.1016/j.fluid.2016.02.004

Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1,4-diaza[2.2.2]bicyclooctanium bis(trifluoromethylsulfonyl)imide using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.026
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Henry's Law Constants of Organic Compounds in Water and n-Octane at T = 293.2 K:	https://www.doi.org/10.1021/je900711h
Density and Viscosity of Binary Mixtures of n-Butyl Acetate with Ketones at (298.15, 303.15, 308.15, and 313.15) K:	https://www.doi.org/10.1021/je300085z
High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid:	https://www.doi.org/10.1016/j.fluid.2017.06.001
Estimated Solubility Method:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Thermodynamics of binary mixtures of N-methyl-2-pyrrolidinone and ketone. Experimental results and modelling of the solid–liquid equilibrium and vapour–liquid equilibrium. The Modified UNIFAC (Dortmund) model characterization:	https://www.doi.org/10.1016/j.jct.2004.11.007
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in [illegible]phosphonium [illegible] Using Gas–Liquid Chromatography: Thermodynamic Properties of Mixtures Containing Ionic Liquids:	https://www.doi.org/10.1021/acs.jced.5b00980
Thermodynamics and activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.06.007
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Separation of binary mixtures based on limiting activity coefficients data using specific ammonium-based ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.fluid.2017.12.029
Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je800658v
Phase Equilibria of (1-Ethyl-3-methylimidazolium Ethylsulfate + Hydrocarbon, + Ketone, and + Ether) Binary Systems:	https://www.doi.org/10.1021/je700591h
Activity coefficients at infinite dilution of organic solutes in methylphosphonate based ionic liquids using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2015.02.023
Excess molar volumes and ultrasonic studies of dimethylsulphoxide with ketones at T = 303.15 K:	https://www.doi.org/10.1016/j.jct.2007.08.006
Infinite Dilution Activity Coefficients of Solutes Dissolved in Two Trihexyl(tetradecyl)phosphonium Ionic Liquids:	https://www.doi.org/10.1021/je500050p
Excess molar volumes, viscosity deviations and excess thermal expansion coefficients for binary and ternary mixtures consist of diethylketone + 2-butanol + ethylchloroacetate at (298.15, 308.15 and 318.15) K:	https://www.doi.org/10.1016/j.tca.2013.03.008
Excess molar volumes and ultrasonic studies of N-methyl-2-pyrrolidone with ketones at T = 303.15 K:	https://www.doi.org/10.1016/j.jct.2009.01.006
Partition Coefficients of Organic Compounds in Four New Tetraalkylammonium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je200637v
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.11.021
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Bearing Short Linear Alkyl Side Chains of Three to Five Carbons:	https://www.doi.org/10.1021/je4001894
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid [illegible] tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.02.012
Physicochemical properties and activity coefficients at infinite dilution for organic solutes and water in a novel bicyclic guanidinium superbase-derived protic ionic liquid:	https://www.doi.org/10.1016/j.jct.2012.09.033
(Liquid + liquid) phase equilibria of 1-alkyl-3-methylimidazolium thiocyanate with alcohols, or ethers, or ketones:	https://www.doi.org/10.1016/j.jct.2005.07.024
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Legend

af:	Acentric Factor
affp:	Proton affinity
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility