

Carbon dioxide

Other names:	Anhydride carbonique CARBONIC ANHYDRIDE CO ₂ Carbon oxide (CO ₂) Carbonic acid anhydride Carbonic acid, gas Carbonica Cardice Dricold Drikold Dry ice Khladon 744 Kohlendioxyd Kohlensäure R 744 UN 1013 UN 1845 UN 2187
Inchi:	InChI=1S/CO2/c2-1-3
InchiKey:	CURLTUGMZLYLDI-UHFFFAOYSA-N
Formula:	CO ₂
SMILES:	O=C=O
Mol. weight [g/mol]:	44.01
CAS:	124-38-9

Physical Properties

Property code	Value	Unit	Source
af	0.2390		KDB
affp	540.50	kJ/mol	NIST Webbook
basg	515.80	kJ/mol	NIST Webbook
dm	0.00	debye	KDB
ea	-0.60	eV	NIST Webbook
ea	-1.60 ± 0.10	eV	NIST Webbook
gf	-394.60	kJ/mol	KDB
gyrad	0.9920		KDB
hf	-393.80	kJ/mol	KDB
hf	-393.51 ± 0.13	kJ/mol	NIST Webbook

hfus	11.09	kJ/mol	Joback Method
hvap	29.78	kJ/mol	Joback Method
ie	13.78 ± 0.00	eV	NIST Webbook
ie	13.78 ± 0.00	eV	NIST Webbook
ie	13.77 ± 0.00	eV	NIST Webbook
ie	13.90 ± 0.20	eV	NIST Webbook
ie	13.78 ± 0.01	eV	NIST Webbook
ie	13.78	eV	NIST Webbook
ie	13.79	eV	NIST Webbook
ie	13.77 ± 0.00	eV	NIST Webbook
ie	13.80 ± 0.01	eV	NIST Webbook
ie	13.79 ± 0.01	eV	NIST Webbook
ie	13.78 ± 0.01	eV	NIST Webbook
ie	13.83 ± 0.05	eV	NIST Webbook
ie	13.79	eV	NIST Webbook
ie	13.78	eV	NIST Webbook
ie	13.78	eV	NIST Webbook
ie	13.77	eV	NIST Webbook
ie	13.77 ± 0.03	eV	NIST Webbook
ie	13.77 ± 0.00	eV	NIST Webbook
ie	13.78 ± 0.00	eV	NIST Webbook
ie	13.78 ± 0.00	eV	NIST Webbook
ie	13.78	eV	NIST Webbook
ie	13.78 ± 0.00	eV	NIST Webbook
ie	13.00 ± 1.00	eV	NIST Webbook
ie	13.89 ± 0.03	eV	NIST Webbook
ie	13.79 ± 0.05	eV	NIST Webbook
ie	13.77	eV	NIST Webbook
ie	13.75 ± 0.05	eV	NIST Webbook
ie	13.77	eV	NIST Webbook
log10ws	0.79		Crippen Method
logp	-0.584		Crippen Method
mcvol	28.090	ml/mol	McGowan Method
pc	7380.00	kPa	Critical Properties of the Reacting Mixture in the Selective Oxidation of Cyclohexane by Oxygen in the Presence of Carbon Dioxide
pc	7340.00 ± 50.00	kPa	NIST Webbook
pc	7420.00	kPa	Critical Properties of the Reacting Mixture in the Esterification of Acetic Acid with Ethanol

pc	7360.00	kPa	Phase Behaviors, Density, and Isothermal Compressibility of Styrene + CO ₂ , Ethylbenzene + CO ₂ , and Ethylbenzene + Styrene + CO ₂ Systems
pc	7420.00	kPa	Critical Properties of Binary and Ternary Mixtures of Hexane + Methanol, Hexane + Carbon Dioxide, Methanol + Carbon Dioxide and Hexane + Carbon Dioxide + Methanol
pc	7375.00	kPa	KDB
pc	7380.00 ± 15.00	kPa	NIST Webbook
pc	7380.00	kPa	Critical Temperatures and Pressures of Several Binary and Ternary Mixtures Concerning the Alkylation of 2-Methylpropane with 1-Butene in the Presence of Methane or Carbon Dioxide
pc	7382.50 ± 0.50	kPa	NIST Webbook
pt	518.50 ± 0.50	kPa	NIST Webbook
pt	518.50	kPa	KDB
rhoc	475.30 ± 8.80	kg/m ³	NIST Webbook
rhoc	466.06 ± 0.31	kg/m ³	NIST Webbook
rhoc	466.50 ± 2.20	kg/m ³	NIST Webbook
rhoc	468.26 ± 1.50	kg/m ³	NIST Webbook
rinpol	152.00		NIST Webbook
rinpol	152.00		NIST Webbook
rinpol	153.00		NIST Webbook
rinpol	154.00		NIST Webbook
sgb	213.78 ± 0.01	J/mol×K	NIST Webbook
tb	194.70	K	KDB
tc	304.23	K	NIST Webbook
tc	304.14	K	KDB
tc	304.30	K	Validation of a New Apparatus Using the Dynamic Method for Determining the Critical Properties of Binary Gas/Gas Mixtures
tc	304.30	K	Development of a Predictive Equation of State for CO ₂ + Ethyl Ester Mixtures Based on Critical Points Measurements
tc	304.18 ± 0.04	K	NIST Webbook
tc	304.10 ± 0.10	K	NIST Webbook

tc	304.35 ± 0.40	K	NIST Webbook
tc	304.20 ± 0.02	K	NIST Webbook
tf	216.58	K	KDB
tt	216.52	K	Solid Liquid Equilibria for the CO ₂ + R23 and N ₂ O + R23 Systems
tt	216.60	K	Solid-liquid equilibria measurements of the carbondioxide + 2,3,3,3-tetrafluoroprop-1-ene and carbondioxide + trans-1,3,3,3-tetrafluoropropene mixtures
tt	216.58 ± 0.01	K	NIST Webbook
tt	216.58 ± 0.03	K	NIST Webbook
tt	216.58	K	KDB
vc	0.092 ± 0.001	m ³ /kmol	NIST Webbook
vc	0.094	m ³ /kmol	KDB
zc	0.2741450		KDB
zra	0.27		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	39.81	J/mol×K	288.60	Joback Method
cpg	41.25	J/mol×K	309.97	Joback Method
cpg	32.51	J/mol×K	203.15	Joback Method
cpg	34.58	J/mol×K	224.51	Joback Method
cpg	36.48	J/mol×K	245.88	Joback Method
cpg	38.22	J/mol×K	267.24	Joback Method
cpg	42.55	J/mol×K	331.33	Joback Method
dvisc	0.0000356	Paxs	800.00	Viscosity of H ₂ -CO ₂ Mixtures at (500, 800, and 1100) K
dvisc	0.0000452	Paxs	1100.00	Viscosity of H ₂ -CO ₂ Mixtures at (500, 800, and 1100) K
dvisc	0.0000236	Paxs	500.00	Viscosity of H ₂ -CO ₂ Mixtures at (500, 800, and 1100) K
hsubt	26.30	kJ/mol	167.00	NIST Webbook
hsubt	25.20	kJ/mol	175.00	NIST Webbook
hsubt	26.10	kJ/mol	207.00	NIST Webbook
hsubt	27.20 ± 0.40	kJ/mol	86.00	NIST Webbook

hsubt	25.90	kJ/mol	188.50	NIST Webbook
hvapt	16.40	kJ/mol	244.50	NIST Webbook
hvapt	16.50	kJ/mol	285.00	NIST Webbook
hvapt	15.33	kJ/mol	216.60	KDB
hvapt	28.84	kJ/mol	82.30	Measurements of enthalpy of sublimation of Ne, N ₂ , O ₂ , Ar, CO ₂ , Kr, Xe, and H ₂ O using a double paddle oscillator
hvapt	16.70	kJ/mol	288.50	NIST Webbook
pvap	682.90	kPa	223.14	Vapor-liquid equilibrium data for the carbon dioxide and nitrogen (CO ₂ +N ₂) system at the temperatures 223, 270, 298 and 303 K and pressures up to 18 MPa
pvap	5627.00	kPa	293.00	Experimental Measurements and Thermodynamic Modeling of the Dissociation Conditions of Clathrate Hydrates for (Refrigerant + NaCl + Water) Systems
pvap	1964.80	kPa	253.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K
pvap	2641.80	kPa	263.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K

pvap	3478.50	kPa	273.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K
pvap	4497.80	kPa	283.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K
pvap	5723.70	kPa	293.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K
pvap	7206.20	kPa	303.15	Vapor liquid equilibria for the carbon dioxide + propane system over a temperature range from 253.15 to 323.15K
pvap	3387.00	kPa	273.00	Experimental Measurements and Thermodynamic Modeling of the Dissociation Conditions of Clathrate Hydrates for (Refrigerant + NaCl + Water) Systems
pvap	6436.90	kPa	298.17	Vapor-liquid equilibrium data for the carbon dioxide and nitrogen (CO ₂ +N ₂) system at the temperatures 223, 270, 298 and 303 K and pressures up to 18 MPa

pvap	7210.50	kPa	303.16	Vapor-liquid equilibrium data for the carbon dioxide and nitrogen (CO ₂ +N ₂) system at the temperatures 223, 270, 298 and 303 K and pressures up to 18 MPa
pvap	1009.00	kPa	233.19	Phase equilibrium data for binary mixtures of carbon dioxide with {1,1,2,3,3,3-hexafluoro-1-propene or 2,2,3-trifluoro-3-(trifluoromethyl)oxirane} at temperatures between (233 and 273) K
pvap	1432.00	kPa	243.22	Phase equilibrium data for binary mixtures of carbon dioxide with {1,1,2,3,3,3-hexafluoro-1-propene or 2,2,3-trifluoro-3-(trifluoromethyl)oxirane} at temperatures between (233 and 273) K
pvap	1970.00	kPa	253.21	Phase equilibrium data for binary mixtures of carbon dioxide with {1,1,2,3,3,3-hexafluoro-1-propene or 2,2,3-trifluoro-3-(trifluoromethyl)oxirane} at temperatures between (233 and 273) K
pvap	2645.00	kPa	263.21	Phase equilibrium data for binary mixtures of carbon dioxide with {1,1,2,3,3,3-hexafluoro-1-propene or 2,2,3-trifluoro-3-(trifluoromethyl)oxirane} at temperatures between (233 and 273) K

pvap	3473.00	kPa	273.19	Phase equilibrium data for binary mixtures of carbon dioxide with {1,1,2,3,3,3-hexafluoro-1-propene or 2,2,3-trifluoro-3-(trifluoromethyl)oxirane} at temperatures between (233 and 273) K
pvap	3487.00	kPa	273.15	Phase equilibrium and critical point data for ethylene and chlorodifluoromethane binary mixtures using a new static-analytic apparatus
pvap	410.40	kPa	213.22	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa
pvap	681.80	kPa	223.15	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa

pvap	1426.30	kPa	243.13	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa
pvap	2646.70	kPa	263.13	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa
pvap	3491.30	kPa	273.22	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa

pvap	4502.30	kPa	283.15	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa
pvap	6595.00	kPa	299.22	Thermodynamics of the carbon dioxide plus argon (CO ₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa
pvap	2294.00	kPa	258.44	(Vapor + liquid) equilibrium data for (carbon dioxide + 1,1-difluoroethane) system at temperatures from (258 to 343) K and pressures up to about 8 MPa
pvap	3977.00	kPa	278.25	(Vapor + liquid) equilibrium data for (carbon dioxide + 1,1-difluoroethane) system at temperatures from (258 to 343) K and pressures up to about 8 MPa

pvap	6502.60	kPa	298.84	(Vapor + liquid) equilibrium data for (carbon dioxide + 1,1-difluoroethane) system at temperatures from (258 to 343) K and pressures up to about 8 MPa
pvap	6020.00	kPa	295.45	A synthetic-dynamic method for water solubility measurements in high pressure CO ₂ using ATR FTIR spectroscopy
pvap	6240.00	kPa	297.05	A synthetic-dynamic method for water solubility measurements in high pressure CO ₂ using ATR FTIR spectroscopy
pvap	6520.00	kPa	299.05	A synthetic-dynamic method for water solubility measurements in high pressure CO ₂ using ATR FTIR spectroscopy
pvap	7090.00	kPa	302.65	A synthetic-dynamic method for water solubility measurements in high pressure CO ₂ using ATR FTIR spectroscopy
pvap	7180.00	kPa	303.25	A synthetic-dynamic method for water solubility measurements in high pressure CO ₂ using ATR FTIR spectroscopy
pvap	3476.00	kPa	272.78	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)

pvap	3958.00	kPa	277.83	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	4487.00	kPa	282.87	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	5069.00	kPa	287.92	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	5705.00	kPa	292.95	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	6404.00	kPa	297.99	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	7176.00	kPa	303.03	Isothermal (vapour + liquid) equilibrium data for binary systems of (n-hexane + CO ₂ or CHF ₃)
pvap	1426.50	kPa	243.15	Experimental research on (vapor + liquid) equilibria for the {trifluoriodomethane (CF ₃ I) + carbon dioxide (CO ₂)} system from 243.150 to 273.150 K
pvap	1968.40	kPa	253.15	Experimental research on (vapor + liquid) equilibria for the {trifluoriodomethane (CF ₃ I) + carbon dioxide (CO ₂)} system from 243.150 to 273.150 K

pvap	2647.50	kPa	263.15	Experimental research on (vapor + liquid) equilibria for the {trifluoriodomethane (CF3I) + carbon dioxide (CO2)} system from 243.150 to 273.150 K
pvap	3483.80	kPa	273.15	Experimental research on (vapor + liquid) equilibria for the {trifluoriodomethane (CF3I) + carbon dioxide (CO2)} system from 243.150 to 273.150 K
pvap	552.50	kPa	218.15	Measurement and modelling of the vapor-liquid equilibrium of (CO2 + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	1004.40	kPa	233.15	Measurement and modelling of the vapor-liquid equilibrium of (CO2 + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	1428.00	kPa	243.15	Measurement and modelling of the vapor-liquid equilibrium of (CO2 + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	2293.70	kPa	258.15	Measurement and modelling of the vapor-liquid equilibrium of (CO2 + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa

pvap	3484.60	kPa	273.15	Measurement and modelling of the vapor-liquid equilibrium of (CO ₂ + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	5089.60	kPa	288.19	Measurement and modelling of the vapor-liquid equilibrium of (CO ₂ + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	7171.20	kPa	302.94	Measurement and modelling of the vapor-liquid equilibrium of (CO ₂ + CO) at temperatures between (218.15 and 302.93) K at pressures up to 15 MPa
pvap	3485.70	kPa	273.16	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	3975.00	kPa	278.20	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	4502.00	kPa	283.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	5087.00	kPa	288.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination

pvap	5726.40	kPa	293.14	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	6430.10	kPa	298.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7212.80	kPa	303.19	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7290.90	kPa	303.64	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7326.10	kPa	303.85	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7358.80	kPa	304.04	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	3485.90	kPa	273.17	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination

pvap	3977.00	kPa	278.20	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	4500.50	kPa	283.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	5090.50	kPa	288.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	5726.50	kPa	293.14	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	6436.10	kPa	298.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7214.40	kPa	303.15	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7297.40	kPa	303.65	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination

pvap	7329.00	kPa	303.85	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	7360.50	kPa	304.04	A novel technique based in a cylindrical microwave resonator for high pressure phase equilibrium determination
pvap	3962.00	kPa	278.15	Experimental Solubility Data for Binary Mixtures of Ethane and 2,2,4-Trimethylpentane at Pressures up to 6 MPa Using a New Variable-Volume Sapphire Cell
pvap	4392.00	kPa	282.15	Phase Equilibrium and Liquid Viscosity of CO ₂ + n-Dodecane Mixtures between 283 and 353 K
pvap	4961.00	kPa	287.11	Phase Equilibrium and Liquid Viscosity of CO ₂ + n-Dodecane Mixtures between 283 and 353 K
pvap	5555.00	kPa	291.84	Phase Equilibrium and Liquid Viscosity of CO ₂ + n-Dodecane Mixtures between 283 and 353 K
pvap	6214.00	kPa	296.64	Phase Equilibrium and Liquid Viscosity of CO ₂ + n-Dodecane Mixtures between 283 and 353 K
pvap	6928.00	kPa	301.38	Phase Equilibrium and Liquid Viscosity of CO ₂ + n-Dodecane Mixtures between 283 and 353 K

pvap	1000.72	kPa	233.07	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	1000.81	kPa	233.07	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	1420.40	kPa	243.01	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	1420.51	kPa	243.01	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	1974.16	kPa	253.22	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	1974.47	kPa	253.22	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	2661.57	kPa	263.31	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	2661.88	kPa	263.32	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	3522.07	kPa	273.55	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus

pvap	3522.39	kPa	273.56	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	4543.19	kPa	283.51	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	4543.58	kPa	283.52	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	5720.90	kPa	293.08	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	5721.42	kPa	293.09	Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Fluids with a New Apparatus
pvap	3960.00	kPa	278.15	Isothermal Vapor-Liquid Equilibria for the Binary System of Carbon Dioxide (CO ₂) + 1,1,1,2,3,3,3-Heptafluoropropane (R-227ea)
pvap	5080.00	kPa	288.15	Isothermal Vapor-Liquid Equilibria for the Binary System of Carbon Dioxide (CO ₂) + 1,1,1,2,3,3,3-Heptafluoropropane (R-227ea)
pvap	6425.00	kPa	298.15	Isothermal Vapor-Liquid Equilibria for the Binary System of Carbon Dioxide (CO ₂) + 1,1,1,2,3,3,3-Heptafluoropropane (R-227ea)

pvap	1672.60	kPa	248.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	3485.10	kPa	273.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	4406.00	kPa	283.00	Experimental Measurements and Thermodynamic Modeling of the Dissociation Conditions of Clathrate Hydrates for (Refrigerant + NaCl + Water) Systems
pvap	6323.00	kPa	298.10	Experimental Measurements and Thermodynamic Modeling of the Dissociation Conditions of Clathrate Hydrates for (Refrigerant + NaCl + Water) Systems
rhog	1.56	kg/m3	344.15	High-pressure densities and interfacial tensions of binary systems containing carbon dioxide + n-alkanes: (n-Dodecane, n-tridecane, n-tetradecane)
srf	0.00	N/m	273.20	KDB

Correlations

Information	Value
Property code	pvap

Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	2.05818e+01
Coeff. B	-3.08771e+03
Coeff. C	-1.26700e+00
Temperature range (K), min.	153.41
Temperature range (K), max.	264.19

Datasets

Kinematic viscosity, m²/s

Temperature, K - Gas	Pressure, kPa - Gas	Kinematic viscosity, m ² /s - Gas
220.00	195.46	0.0000024
220.00	226.71	0.0000020
220.00	254.44	0.0000018
220.00	288.21	0.0000016
220.00	324.47	0.0000014
220.00	362.71	0.0000012
220.00	393.15	0.0000011
220.00	425.51	0.0000010
220.00	449.74	0.0000010
220.00	479.78	0.0000009
235.00	224.78	0.0000023
235.00	359.42	0.0000014
235.00	452.88	0.0000011
235.00	545.18	0.0000009
235.00	638.46	0.0000008
235.00	723.88	0.0000007
235.00	803.31	0.0000006
250.00	274.20	0.0000022
250.00	372.33	0.0000016
250.00	491.25	0.0000012
250.00	595.84	0.0000010
250.00	702.28	0.0000008
250.00	818.89	0.0000007
250.00	953.03	0.0000006
250.00	1060.25	0.0000005
250.00	1177.57	0.0000004
250.00	1291.60	0.0000004

250.00	1438.24	0.0000004
270.00	196.61	0.0000035
270.00	354.95	0.0000020
270.00	526.01	0.0000013
270.00	705.02	0.0000009
270.00	898.25	0.0000007
270.00	1076.54	0.0000006
270.00	1287.09	0.0000005
270.00	1446.70	0.0000004
270.00	1630.02	0.0000004
270.00	1846.01	0.0000003
270.00	2079.64	0.0000003
270.00	2316.11	0.0000002
270.00	2509.61	0.0000002
285.00	371.48	0.0000021
285.00	575.25	0.0000013
285.00	787.74	0.0000009
285.00	1012.98	0.0000007
285.00	1228.51	0.0000006
285.00	1500.40	0.0000005
285.00	1707.23	0.0000004
285.00	1947.24	0.0000004
285.00	2221.25	0.0000003
285.00	2559.66	0.0000003
285.00	2841.89	0.0000002
285.00	3057.76	0.0000002
300.00	313.43	0.0000027
300.00	531.89	0.0000016
300.00	748.54	0.0000011
300.00	993.26	0.0000008
300.00	1237.01	0.0000007
300.00	1526.14	0.0000005
300.00	1747.76	0.0000005
300.00	1991.46	0.0000004
300.00	2272.21	0.0000003
300.00	2577.86	0.0000003
300.00	2893.94	0.0000003
300.00	3145.79	0.0000002
304.77	317.81	0.0000028
304.77	527.05	0.0000016
304.77	745.25	0.0000012
304.77	979.77	0.0000009
304.77	1201.73	0.0000007
304.77	1478.86	0.0000006

304.77	1699.38	0.0000005
304.77	1949.98	0.0000004
304.77	2235.18	0.0000004
304.77	2552.72	0.0000003
304.77	2881.02	0.0000003
304.77	3135.88	0.0000002
310.00	342.08	0.0000027
310.00	552.05	0.0000016
310.00	773.76	0.0000011
310.00	1018.04	0.0000009
310.00	1259.27	0.0000007
310.00	1549.04	0.0000006
310.00	1786.58	0.0000005
310.00	2055.58	0.0000004
310.00	2362.96	0.0000003
310.00	2708.78	0.0000003
310.00	3077.99	0.0000003
330.00	201.27	0.0000052
330.00	402.42	0.0000026
330.00	626.86	0.0000016
330.00	846.53	0.0000012
330.00	1059.35	0.0000009
330.00	1332.98	0.0000007
330.00	1559.26	0.0000006
330.00	1813.09	0.0000005
330.00	2103.20	0.0000005
330.00	2423.40	0.0000004
330.00	2778.74	0.0000003
330.00	3067.62	0.0000003
350.00	234.16	0.0000050
350.00	440.84	0.0000026
350.00	660.10	0.0000017
350.00	905.02	0.0000013
350.00	1143.54	0.0000010
350.00	1347.77	0.0000008
350.00	1584.41	0.0000007
350.00	1872.95	0.0000006
350.00	2228.51	0.0000005
350.00	2447.32	0.0000004
350.00	2827.17	0.0000004
350.00	3133.84	0.0000003
370.00	335.49	0.0000038
370.00	537.70	0.0000024
370.00	749.31	0.0000017

370.00	964.83	0.0000013
370.00	1228.54	0.0000010
370.00	1449.30	0.0000009
370.00	1714.08	0.0000007
370.00	2019.04	0.0000006
370.00	2382.70	0.0000005
370.00	2771.65	0.0000004
370.00	3095.89	0.0000004

Reference

<https://www.doi.org/10.1016/j.jct.2007.07.002>

Mass density, kg/m3

Temperature, K - Gas	Pressure, kPa - Gas	Mass density, kg/m3 - Gas
273.15	1001.00	21.2
273.15	2001.00	45.7
273.15	3001.00	77.3
273.15	3300.00	89.9
283.15	1000.00	20.4
283.15	2001.00	43.1
283.15	3001.00	70.9
283.15	4000.00	108.4
283.15	4400.00	130.1
293.15	1011.00	19.8
293.15	2001.00	41.0
293.15	3001.00	66.2
293.15	4000.00	97.3
293.15	5000.00	140.4
293.15	5600.00	182.2

Reference

<https://www.doi.org/10.1021/je300590v>

Mass density, kg/m3

Pressure, kPa - Fluid (supercritical or subcritical phases)	Temperature, K - Fluid (supercritical or subcritical phases)	Mass density, kg/m3 - Fluid (supercritical or subcritical phases)
6100.00	291.30	817.0
7100.00	291.50	834.0

8100.00	291.70	849.0
9200.00	291.90	861.0
10200.00	292.00	872.0
15200.00	292.10	916.0
20200.00	292.10	948.0
25300.00	292.10	975.0
30100.00	292.10	997.0

Reference

<https://www.doi.org/10.1016/j.fluid.2010.07.009>

Temperature, K	Pressure, kPa	Mass density, kg/m3
323.20	10968.00	497.981
323.25	12967.00	633.58
323.24	14966.00	698.2
323.25	16965.00	739.597
323.24	18962.00	770.445
323.26	20966.00	795.018
323.25	22962.00	815.681
323.24	24961.00	833.514
323.25	26959.00	849.181
323.25	28037.00	856.967

Reference

<https://www.doi.org/10.1016/j.jct.2015.07.014>

Temperature, K	Pressure, kPa	Mass density, kg/m3
308.00	10000.00	714.84
308.00	12000.00	768.42
308.00	15000.00	816.06
308.00	18000.00	848.87
308.00	21000.00	874.4
318.00	10000.00	502.57
318.00	12000.00	659.73
318.00	15000.00	743.17
318.00	18000.00	790.18
318.00	21000.00	823.71
328.00	10000.00	326.4
328.00	12000.00	506.85
328.00	15000.00	654.94
328.00	18000.00	724.13
328.00	21000.00	768.74

Reference

<https://www.doi.org/10.1016/j.tca.2011.01.039>

Temperature, K	Pressure, kPa	Mass density, kg/m3
304.29	10390.00	768.95
398.60	20324.00	391.9
469.77	20740.00	272.48
304.29	21349.00	895.09
304.29	31385.00	950.66
469.78	36325.00	466.7
398.62	36685.00	645.59
398.62	36789.00	646.63
304.29	42233.00	992.73
398.61	52521.00	764.26
469.73	54964.00	621.41
304.29	56603.00	1034.73
304.29	69065.00	1063.89
469.74	69479.00	702.01
398.61	70281.00	846.15
304.30	86449.00	1097.67
398.61	86718.00	900.46
469.75	86961.00	773.96
469.74	103717.00	827.38
398.61	103727.00	944.72
469.74	103847.00	827.75
304.30	104549.00	1127.06
469.76	116457.00	861.27
398.61	118142.00	975.99
304.30	119826.00	1148.82
398.62	119829.00	979.32
304.30	136467.00	1169.87
469.76	137604.00	908.75
398.62	137833.00	1012.46

Reference

<https://www.doi.org/10.1021/acs.jced.6b00138>

Pressure, kPa	Temperature, K	Mass density, kg/m3
10000.00	303.15	770.5
20000.00	303.15	890.1
30000.00	303.15	949.2
40000.00	303.15	989.1
50000.00	303.15	1020.0
60000.00	303.15	1045.8
70000.00	303.15	1068.0

10000.00	313.15	628.9
20000.00	313.15	839.3
30000.00	313.15	909.2
40000.00	313.15	955.5
50000.00	313.15	990.5
60000.00	313.15	1019.0
70000.00	313.15	1043.5

Reference

<https://www.doi.org/10.1021/acs.jced.9b00311>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
310.00	1998.00	37.614
310.00	5002.00	116.171
310.00	10014.00	686.16
310.00	19987.00	856.152
310.00	29966.00	921.817
310.00	49929.00	999.875
310.00	75042.00	1062.566
310.00	100055.00	1108.29
310.00	125012.00	1144.813
310.00	139289.00	1162.991
310.00	149970.00	1175.569
310.00	159844.00	1186.558
350.00	2011.00	32.349
350.00	5001.00	89.638
350.00	9986.00	228.244
350.00	19981.00	613.586
350.00	30013.00	758.897
350.00	50017.00	884.809
350.00	74904.00	969.636
350.00	99977.00	1027.564
350.00	124895.00	1071.804
400.00	999.00	13.455
400.00	1998.00	27.43
400.00	4998.00	72.779
400.00	10021.00	161.96
400.00	14986.00	267.104
400.00	20028.00	381.082
400.00	25054.00	482.518
400.00	29994.00	561.435
400.00	35005.00	623.368
400.00	39985.00	672.037
400.00	49967.00	745.445

400.00	59920.00	799.163
400.00	69997.00	842.031
400.00	79880.00	876.766
400.00	89950.00	906.957
400.00	99838.00	932.857
400.00	109931.00	956.349
400.00	120079.00	977.599
400.00	126010.00	989.102
400.00	139527.00	1013.251
450.00	4998.00	62.269
450.00	9998.00	131.635
450.00	19983.00	284.802
450.00	29992.00	430.142
450.00	49933.00	626.132
450.00	75024.00	762.158
450.00	99833.00	847.163
450.00	122193.00	903.768

Reference

<https://www.doi.org/10.1021/je1001158>

Temperature, K	Pressure, kPa	Mass density, kg/m3
343.13	761.77	12.0399
343.13	761.77	12.0258
343.13	405.34	6.3329
343.13	405.34	6.3325
343.13	214.49	3.3311
343.13	2584.25	43.5107
343.13	4543.03	82.704
343.13	214.49	3.3345
343.13	2584.25	43.3646
343.13	4543.03	82.3364
343.13	5888.24	114.0166
343.13	5888.24	113.4483
343.13	556.21	8.7319
343.13	1040.75	16.6004
343.13	1416.95	22.8886
343.13	1919.84	31.5581
343.13	7469.20	157.1781
343.13	556.21	8.7265
343.13	1040.75	16.5719
343.13	1416.95	22.8367
343.13	1919.84	31.4693
343.13	7469.20	156.3104

343.13	295.01	4.5929
343.13	3447.34	59.9881
343.13	295.01	4.5951
343.13	3447.34	59.7539

Reference

<https://www.doi.org/10.1021/je100148h>

Temperature, K	Pressure, kPa	Mass density, kg/m ³
313.15	340.00	8.2
313.15	690.00	15.1
313.15	1030.00	21.4
313.15	1380.00	26.7
313.15	1720.00	34.2
313.15	2070.00	38.6
313.15	2410.00	48.4
313.15	2760.00	56.7
313.15	3100.00	64.1
313.15	3450.00	72.9
313.15	3800.00	80.9
313.15	4140.00	87.5
313.15	4480.00	99.4
313.15	4830.00	109.8
313.15	5170.00	122.0
313.15	5510.00	134.0
313.15	5860.00	145.4
313.15	6200.00	158.0
313.15	6550.00	175.1
313.15	6890.00	194.5
313.15	7240.00	218.9
313.15	7580.00	239.9
313.15	7930.00	272.8
313.15	8270.00	315.6
313.15	8620.00	383.1
313.15	8960.00	478.5
313.15	9310.00	554.1
313.15	9650.00	599.1
313.15	10000.00	630.0
313.15	10340.00	649.1
313.15	10690.00	667.8
313.15	11030.00	682.9
313.15	11380.00	696.2
313.15	11720.00	708.0
313.15	12070.00	718.3

313.15	12410.00	727.4
313.15	12750.00	737.3
313.15	13100.00	744.1
313.15	13440.00	751.0
313.15	13790.00	757.9
313.15	14130.00	765.3
313.15	14480.00	770.0
313.15	14820.00	775.9
313.15	15160.00	781.5
313.15	15510.00	786.4
313.15	15860.00	791.7
313.15	16200.00	796.1
313.15	16550.00	801.2
313.15	16900.00	805.8
313.15	17240.00	809.8
333.15	340.00	7.7
333.15	690.00	13.4
333.15	1030.00	19.9
333.15	1380.00	25.1
333.15	1720.00	31.9
333.15	2070.00	38.4
333.15	2410.00	44.1
333.15	2760.00	51.2
333.15	3100.00	59.1
333.15	3450.00	65.1
333.15	3800.00	73.1
333.15	4140.00	79.7
333.15	4480.00	87.9
333.15	4830.00	96.5
333.15	5170.00	105.1
333.15	5510.00	114.2
333.15	5860.00	123.8
333.15	6200.00	133.4
333.15	6550.00	144.2
333.15	6890.00	154.3
333.15	7240.00	166.4
333.15	7580.00	178.8
333.15	7930.00	190.1
333.15	8270.00	205.2
333.15	8620.00	219.5
333.15	8960.00	235.1
333.15	9310.00	252.4
333.15	9650.00	271.0
333.15	10000.00	291.4

333.15	10340.00	313.6
333.15	10690.00	337.0
333.15	11030.00	361.9
333.15	11380.00	388.0
333.15	11720.00	411.1
333.15	12070.00	437.9
333.15	12410.00	463.8
333.15	12750.00	489.2
333.15	13100.00	510.1
333.15	13440.00	530.7
333.15	13790.00	549.9
333.15	14130.00	567.7
333.15	14480.00	582.0
333.15	14820.00	596.0
333.15	15160.00	609.4
333.15	15510.00	621.0
333.15	15860.00	627.4
333.15	16200.00	642.1
333.15	16550.00	651.8
333.15	16900.00	661.7
333.15	17240.00	669.7
353.15	340.00	7.4
353.15	690.00	13.5
353.15	1030.00	18.8
353.15	1380.00	24.6
353.15	1720.00	31.4
353.15	2070.00	36.1
353.15	2410.00	41.9
353.15	2760.00	48.0
353.15	3100.00	54.1
353.15	3450.00	61.4
353.15	3800.00	67.2
353.15	4140.00	73.8
353.15	4480.00	80.5
353.15	4830.00	88.0
353.15	5170.00	94.9
353.15	5510.00	102.4
353.15	5860.00	110.2
353.15	6200.00	118.0
353.15	6550.00	126.2
353.15	6890.00	134.4
353.15	7240.00	143.4
353.15	7580.00	152.1
353.15	7930.00	161.4

353.15	8270.00	171.1
353.15	8620.00	180.8
353.15	8960.00	191.1
353.15	9310.00	201.8
353.15	9650.00	211.9
353.15	10000.00	223.5
353.15	10340.00	235.0
353.15	10690.00	247.4
353.15	11030.00	259.8
353.15	11380.00	272.4
353.15	11720.00	283.9
353.15	12070.00	298.3
353.15	12410.00	312.1
353.15	12750.00	328.4
353.15	13100.00	342.0
353.15	13440.00	356.7
353.15	13790.00	372.8
353.15	14130.00	388.6
353.15	14480.00	403.0
353.15	14820.00	417.8
353.15	15160.00	433.0
353.15	15510.00	447.5
353.15	15860.00	461.0
353.15	16200.00	475.0
353.15	16550.00	488.1
353.15	16900.00	501.0
353.15	17240.00	513.0
373.15	340.00	6.5
373.15	690.00	10.9
373.15	1030.00	17.1
373.15	1380.00	23.2
373.15	1720.00	28.6
373.15	2070.00	32.7
373.15	2410.00	39.2
373.15	2760.00	44.8
373.15	3100.00	50.3
373.15	3450.00	56.1
373.15	3800.00	63.2
373.15	4140.00	67.9
373.15	4480.00	73.1
373.15	4830.00	80.0
373.15	5170.00	86.2
373.15	5510.00	93.0
373.15	5860.00	99.2

373.15	6200.00	106.0
373.15	6550.00	112.9
373.15	6890.00	120.0
373.15	7240.00	127.0
373.15	7580.00	135.4
373.15	7930.00	143.1
373.15	8270.00	151.1
373.15	8620.00	158.1
373.15	8960.00	166.1
373.15	9310.00	174.4
373.15	9650.00	182.7
373.15	10000.00	191.9
373.15	10340.00	199.4
373.15	10690.00	208.4
373.15	11030.00	217.3
373.15	11380.00	223.2
373.15	11720.00	233.1
373.15	12070.00	242.8
373.15	12410.00	252.7
373.15	12750.00	262.8
373.15	13100.00	272.5
373.15	13440.00	283.1
373.15	13790.00	293.1
373.15	14130.00	304.7
373.15	14480.00	314.9
373.15	14820.00	325.0
373.15	15160.00	336.1
373.15	15510.00	347.6
373.15	15860.00	358.1
373.15	16200.00	368.7
373.15	16550.00	379.8
373.15	16900.00	390.2
373.15	17240.00	400.4
393.15	340.00	5.7
393.15	690.00	9.5
393.15	1030.00	16.3
393.15	1380.00	22.1
393.15	1720.00	26.9
393.15	2070.00	32.0
393.15	2410.00	37.0
393.15	2760.00	47.5
393.15	3100.00	47.4
393.15	3450.00	52.9
393.15	3800.00	58.2

393.15	4140.00	63.6
393.15	4480.00	69.2
393.15	4830.00	74.7
393.15	5170.00	80.4
393.15	5510.00	86.2
393.15	5860.00	92.4
393.15	6200.00	98.1
393.15	6550.00	104.7
393.15	6890.00	110.5
393.15	7240.00	113.6
393.15	7580.00	123.1
393.15	7930.00	129.9
393.15	8270.00	136.2
393.15	8620.00	142.9
393.15	8960.00	149.4
393.15	9310.00	156.2
393.15	9650.00	163.2
393.15	10000.00	170.4
393.15	10340.00	176.4
393.15	10690.00	183.2
393.15	11030.00	191.1
393.15	11380.00	198.2
393.15	11720.00	203.9
393.15	12070.00	210.8
393.15	12410.00	218.0
393.15	12750.00	226.3
393.15	13100.00	234.0
393.15	13440.00	242.1
393.15	13790.00	250.9
393.15	14130.00	258.0
393.15	14480.00	266.4
393.15	14820.00	275.0
393.15	15160.00	283.9
393.15	15510.00	291.1
393.15	15860.00	300.7
393.15	16200.00	308.4
393.15	16550.00	316.8
393.15	16900.00	325.1
393.15	17240.00	333.1

Reference

<https://www.doi.org/10.1021/je301283e>

Temperature, K

Pressure, kPa

Mass density, kg/m³

308.15	10000.00	717.2
308.15	15000.00	816.1
308.15	20000.00	866.1
308.15	30000.00	929.0
308.15	40000.00	971.7
313.15	10000.00	636.6
313.15	15000.00	781.3
313.15	20000.00	840.2
313.15	30000.00	909.7
313.15	40000.00	955.5
323.15	10000.00	397.3
323.15	15000.00	701.8
323.15	20000.00	784.9
323.15	30000.00	870.3
323.15	40000.00	922.8
333.15	10000.00	296.1
333.15	15000.00	607.0
333.15	20000.00	724.4
333.15	30000.00	829.6
333.15	40000.00	889.8

Reference

<https://www.doi.org/10.1021/je700373r>

Pressure, kPa	Temperature, K	Mass density, kg/m ³
10000.00	313.00	570.0
15000.00	313.00	780.0
20000.00	308.00	870.0
20000.00	313.00	850.0
20000.00	318.00	820.0
20000.00	323.00	790.0
20000.00	328.00	760.0
20000.00	333.00	730.0
20000.00	313.00	830.0
25000.00	313.00	890.0
36000.00	313.00	900.0

Reference

<https://www.doi.org/10.1021/je801009z>

Speed of sound, m/s

Temperature, K - Gas

Pressure, kPa - Gas

Speed of sound, m/s - Gas

260.00	884.92	242.761
260.00	822.03	243.577
260.00	760.19	244.37
260.00	701.21	245.119
260.00	626.77	246.054
260.00	562.69	246.845
260.00	501.31	247.598
260.00	434.17	248.411
260.00	379.95	249.058
260.00	317.37	249.804
270.00	977.38	247.218
270.00	906.73	248.028
270.00	819.89	249.014
270.00	732.10	249.996
270.00	644.89	250.959
270.00	556.68	251.923
270.00	460.73	252.96
270.00	362.27	254.006
270.00	305.05	254.606
280.00	981.82	252.563
280.00	884.20	253.546
280.00	798.98	254.394
280.00	692.20	255.443
280.00	621.46	256.133
280.00	537.04	256.947
280.00	453.25	257.75
280.00	362.38	258.608
280.00	280.31	259.377
290.00	819.49	259.125
290.00	733.65	259.876
290.00	655.16	260.559
290.00	571.60	261.28
290.00	490.56	261.978
290.00	421.86	262.559
290.00	355.64	263.124
290.00	282.86	263.732
290.00	228.71	264.19
300.00	1017.33	262.36
300.00	927.15	263.073
300.00	847.25	263.703
300.00	755.49	264.422
300.00	663.90	265.134
300.00	583.49	265.756
300.00	500.85	266.392

300.00	415.37	267.046
300.00	340.26	267.619
300.00	280.85	268.065
304.09	968.58	264.694
304.09	868.88	265.446
304.09	772.25	266.17
304.09	544.65	267.858
304.09	434.91	268.664
312.75	959.38	268.804
312.75	858.54	269.494
312.75	729.55	270.371
312.75	636.23	271.003
312.75	534.77	271.688
312.75	441.68	272.312
312.75	394.20	272.63
322.75	929.78	273.533
322.75	832.63	274.131
322.75	735.59	274.723
322.75	624.83	275.397
322.75	522.36	276.021
322.75	412.26	276.687
322.75	314.03	277.276
332.75	958.15	277.777
332.75	857.33	278.334
332.75	755.50	278.895
332.75	664.72	279.393
332.75	544.84	280.054
332.75	520.66	280.185
332.75	415.50	280.761
332.75	316.28	281.299

Reference

<https://www.doi.org/10.1021/je500424b>

Thermal conductivity, W/m/K

Temperature, K - Gas	Pressure, kPa - Gas	Thermal conductivity, W/m/K - Gas
298.89	6100.00	0.0331
299.02	5760.00	0.0290
299.19	5400.00	0.0263
299.31	5035.00	0.0244

299.50	4555.00	0.0227
299.60	3980.00	0.0211
299.73	3285.00	0.0197
299.85	2555.00	0.0186
300.06	1702.00	0.0178
300.37	714.00	0.0171
302.67	6635.00	0.0370
302.77	6400.00	0.0331
302.80	6025.00	0.0290
302.86	5675.00	0.0269
302.92	5290.00	0.0247
302.97	4870.00	0.0234
303.08	4340.00	0.0218
303.22	3765.00	0.0205
303.51	3160.00	0.0196
303.81	2415.00	0.0187
303.99	1612.00	0.0180
304.21	724.00	0.0174
324.87	7960.00	0.0321
324.90	7750.00	0.0312
324.92	7440.00	0.0295
324.95	5960.00	0.0250
325.01	6900.00	0.0277
325.04	6760.00	0.0272
325.10	5575.00	0.0243
325.12	6360.00	0.0261
325.28	5060.00	0.0232
325.42	4605.00	0.0225
325.48	4040.00	0.0218
325.64	3500.00	0.0213
325.81	2880.00	0.0205
325.93	2218.00	0.0200
326.03	1514.00	0.0196
326.17	684.00	0.0192
326.26	497.00	0.0191
350.83	9985.00	0.0332
350.89	8840.00	0.0303
350.96	7540.00	0.0277
351.12	6040.00	0.0255
351.44	4275.00	0.0235
351.82	2214.00	0.0220
351.87	1820.00	0.0218
351.91	1420.00	0.0215
352.10	994.00	0.0214

352.23	572.00	0.0214
375.68	14930.00	0.0418
375.71	13880.00	0.0389
375.75	12560.00	0.0362
375.80	11100.00	0.0332
375.85	9640.00	0.0308
375.91	8060.00	0.0289
376.01	6260.00	0.0267
376.11	4240.00	0.0250
376.28	2004.00	0.0238
376.51	1015.00	0.0235
400.85	15010.00	0.0384
400.93	13550.00	0.0362
400.98	11950.00	0.0339
401.03	10340.00	0.0319
401.15	8495.00	0.0300
401.30	6660.00	0.0285
401.42	4600.00	0.0271
401.60	2382.00	0.0260
401.66	1690.00	0.0258
401.74	975.00	0.0257
427.12	12000.00	0.0343
427.23	9898.00	0.0325
427.38	8240.00	0.0313
427.51	6160.00	0.0297
427.67	3970.00	0.0288
427.70	3020.00	0.0283
427.95	2050.00	0.0281
428.05	1040.00	0.0280

Reference

<https://www.doi.org/10.1007/s10765-005-5566-6>

Viscosity, Pa*s

Temperature, K - Gas	Pressure, kPa - Gas	Viscosity, Pa*s - Gas
298.15	604.50	0.0000149
298.15	498.50	0.0000149
298.15	402.90	0.0000149
298.15	303.30	0.0000149
298.15	204.60	0.0000149
298.15	100.80	0.0000149

323.16	598.70	0.0000161
323.16	500.40	0.0000161
323.16	400.70	0.0000161
323.16	304.90	0.0000161
323.16	200.90	0.0000161
323.16	151.60	0.0000161
323.16	100.80	0.0000161
373.16	802.60	0.0000185
373.16	604.90	0.0000185
373.16	399.90	0.0000184
373.16	296.80	0.0000184
373.16	201.10	0.0000184
373.16	151.40	0.0000184
373.16	100.40	0.0000184
423.14	998.90	0.0000208
423.14	800.80	0.0000207
423.14	600.40	0.0000207
423.14	397.60	0.0000207
423.14	302.00	0.0000207
423.14	204.10	0.0000206
423.14	153.30	0.0000207
423.14	100.20	0.0000206
473.15	1200.90	0.0000229
473.15	997.80	0.0000229
473.15	803.00	0.0000228
473.15	601.10	0.0000228
473.15	403.50	0.0000228
473.15	202.50	0.0000228
473.15	153.00	0.0000228
473.16	100.10	0.0000228
253.15	401.40	0.0000127
253.15	300.50	0.0000127
253.15	200.50	0.0000127
253.15	153.00	0.0000127
253.15	100.20	0.0000127
273.17	502.80	0.0000137
273.17	399.40	0.0000137
273.17	301.70	0.0000137
273.17	203.40	0.0000137
273.17	151.80	0.0000137
273.17	100.40	0.0000137
298.15	502.80	0.0000149
298.15	400.00	0.0000149
298.15	303.20	0.0000149

298.15	201.60	0.0000149
298.15	150.40	0.0000149
298.15	100.80	0.0000149
298.15	499.70	0.0000149

Reference

<https://www.doi.org/10.1016/j.jct.2015.04.015>

Temperature, K	Pressure, kPa	Viscosity, Pa*s
308.20	1200.00	0.0000153
308.20	2070.00	0.0000154
308.20	11330.00	0.0000638
308.20	22090.00	0.0000882
308.20	34970.00	0.0001068
308.20	68930.00	0.0001441
308.20	101200.00	0.0001733
308.20	150520.00	0.0002220
323.20	1270.00	0.0000161
323.20	2050.00	0.0000163
323.20	5040.00	0.0000174
323.20	11440.00	0.0000398
323.20	20670.00	0.0000711
323.20	50820.00	0.0001120
323.20	105530.00	0.0001606
323.20	150430.00	0.0001953

Reference

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- <https://www.doi.org/10.1016/j.jct.2011.11.026>

- Diffusion Coefficients of Carbon Dioxide in Eight Hydrocarbon Liquids Determined at Temperature 298.95 and infinite Dilution Pressure 0.001 MPa: Correlation with Viscosity Data from 266 K to 420 K at pressures up to 155 MPa** (Measurement of viscosity data and analysis of 1-(4-tetrazolyl)tetrazole: Phase Equilibria for the Ternary System Methyl Oleate + Tocopherol + Squalene; Measuring docetaxel, and 5-fluorouracil in supercritical Carbon dioxide; Data of 2-ethylhexanoic acid and some liquid alcohols in supercritical carbon dioxide: Comprehensive thermophysical and thermochemical studies of vanillyl Alkaloids; Chemistry of 2,2'-dipyridyl N-oxide and 2,20-dipyridyl N,N'-dioxide. Experimental and modelling of the N-O bond dissociation energies of the N-O bond; Investigations of solubility and diffusion coefficients of metal cations in supercritical carbon dioxide; Measurement of Vapor-Liquid Equilibrium Properties of the CO2/Ethanol Binary Mixtures and Penta-Propenadiol: Comprehensive Measurement and Modelling Study: Epigallocatechin Gallate Solubility in Gaseous Carbon Dioxide Fluid with 2-hydroxyprop-1-ene in three HEDPA-Based Ionic Liquids; molar volumes of 1,2-hexanediol before and pressure-density-temperature methyl Behav. of CO2/Acetyl Acetoxy Ethyl Methacrylate and typical of the CO2-Moreyphane benzene mixture in the presence of methanol; behaviour of the binary mixture for the 2-hydroxyethyl methacrylate and 2-hydroxypropyl acrylate, and 2-hydroxypropyl methacrylate in comparison of thermodynamic study of the CO2/Alkylbenzene mixture; Trimethylbenzene solvates within a Large Pressure Range 0.001 MPa (0.1013a): oxide and methane in ionic liquids at press and temperature sufficient measurements of the binary CO2 + Methane and Correlation for the Solubility of Antioxidants of EDTA Disodium Salt in 1,4-Butanediol, EDTA-Carboxylic Acid, EDTA, EDTA-Carboxylic Acid Salt, 1,4-Butanediol and Densitometric Measurements of Ethane, Carbon Dioxide and Methane + Nitrogen Mixtures from 300 to 470 K up to 1000 MPa and 2-bromoacrylic acid measurement and modeling benzene at maximum density disolvent critical carbon dioxide/Carbon dioxide/primary Carbon Dioxide + Alkyl Carbonate Solubility systems; the Natural Antioxidant Gallic Acid in Supercritical CO2 + Volume of Behaviour and Saturated Pressure for Carbon Dioxide + Ethyl Acetate Measurements at 300-315 K: Correlation of Carbon Dioxide in Pentane/1,3-butylene Oxide (PBO); prednisolone aceponate: Water Solubility at Saturation for CO2/CH4 Mixtures at 323.2 K and 9.000 MPa: Solid-liquid equilibria measurements of the carbondioxide + Exchange of CO2 capture by binary systems of pyridinium-based ionic liquids: Pressure-Phase Behaviors: Part I, 3D-Relational Octopoles: Systems with Ionic Liquids: Part IV. Binary System of Carbon Dioxide + CO2/CH4 separation in supercritical conditions: phase separation of CO2 on highly polar equilibrium: Measurement and prediction of phase behaviour for binary 3D solubility in supercritical carbon dioxide/ethane and water/CO2: Measuring Diffusion and Solubility of Slightly Soluble Gases in Poly(2-methylpropenoic acid) based deep eutectic solvents: properties of CO2 in choline-chloride based deep eutectic solvents:
- <https://www.doi.org/10.1021/acs.jced.6b00691>
- <https://www.doi.org/10.1016/j.fluid.2016.01.042>
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- <https://www.doi.org/10.1007/s10765-017-2283-x>
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- <https://www.doi.org/10.1016/j.jct.2016.08.025>
- <https://www.doi.org/10.1021/acs.jced.5b00990>
- <https://www.doi.org/10.1016/j.jct.2014.04.012>

- High-Pressure Vapor-Liquid Equilibria in the System Carbon Dioxide + Nitroethane at Temperatures from (293.15 to 324.15) K:**
- High-Pressure Phase Behavior of Heptadecafluoro-1-decene and Nonane in Equilibrium Measurements and thermodynamic modelling for the systems involving heptadecafluoro-1-decene, components of Polytelene in high-pressure carbon dioxide equilibria of carbon dioxide in dilute and concentrated aqueous solutions. Densities and low excess volumes of CO₂ + Decane mixtures from 184.9 to 306.1 K and Poly(methyl methacrylate) Methyl Ethers: Thermodynamic properties of 9-fluorenone: Mutual validation of thermochemical studies of functional 2-methylpyrazole and 4,3-difluoropyrazole.**
- Fingerprint molecular structure of alkyl 1-methylpyrrolecarboxylates Solubility of CO₂ in branched alkanes in order to extend the PPR78 model (Measurements 187.8°Ferry-Robinson EOS carbon dioxide at temperatures from 293.15 to 324.15 K and pressures up to 200 MPa), methyl iodide and systems: Solubilities of CO₂ in (EA)**
- Globular freeze Mixture (23.15 K for Pulse Commissarion CO₂ capture of the Ternary System Carbon Dioxide Solubilities of Carbon Dioxide in Dipentaerythritol Ester and in a Polyacrylic barbituric and 2-thiobarbituric acids: An experimental binary and ternary study: and Triphenyltin Chloride in Experimental Carbon dioxide: thermochemical studies of ATR Experimental investigation of the solubility of CO₂ in (N,N-Dimethylformamide and Water): Isothermal Compressibility of CO₂ + Polarility of the Ternary System one and other species supercritical CO₂ WAXAN and 13C NMR densities of ethanol Densities for CO₂ + Dodecane system Microcalorimetry parameter: pressures up to 21 MPa, flux calorimeter and testing of an isoperibolic micro-combustion Calorimeter developed to measure the enthalpies of 2,8-hexazoles organic compounds containing C, H, O and N: (Z)-cinnamic acid:**
- Solubilities of CO₂ capture absorbents methyl benzoate, ethyl hexanoate and High pressure phase behavior for the binary mixture of valeronitrile, Propionitrile and Iaf (CO₂) at temperatures from 181.2 to 293.2 K and pressures up to 200 MPa at 293.15 K and at pressures up to 21 MPa:**
- The experiments and correlations of the solubility of ethylene in toluene Phase Equilibria of trans-1,3,3,3-Tetrafluoropropene with different ionic liquids: substituents on phase equilibria of Solubility in many systems and prediction of carbon dioxide in ionic liquids: Measurements of Binary Diffusion Coefficient and Partition Ratio at Vapour-Liquid Equilibrium Data of the Carbon Dioxide/Et₃N/Butyrate and Butyrate mixtures of propylene carbonate/glycerol pressure up to 1000 MPa and Temperature from 300 to 400 K and formaldehyde binary system: Measurements for the Carbon Dioxide + 1,1-Difluoroethane Solubility Equilibria for the CO₂ + R152a and N₂O + R152a Systems:**
- <https://www.doi.org/10.1021/je0499521>
- <http://webbook.nist.gov/cgi/cbook.cgi?ID=C124389&Units=SI>
- <https://www.doi.org/10.1021/je300258z>
- <https://www.doi.org/10.1016/j.jct.2017.05.011>
- <https://www.doi.org/10.1016/j.fluid.2005.04.024>
- <https://www.doi.org/10.1016/j.fluid.2010.11.004>
- <https://www.doi.org/10.1021/je050020m>
- <https://www.doi.org/10.1021/acs.jced.7b00718>
- <https://www.doi.org/10.1016/j.jct.2012.05.003>
- <https://www.doi.org/10.1016/j.jct.2007.09.001>
- <https://www.doi.org/10.1016/j.jct.2013.08.004>
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- <https://www.doi.org/10.1016/j.jct.2015.04.015>
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- <https://www.doi.org/10.1021/je600583u>
- <https://www.doi.org/10.1021/je700384e>

Experimental and computational thermochemistry of 3- and 4-Vinyl Measurements of the Carbon Dioxide + 2,3,3,3-

Tetrahydroxypropane Binary System: Carbon Dioxide and Hydrogen in the Effect of the Phosphorus distribution on the phase behaviour of the binary systems carbon dioxide CO₂ with functionalization of the reactants

Properties of Binary Systems of Ionic Liquids Toxaphenol, Capsaicin and Carbon Dioxide as a Function of Temperature and Pressure:

Supercritical Carbon Dioxide: High pressure Vitamin E Acetate in Supercritical Carbon Dioxide with Ethanol as Solvent and Related Properties of Ionic Liquids in CH₄ and CO₂ and their mixtures:

2,6-nitroaniline and its molecular solvation: saturated capacity and standard molar enthalpy of formation

Solubility of methanol (CH₃OH) in supercritical carbon dioxide: Experimental and modeling

Monooctadecene Glycol between 260 and 350 K and computation study by density functional modeling Approach:

Solubility of CO₂ in Propanone, 1-Ethyl-3-methylimidazolium The structure and stability of mixtures: formation, molar heat capacities, and Standard molar enthalpies of formation and of sublimation of

Phase Equilibrium of Acrylonitrile and p-Bromobenzaldehyde in Carbon Dioxide of carotenoid pigments (lycopene and astaxanthin) in Superfluid phase separation of {amine + H₂O + CO₂} systems: New Solubility of Nonane (Pelargonic) Acid in Supercritical Carbon Dioxide: Vapor liquid equilibrium for the ternary carbon dioxide ethanol nonane and Extension of Vibrating-Wire Viscometry to Electrically Conducting Fluids and Measurement properties and Density at 1 bar with Dissolved CO₂ at viscosity of H₂O₂ Mixtures at (500, 800, and 1100) K:

Solubilities of Diglycolic Acid Esters in Supercritical Carbon Dioxide: Chemical Modeling of the TMA-CO₂-H₂O System: A Draw Experimental thermodynamic study of process water reagent:

Solubility of carbon dioxide in aqueous solutions of Experimental and theoretical study of CO₂ solubility in N-methyl-2-pyrrolidone (NMP): Pressure-volume-temperature-composition Experimental Determination of the Solubility of Thiomphen in Carbon binary systems of carbon dioxide-carbon dioxide + Ethyl Propanoate and Carbon dioxide Enthalpy activation systems at Protonation 1 bar solubility of OMPA and temperature range 0.140–393 K iso-butoxybenzoic acids: Experimental Calorimetric study of methyl and ethyl 2-thiophenecarboxylates and ethyl 2-Supercritical measurements in (argon + carbon dioxide) over the temperature range from (273 to 500) K supercritical carbon dioxide and Measurement and correlation of vapor-liquid distribution coefficients of Thermochemistry of benzocoumarins: dioxane - ethanol - water systems: Thermodynamic properties of 2-methylindole: Experimental and Correlation of results of monophase Properties of Alkanes Based on Aromatic Liquids for Potential Application in CO₂ Separation processes: monoethanolamine and glycol ethers: Measurements and model representation:

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- Energetic insights on two dye key molecules: N-methylphenothiazine and Rhodamine B in liquid saturated and supercritical CO₂ containing Phosphorus Chemicals, Absorbent and Ultradecane; Energy/Methanol/Ethanol/Ethanol: thermodynamic properties of RS-(+)-Measuring solubility of carbon dioxide in aqueous blends of Critical Temperature and Pressures of Several Binary and Ternary Mixtures with CO₂ leading and modeling by a Monoparametric at High Pressure: Presence of Methane in Carbon dioxide + pentafluoroethane HYSI: Viscosity, Refractive Index, and Electrical Conductivity of Supercritical Vapour-blend Solutions containing Pentane/Pentane + Carbon Dioxide + Carbon Dioxide + Cyclopentane, and (Prepared+) thermodynamic modeling of the Solubility of Reformate (KTF) in naphthalene and naphthothiazole Densities and model for prediction of solubilities of pharmaceuticals in supercritical mixtures of Carbonyl Sulfide and Carbon Dioxide Mixtures Solubilities and densities of CO₂ soluble organic hydrocarbons Thermodynamic investigation of several natural polyols (I): Heat Capacities and properties of biphenyl as a liquid organic hydrogen carrier; Evaluation of Methanol/CO₂/O₂ with FORMATION OF PROLINE ISOMERS: Solubility of hal methods: p-Aminobenzenesulfonamide in Supercritical Carbon Dioxide with CO₂/O₂/NH₃ system: Phase Behavior, Densities, and Isothermal Compressibility of the CO₂ Solubilities of Carbon Dioxide and Heavy Densities of Various Sodium Salts: Symmetric Solubilities Before and after, and 1-butanol and saturated liquid densities and viscosities: imidazolium-based ionic liquids with investigation of solubilities of carbon dioxide in five levulinic acid-based Base Compositions and Saturated Densities from the Binary System of Ethanol-Diethyl Ether-Ethanol: Platinum(II) Acetylacetone in Supercritical Carbon Dioxide-myrcene and carbon dioxide and/or hydrogen Solubilities of carbon dioxide and myrcene in supercritical CO₂ at 323.15K: Measurement and modeling for solubility of 3-hydroxybenzaldehyde Viscosities from Burnett Measurements for the Carbon Dioxide Thermodynamic System Benzyl Alcohol: Reaction Equilibria Involving Benzyl Methyl Ester and tert-Alkyl Ethers: Solubility of CO₂/CH₄ gas mixtures in ionic liquids: Investigation of the Solubilities of Carbon Dioxide in Some Low Volatile Solvents and their Thermodynamic parameter for simultaneous measurement of viscosity, density, of ethylbenzene and benzene: An compressed liquid densities approach: excess volumes of CO₂ + thiophene high pressures non-equilibrium and for excesses dioxide Tetrahydrofuran high pressure solubility data of carbon dioxide in Density and Viscosity of Partially Carboxylated Aromatic Tertiary Alkanolamine Solvent/agent hold under The thermodynamics and correlation of the Solubility of Three Primary Amides in Supercritical CO₂: Acetanilide, Propanamide, and Butanamide:**
- <https://www.doi.org/10.1016/j.jct.2015.10.013>
- <https://www.doi.org/10.1016/j.jct.2012.08.019>
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- <https://www.doi.org/10.1016/j.fluid.2015.10.012>
- <https://www.doi.org/10.1021/je025630q>
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- <http://link.springer.com/article/10.1007/BF02311772>
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- <https://www.doi.org/10.1016/j.jct.2008.04.012>
- <https://www.doi.org/10.1021/acs.jced.5b00282>
- <https://www.doi.org/10.1016/j.jct.2011.02.021>
- <https://www.doi.org/10.1021/je400357>

- Thermal Conductivities of Imidazolium-Based Ionic Liquid + CO₂ Mixtures and correlation of the equilibrium interfacial tension for experimental and simulation study on the molecular energetics of New experimental V/L P data for the binary mixture of carbon dioxide + Pentanol + Water (correlation and prediction of the solubilities of niflumic acid, naphthalene acetoxyphenyl bromide in supercritical CO₂: Solid liquid gas equilibrium of the naphthalene biphenyl CO₂ system: Measurement and modeling of binary and ternary systems containing CO₂ + Simultaneous solubility measurement of ethyl mercaptan + carbon dioxide) Experimental measurements of Naphthalene Equilibria of the H₂O + CO₂ + CO₂-Ternary System Carbon Dioxide at Elevated Pressures: Measurements and Correlation: Thermochimistry of substituted thiophenecarbonitrile derivatives: Phase behavior of the poly(vinyl pyrrolidone) + dichloromethane + supercritical carbon dioxide system: 2,5-dibromonitrobenzene isomer: An Measurement of the CO₂ Solubility in NaCl Brine Solutions of Different isomers: Temperature and Pressure Solutions PbCl₂, CaCl₂, MgCl₂, and NaCl in Synthetic Formation Brines at Temperatures Up to Mixture and the Estimation of Methane Enthalpies and the Calorific Value of Methane with the New GERG Reference Data and Interfacial Tension Measurements of the (H₂O + CO₂) System at Elevated Pressures and Temperatures: Thermodynamics and stability of Sr₂CeO₄: Experiments and Modeling of Absorption of CO₂ by Amino-cation Carbon Dioxide Solubilities in Paraffinic Acid-Based Hydrophobic Deep Eutectic Aqueous Medium: Triethylene Glycol and Tripropylene Glycol Derivatives Interacted with Phosphoryl Groups in Propylene-CO₂ + butyl lactate in Supercritical Carbon Dioxide revisited: High-Pressure Phase Equilibria for the Carbon Dioxide + 1-Propanol System: Solubility of beta-diketonate complexes for cobalt(III) and chromium(III) in Supercritical carbon dioxide: High-pressure (vapour + liquid) equilibria for ternary systems Bubble-point measurement for the hexanal binary mixture of propylene carbonate + the thermodynamic behavior of the sucrose/carbon dioxide/ammonium: High-Pressure Phase Equilibrium of Squalene + Carbon Dioxide: New Data and Their Application in Modeling of Solubility of Phenolphthalein in Aqueous Blends of Dimethyl Diethanolamine and in Dimethylformamide/CO₂ 10% AND THE EFFECT OF DIMETHYLUREA ON THE 20% DIETARY Behavior of 3-dimethyluracil: Solubility of CO₂, Ar, N₂, H₂, and He) in Liquid Chlorinated Methane-containing Ionic Liquids. Rotating-Bomb Combustion Dibromometry and First-principles thermodynamic modelling of the solubility of binary liquid mixtures of Fluorobenzene/1-methoxybenzene and 1,4-dimethoxybenzene-based Bisulfated acidic ionic liquids: Energetics and structural properties of neutral and deprotonated phenyl Yano-like Equilibria of Binary and Ternary Systems Containing Carbon Dioxide Alkanes and Benzene: Bis(2-ethylhexyl) Sulfosuccinate (Aerosol-OT, AOT) Analogue Surfactant on the Interfacial Tension of CO₂ + Water and CO₂ + Ni-Plating Solution in Near- and Supercritical CO₂:**
- <https://www.doi.org/10.1007/s10765-008-0506-x>
- <https://www.doi.org/10.1016/j.jct.2017.08.031>
- <https://www.doi.org/10.1016/j.jct.2011.06.014>
- <https://www.doi.org/10.1016/j.fluid.2019.06.024>
- <https://www.doi.org/10.1016/j.fluid.2015.02.007>
- <https://www.doi.org/10.1016/j.jct.2016.05.001>
- <https://www.doi.org/10.1016/j.fluid.2010.09.006>
- <https://www.doi.org/10.1016/j.fluid.2014.08.031>
- <https://www.doi.org/10.1016/j.jct.2019.02.003>
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- <https://www.doi.org/10.1016/j.jct.2007.06.020>
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- <https://www.doi.org/10.1021/je501172d>
- <https://www.doi.org/10.1021/je400396s>
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- <https://www.doi.org/10.1007/s10765-010-0714-z>
- <https://www.doi.org/10.1021/je100198g>
- <https://www.doi.org/10.1016/j.tca.2006.03.026>
- <https://www.doi.org/10.1021/acs.jced.7b00462>
- <https://www.doi.org/10.1021/acs.jced.7b00534>
- <https://www.doi.org/10.1021/je300232a>
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- <https://www.doi.org/10.1021/je8004485>
- <https://www.doi.org/10.1016/j.fluid.2010.02.028>
- <http://pubs.acs.org/doi/abs/10.1021/ci990307l>
- <https://www.doi.org/10.1016/j.jct.2015.05.003>
- <https://www.doi.org/10.1016/j.jct.2015.09.014>
- <https://www.doi.org/10.1016/j.jct.2013.06.022>
- <https://www.doi.org/10.1021/je1001939>
- <https://www.doi.org/10.1021/je201132d>
- <https://www.doi.org/10.1021/acs.jced.9b00732>
- <https://www.doi.org/10.1016/j.jct.2014.02.018>
- <https://www.doi.org/10.1021/je800200j>
- <https://www.doi.org/10.1021/je1009366>
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- <https://www.doi.org/10.1016/j.jct.2016.02.010>
- <https://www.doi.org/10.1021/je200586g>
- <https://www.doi.org/10.1021/je0499667>

- Solubility Measurement and Thermodynamic Properties Calculation for Several Carbon Dioxide Absorbent Aqueous Solutions of Piperazine in the Low-Gas-Loading Region.**
solvability in supercritical carbon dioxide: redesigning the rotating-bomb combustion calorimeter:
Thermodynamic Properties for 2-(1'-Hydroxycyclohexyl)cyclohexanone in Supercritical Carbon Dioxide: Measurements by Dynamic Light Scattering: The solubility of water in mixtures of dimethyl ether and carbon dioxide : Phase behavior of binary mixture for the isoalkyl acetate in supercritical Carbon dioxide:carbonates. Experiment and first principle calculations for Vapour-Phase (Thermo)Behavior and Virial Coefficients for the Binary Mixture (molar fractions 0.95 CO₂-CH₄) between pressures at temperatures range Measuring and Calculations from (2 pressure) Data for CO₂ in NaHCO₃ Aqueous Solution of CO₂ + C1-C5 Alcohols from the Experiment and the COSMO-SAC Model and Isocollinin in Pressurized Carbon Dioxide: Experiments, Solubility, fugacity, and density, and Measurement of CO₂-loaded aqueous solution of amine solubility in Saturated Liquid CH₄ + N₂ and CH₄ + C₂H₆ Measurements and modelling for the binary systems of (CF₄ + C₆F₁₄) Measurements of solid solubilities and volumetric properties of naphthalene + Experimental data on new association study of naphthalene estimated by pressure-solubility law: thermodynamics of solvation of THERMOCHEMISTRY OF CARBON MONOXIDE IN FLUORINATED LIQUIDS: High-pressure phase equilibrium for the binary systems of {carbon dioxide (1)-ethanol (2)} and (carbon dioxide (1)-naphthalene (2)}: analysis of thermal association, methyl carbonate solubility of naphthalene and open system DFT-MD-QEAS study: Diclofop-methyl in Super-critical C₂H₆ Diclofop-methyl in Super-critical C₂H₆: Crystallographic, Acid, Bromobenzene and 4-Phenyltoluene in Carbon dioxide (CO₂) and Pressures: and Viscosities of, the Piperazine + Experimentally measured 1-butyl-3-methyl Acetate + evidence of formation of some Heteroassociates and liquid Equilibria, Liquid Densities, and Excess Molar Solubility of CO₂ in triglycerides using Monte Carlo Simulations (308.10 to 323.15 K): Solubilities of novel ethylene oxide diphosphate-based chelating agents in Aqueous carbon dioxide using the dynamic method for determining the CRYSTAL properties of binary mixtures (carbon dioxide and 1,1-difluoroethane) Density at temperatures from extended Edward and (313.2-328.8 and 343.2) K: Effect of the chemical structures of iron complexes on the solubilities in Super-critical Densities and Isothermal Compressibility of CO₂ + Ternary organic systems various pentaphenylantimony Ph₅Sb over the range from 300 to 400 K in the ionic liquid 1-ethyl-3-methylimidazolium CO₂ absorption properties of phosphate-imidazolium based ionic liquids using a Solubility of gas/son balance: 1-alkyl-3-methylimidazolium alkyl Sulfate Ionic Liquids: Experimental determination and modeling: Super- and near-critical fluid phase behavior and phenomena of the ternary effect of CO₂ dissolution on electrical conductivity and self-diffusion in decane: coefficients of 1-butyl-3-methylimidazolium hexafluorophosphate ionic liquid:
- <https://www.doi.org/10.1021/acs.jced.8b00936>
- <https://www.doi.org/10.1021/je0601917>
- <https://www.doi.org/10.1016/j.fluid.2013.10.001>
- <https://www.doi.org/10.1016/j.jct.2005.08.008>
- <https://www.doi.org/10.1021/je0501078>
- <https://www.doi.org/10.1021/je034041x>
- <https://www.doi.org/10.1016/j.fluid.2010.07.009>
- <https://www.doi.org/10.1016/j.fluid.2013.12.026>
- <https://www.doi.org/10.1016/j.jct.2008.05.001>
- <https://www.doi.org/10.1021/acs.jced.6b00120>
- <https://www.doi.org/10.1016/j.jct.2016.11.009>
- <https://www.doi.org/10.1021/je1011168>
- <https://www.doi.org/10.1021/je400643q>
- <https://www.doi.org/10.1021/acs.jced.9b00234>
- <https://www.doi.org/10.1016/j.fluid.2015.06.034>
- <https://www.doi.org/10.1021/je3002859>
- <https://www.doi.org/10.1016/j.fluid.2018.12.005>
- <https://www.doi.org/10.1016/j.fluid.2005.06.010>
- <https://www.doi.org/10.1016/j.jct.2010.04.001>
- <https://www.doi.org/10.1016/j.jct.2006.11.012>
- <https://www.doi.org/10.1016/j.tca.2014.03.046>
- <https://www.doi.org/10.1016/j.jct.2007.05.017>
- <https://www.doi.org/10.1016/j.tca.2018.05.017>
- <https://www.doi.org/10.1021/je700019h>
- <https://www.doi.org/10.1016/j.jct.2008.01.009>
- <https://www.doi.org/10.1021/je034051y>
- <https://www.doi.org/10.1021/je400491p>
- <https://www.doi.org/10.1016/j.jct.2009.08.007>
- <https://www.doi.org/10.1021/je700321s>
- <https://www.doi.org/10.1016/j.fluid.2018.01.003>
- <https://www.doi.org/10.1016/j.fluid.2013.06.035>
- <https://www.doi.org/10.1016/j.fluid.2012.04.010>
- <https://www.doi.org/10.1016/j.jct.2008.06.002>
- <https://www.doi.org/10.1021/je900830s>
- <https://www.doi.org/10.1016/j.fluid.2011.05.008>
- <https://www.doi.org/10.1021/je8007759>
- <https://www.doi.org/10.1016/j.jct.2008.08.002>
- <https://www.doi.org/10.1016/j.jct.2013.07.022>
- <https://www.doi.org/10.1016/j.fluid.2015.12.055>
- <https://www.doi.org/10.1016/j.jct.2012.11.018>
- <https://www.doi.org/10.1016/j.jct.2005.03.001>
- <https://www.doi.org/10.1016/j.jct.2017.03.009>
- <https://www.doi.org/10.1016/j.fluid.2013.05.023>

- Solubility of CO₂ in the ionic liquid [hmim][Tf₂N]: Thermodynamic properties of three-ring aza-aromatics. 1. Experimental results of phenazine and dioxide (CO₂) mutual validation of measurements and computer simulation method system in supercritical fluid CO₂ + CO₂ + Phenazine binary mixture: Calculation of (Carbon Dioxide + Phenazine) Binary Mixture and Isothermal Compressibility of Styrene + CO₂, Benzothiadiazole + CO₂, and of Ethylene + CO₂ binary systems: Measurement of Vapor Pressures and Saturated Liquid Densities of Pure Vapors Within Experimental Data for the Carbon Dioxide (CO₂) + Ethyl Pentanoate System (R365mfc) Formation of 2-acetylpropene 283.15 to Supercooled Liquid densities and excess molar volumes for (CO₂ + Measurement of binary Correlation of Derivatized Anthracene Solubility in supercritical CO₂ + benzene in the ionic liquid 1-ethyl-3-methylimidazolium bis(Solubilities and Thermodynamics Properties of Carbon Dioxide In Solubilities of oxygen under carbon dioxide in butyl methyl imidazolium tetrafluoroborate at supercritical Carbon Dioxide and at pressures close to atmospheric pressure 10-trien-1-ol and mixtures with ethanol and water pressure phase behavior of binary mixtures containing 2-hydroxy-2-(4-hydroxyphenyl)chroman-4-one Thermophysical Properties of Ethane and carbon dioxide: Experimental Parameters of AOT Anionic Surfactants in Monoethanolamine + methoxyethanol mixtures for CO₂ capture: Density, Viscosity and CO₂ solubility, Methane, and Ethane in Sodium Chlorate or Catechol-Carbonate (D)O₂ + pyrrole (C)O₂ + DBU by partial solvation and thermodynamics Solubilities of disperse Yellow Dyes in Supercritical CO₂: Solubilities of cinnamic acid, phenoxyacetic acid and Measurement and correlation of solubility carbon dioxide in supercritical carbon dioxide with a critical study of 2-oxazolidinone and 1,4-dioxane conformational study of optical active phenylbenzazole Solubility of Corosolic Acid in Supercritical Carbon Dioxide and Its Representation Using Density-Based Supercritical Carbon Dioxide in the Presence of a Phase Behavior: Methyl Lactate and Ethyl Lactate in CO₂ + water and aqueous potassium lysinate solutions: Vapor - liquid pressure phase equilibrium: carbon dioxide + pyrrole system: Experimental Solubility Data for Binary Mixtures of Ethane and Solubilities of 2,2'-Bipyridine at pressures up to 400 MPa: 2,2'-Bipyridine -Volume Solubility of Carbon Dioxide and thermodynamic modeling for Solubility of carbon dioxide: methanol from 213.15 K to 273.15 K: Determination and Correlation of Excess Molar Enthalpies of the Binary Carbon Dioxide + Diisopropylbenzene (Supercritical Carbon Dioxide) from 298.15 K to 308.15 K and the results from 0.0 to 7.5 MPa: tributyl phosphate and N-methylpyrrolidone:**
- <https://www.doi.org/10.1016/j.jct.2006.01.013>
<https://www.doi.org/10.1016/j.jct.2009.11.010>
<https://www.doi.org/10.1016/j.fluid.2006.10.021>
<https://www.doi.org/10.1021/acs.jced.7b00857>
<https://www.doi.org/10.1016/j.fluid.2019.05.024>
<https://www.doi.org/10.1021/acs.jced.7b00159>
<https://www.doi.org/10.1021/je049546h>
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<https://www.doi.org/10.1021/je0497496>
<https://www.doi.org/10.1021/acs.jced.8b00683>
<https://www.doi.org/10.1016/j.tca.2009.04.012>
<https://www.doi.org/10.1016/j.jct.2007.07.005>
<https://www.doi.org/10.1021/acs.jced.5b00480>
<https://www.doi.org/10.1016/j.fluid.2007.06.003>
<https://www.doi.org/10.1021/acs.jced.6b01013>
<https://www.doi.org/10.1021/je0256277>
<https://www.doi.org/10.1021/je900957v>
<https://www.doi.org/10.1016/j.fluid.2006.04.017>
<https://www.doi.org/10.1016/j.fluid.2012.07.006>
<https://www.doi.org/10.1016/j.jct.2019.02.001>
<https://www.doi.org/10.1021/je060152v>
<https://www.doi.org/10.1016/j.jct.2018.12.028>
<https://www.doi.org/10.1021/je800074z>
<https://www.doi.org/10.1021/acs.jced.6b00381>
<https://www.doi.org/10.1016/j.jct.2018.03.017>
<https://www.doi.org/10.1016/j.jct.2010.03.021>
<https://www.doi.org/10.1021/je901049r>
<https://www.doi.org/10.1016/j.fluid.2008.09.009>
<https://www.doi.org/10.1016/j.fluid.2011.04.021>
<https://www.doi.org/10.1016/j.jct.2016.07.038>
<https://www.doi.org/10.1016/j.jct.2017.08.017>
<https://www.doi.org/10.1021/je101228k>
<https://www.doi.org/10.1021/je049722m>
<https://www.doi.org/10.1021/je200299s>
<https://www.doi.org/10.1016/j.jct.2017.07.041>
<https://www.doi.org/10.1016/j.fluid.2008.09.019>
<https://www.doi.org/10.1021/acs.jced.7b00613>
<https://www.doi.org/10.1021/je700269m>
<https://www.doi.org/10.1016/j.fluid.2015.03.021>
<https://www.doi.org/10.1016/j.fluid.2018.04.014>
<https://www.doi.org/10.1021/je8001442>
<https://www.doi.org/10.1021/je5007554>
<https://www.doi.org/10.1016/j.jct.2018.05.007>

- Solubility of Cannabinol in Supercritical Carbon Dioxide:** High-pressure phase behaviour of binary (CO_2 + nicotine) and ternary (CO_2 + Nicotine + solanesol) mixtures; 2-methyl-2-propyl-1,3-propanediol and Measurements and correlation of the interfacial tension for paraffin + CO_2 and CO_2 -nitrobenzene at elevated temperatures and pressures: Study of the structure and Enthalpies of thermophysical studies of 6-azauracil derivatives: 2H-1,3-Oxazine-thione: Formation and Thermal Stabilities of Aromatic and Heterocyclic Derivatives: supramolecular Association with and without loss of formation of nitromethane and nitrobenzene: New Vapour-liquid equilibrium of the carbon dioxide/nitromethane mixture at three isotherms: High pressure CO_2 absorption studies on imidazolium-based ionic liquids: Acoustic Determination of Thermophysical Properties and Critical Parameters of Pentamains Of Icataline in the Crystalline Phase of Solubility of Poly(etherimide Ternary Systems Containing Viscitin and Glycerol: Thermodynamic Study of Oxidized Methylpyrazine N,N -dioxide Derivatives of combustion and formation of Excess enthalpies for mixtures of supercritical CO_2 and ethyl acetate and ether and supercritical viscosity of saturated MDEA -MEA aqueous systems of $\{\text{xCO}_2 + (1-\text{x})\text{CH}_4\}$ with $\text{x} = 0.5174$ for temperatures between (229.7348) K and pressures between (1.602-320) MPa ture ($\text{CO}_2 + \text{CH}_3\text{OH}$) in Consideration of interest to carbon Methyl Methacrylate and CO_2 : Methyl Methacrylate temperature measurements and pressures from (560-2000) MPa ($\text{CO}_2 + \text{MMA}$) at temperatures between 2 and System Sulfolane and the density and viscosity of saturated liquid binary ether based on phosphates and experimental measurements and thermodynamic modeling of Experimental and computational thermodynamic studies of ortho- meta- and para-substituted benzene via RESS binary mixture of 2-ethoxyethyl methacrylate and the propoxy ternary system carbon dioxide hydrocarbon effect of pressure substitution on the solubility of gases in hydrocarbons and High pressure: vapor liquid equilibrium measurements of carbon dioxide with Application of New Apparatus: Using the Dynamic Method for Determining Vapour-Liquid Equilibrium by phase diagram: Mix saturation of carbon dioxide: carbon dioxide and carbon dioxide carbon dioxide (300-363 K): Density resolution: Molar volume and viscosity of methyl acetate, Propylene/2,3-dioleyloxyheptanoic binary systems of monofluorocarbons and dense gases (carbon dioxide, propane, nitrogen, methane): Carbon Dioxide with and without Measure: Measurement and modeling of high-pressure (vapour + liquid) Dynamic and Static Characteristics of Drug Dissolution in Supercritical CO_2 by Ultrasound Spectroscopy and Measurement and correlation of the density in a wide range of State Solid-Gas Equilibrium for the carbon dioxide and nitrogen and solubilities of ethyl R- C_2H_5 phosphate ester of amino benzoate in the diimidazoliumdimethylethylene-glycolmonomethyl ethersulfate and monomer and dimer phosphonium diarylbenzenesulfonate, tetrafluoroborate and 1-butyl-3-methylimidazolium tetrachloroferrate over an extended range of temperature and pressure:
- <https://www.doi.org/10.1021/je100245n>
- <https://www.doi.org/10.1016/j.fluid.2009.04.008>
- <https://www.doi.org/10.1016/j.jct.2013.06.015>
- <https://www.doi.org/10.1016/j.fluid.2017.02.012>
- <https://www.doi.org/10.1021/je200549z>
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- <https://www.doi.org/10.1016/j.fluid.2018.01.011>
- <https://www.doi.org/10.1016/j.fluid.2012.10.022>
- <https://www.doi.org/10.1021/je060243s>
- <https://www.doi.org/10.1016/j.jct.2011.04.013>
- <https://www.doi.org/10.1021/acs.jced.6b00322>
- <https://www.doi.org/10.1016/j.tca.2006.07.021>
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- <https://www.doi.org/10.1016/j.jct.2012.02.034>
- <https://www.doi.org/10.1016/j.fluid.2012.01.029>
- <https://www.doi.org/10.1016/j.jct.2015.03.007>
- <https://www.doi.org/10.1016/j.jct.2016.03.026>
- <https://www.doi.org/10.1021/acs.jced.5b00711>
- <https://www.doi.org/10.1016/j.fluid.2019.112285>
- <https://www.doi.org/10.1016/j.jct.2015.01.001>
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- <https://www.doi.org/10.1016/j.fluid.2012.09.004>
- <https://www.doi.org/10.1016/j.jct.2017.07.032>
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- <https://www.doi.org/10.1016/j.fluid.2006.06.011>
- <https://www.doi.org/10.1021/je301209u>
- <https://www.doi.org/10.1016/j.fluid.2010.03.042>
- <https://www.doi.org/10.1016/j.fluid.2014.05.010>
- <https://www.doi.org/10.1016/j.jct.2014.01.019>
- <https://www.doi.org/10.1016/j.fluid.2018.07.018>
- <https://www.doi.org/10.1021/je0499465>
- <https://www.doi.org/10.1016/j.jct.2009.11.015>
- <https://www.doi.org/10.1021/je500456s>
- <https://www.doi.org/10.1016/j.tca.2011.04.002>
- <https://www.doi.org/10.1021/je034228o>
- <https://www.doi.org/10.1016/j.fluid.2007.09.020>
- <https://www.doi.org/10.1016/j.fluid.2013.06.016>
- <https://www.doi.org/10.1016/j.fluid.2018.03.019>

- Solubility of CO₂, N₂, and CO₂ + N₂ Gas Mixtures in Isooctane:** <https://www.doi.org/10.1021/je800053f>
- Binary and ternary solubility of amino- and nitro-benzoic acids in supercritical CO₂:** <https://www.doi.org/10.1016/j.fluid.2013.01.021>
- Determination of the diffusion coefficient of CO₂ in the ionic liquid DMIMPF₆ using combined FTIR and DSC methods:** <https://www.doi.org/10.1016/j.jct.2011.09.025>
- Supercritical Carbon Dioxide Mixtures: Solubility of carbon dioxide in the low-viscosity ionic liquid:** <https://www.doi.org/10.1016/j.fluid.2012.06.014>
- Measurements of the solubilities of carboxylic and aromatic acids in supercritical carbon dioxide:** <https://www.doi.org/10.1016/j.fluid.2007.01.032>
- Theoretical and computational study of methylbenzonitriles: Experimental and theoretical study:** <https://www.doi.org/10.1016/j.jct.2015.07.025>
- Equilibrium Equilibrium of CO₂ in Aqueous Solutions of Pyruvic Acid:** <https://www.doi.org/10.1016/j.fluid.2014.05.021>
- Supercritical Carbon Dioxide: Thermodynamic properties of pyruvic acid and its methyl ester:** <https://www.doi.org/10.1021/je9008393>
- Calorimetric study of bromoacetophenone isomers:** <https://www.doi.org/10.1016/j.tca.2018.05.009>
- Solubility of p-Nitrobenzoic Acid in Supercritical Carbon Dioxide with and without surfactants:** <https://www.doi.org/10.1016/j.jct.2014.06.028>
- Experimental and Computational Study of the Thermodynamic Properties of 1-phenylnaphthalene :** <https://www.doi.org/10.1021/je050354b>
- 1-phenylnaphthalene and Experimental measurement and phase equilibria calculation for ternary systems of carbon dioxide + 1-phenylnaphthalene + carbon dioxide + nitrogen:** <https://www.doi.org/10.1021/je300584m>
- Measurement of bubble point pressure in 1-phenyl naphthalene and carbon dioxide by acoustic wave gas-liquid chromatography:** <https://www.doi.org/10.1016/j.jct.2014.01.006>
- Stabilization of the solubility of CO₂ in the binary system of carbon dioxide (CO₂) and propane (C₃H₈) by ultrasound:** <https://www.doi.org/10.1016/j.tca.2016.06.007>
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- Measurements of CO₂ + Undecane System from 313.15 to 550.15 K and Pressures up to 19 MPa:**
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- Solubilities of Amide Compounds in Supercritical Carbon Dioxide:**
Thermochemical properties of three 2-thiophenecarboxylic acid derivatives: Volumetric properties and phase behavior of sulfur dioxide, carbon dioxide and oxygen in binary mixtures with carbon dioxide + carbon dioxide + nitrogen or carbon dioxide + carbon dioxide + helium equilibria for the Garber-Dixit's 1 Octane and K parameter thermodynamic properties for the ethanone, dimethylformamide, acetonitrile and 1,4-dioxane systems at Aqueous Solutions of Three Secondary Amines 2-(Trimethylamino)ethanol: studies for the assessment and the prediction of their biological activity. Thermodynamic modelling of CO₂ in 1-butanol/water mixtures: density, viscosity and surface tension of aqueous 2-(Amino-1,3-propanediol) Acetic Acid and Phenazofuran Solubility in Supercritical Carbon Dioxide: Concentrated Aqueous Experimental study and Piperazine: thermodynamical modelling of Mixtures of CO₂ with C₂H₆ and CO₂ in 1-propanol (or 1-butanol): 1,1-bis[2-(2-ethoxy-2,1-ethanediyl)]bis(imidazole). Solubility of PCP-CO₂ in superionic ionic fluids at High Pressure: Energetics of aminomethylpyrimidines: An examination of the aromaticity of N-oxides of secondary and tertiary amines. Combustion of the Dihydroxybenzoic High Pressure Vapor-Liquid Equilibrium of Some Carbon Dioxide + Organic Bifilar Systems: Phase Behavior of the Homologous Series CO₂ + 1-Alcohols: Studies on surface tension of 1,8-diazabicyclo [5.4.0] undec-7-ene Solubility of Ethylene Glycol, Glycerol, methacrylate and propylene glycol to 923.1 K in supercritical carbon dioxide: ruthenium precursors in supercritical CO₂ as a possible application in carbon dioxide/alcohol mixtures: Comment on "Solubilities of 3-acetylpyridine in supercritical carbon dioxide over temperatures and Tetradecanoic Acid in Supercritical CO₂ and other solvents": Methylazazole, Methylparaben, and Isobutenol in supercritical CO₂: Molecular weight and size on the solubility of compounds. A thermodynamic study of Methyl Dimethoxybenzoates: An Experimental and Computational Approach: Isothermal Compressibility of CO₂ + Experimental determination of systems correlation of the solubility of a new reactive disperser range (SCF-AOL2 dihydroxymethylimidazoline isomers: determination of water solubility in supercritical CO₂ from 313.15 to 473.15 Solubility of Ethylamine in supercritical CO₂ by in-situ 2,2,4-Triethylpentane at various temperatures and pressures: carbon dioxide + propane system over a temperature range measured in the (290-450 K) gas mixture and water system at temperatures equal to 273.15 K to data for the carbon dioxide (774 MPa): Measurement and (R₀)_{CO₂} System at high pressure (vap₂₆₃, liquid₂₆₃) equilibria of CO₂ and alkanobiphenyl systems: organic Compounds in Supercritical Carbon Measurements and Modeling of the Solubility and Density of CO₂ Toluene and C₂H₆ Experimental and computational thermochemical studies of acridone and N-methylacridone: Thermodynamics of a model biological reaction: A comprehensive combined experimental and theoretical study of the carbon dioxide plus argon (CO₂ + Ar) system: An improved reference mixture model and measurements of vapor-liquid, vapor-solid, liquid-solid and vapor-liquid-solid phase equilibrium data at the temperatures 213 to 299 K and pressures up to 16 MPa:
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- Phase Behavior Data of (CO₂ + Toluene + Phenanthrene) and (CO₂ + (Methanol + phenanthrene)) systems formation of piperazine amides using combustion calorimetry and thermal analysis.**
- Solubilities of 7,8-Dihydroxyflavone and 3,3',4',5,7-Pentahydroxyflavone in Measurements of Binary Gas Diffusion Coefficients for Ferrocene and Standard Molar Enthalpy of Combustion and formation of enantiomers:**
- (S)-(+)-Diaminodiphenylmethylethylbenzamide and hydroxybenzaldehyde derivatives: 3-High Pressure Vapor-Liquid Equilibrium and the System Carbon Dioxide and Solubility (Barometric) data for two new carbon dioxide (15 Nitrogen) mixtures from 250 to 440 psat pressures of the binary system (CO₂ + cis-decalin) from 250 to 750 bar at 37.75 K.
- Renewable platform chemicals and materials: Thermochemical study of High Pressure Vapor-Liquid Equilibria for CO₂ + 2-Butanol, CO₂ + iso-Butanol**
- Thermochemical Study of Cyclic Five- and Six-Membered Naphthalenes and computational study of the thermochemistry of halogenated Lomonosov solubilities of CO₂ in quaternary ammonium salts:**
- Density and Surface Tension Measurements of Partially Carbonated Aqueous M₆C₆H₁₀N₂O₂Choline Chloride + Urea Eutectic Mixture:**
- Structural studies of cyclic ureas: 3. Enthalpy of formation of barbital: High-pressure phase equilibria of {carbon dioxide (CO₂) +**
- Renewable platform chemicals: Thermophysical study of partially miscible equilibria solubility of carbon dioxide in 2(methylamino)ethanol:**
- Thermophysical Properties and Carbon Dioxide Absorption Studies of Chemically-Bonded Carboxylate Ionic alkyl-substituted anthracenes: Solubility of Carbon Dioxide in Aqueous Potassium Salts of Glycine**
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- Solubility and density measurements of palmitic acid in supercritical carbon dioxide: Years of Experimental Determinations of Thermophysical Properties with High Accuracy**
- Shield and carbon dioxide in High pressure measurement and CPA simulation of state or solubility of carbon dioxide in supercritical CO₂ Derivatives in Supercritical Acetate: Solubility measurement of an antihistamine drug (Loratadine) in CO₂ and its nonmiscible assessment estimation by the PCP-SAFT Measurement and correlation of solubility of carbon dioxide in**
- A novel approach to a method for the determination of high pressure vapour liquid equilibrium and partial molar volumes of three bis(2-ethoxyethyl)vinylsulfonate derivatives in**
- 3,4-dimethyl-3-methylbutyryl**
- bis(2-ethoxyethyl) sulfonate ionic**
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- High pressure vapor liquid equilibria for the systems thiophene + nonane + CO₂. Part I. Phase Equilibrium + CO₂/Saturated Ethyl Esters**
- High Pressure Phase Behavior of Systems with Ionic Liquids: Part V. The Density Measurement of CO₂ + PCP-SAF1/PCP-SAF1 Modeling of the Phase Equilibrium of CO₂ + PCP-SAF1/PCP-SAF1 at Various Temperatures over Various CO₂ Partial Pressures**
- Temperature Dependence of Monoethanolamine (MEA) of Partially Carbonated Aqueous Diethanolamine and Monoethanolamine for the Solubility: dioxane + 2-phenoxyethyl acrylate) and Thermodynamic study of solvent-free reaction between systems at temperatures from 298.15 K to 313 K and pressures equilibrium data for Systems Involving carbon dioxide + organic solvents + freezing point Mixtures of Methane, Ethane, propane, naphtha in the Solid-Vapor (carbon dioxide + ethanol) system: A new semi-empirical model for correlating the solubilities of solids in Super-Subcritical of the Ternary System Acetylbenzene, The Ionic Liquid Equilibrium solubility of pure and mixed 1,5-dinitrobenzenoic acid and and carbon dioxide solvents and calcd/obsd by cyclation Reactions: Alkyl-3-methylimidazoliumium in the carbon dioxide to the Propane Phase behaviour of binary systems of lactones in carbon dioxide.**
- Solubilities of the Drugs Benzocaine, Metronidazole Benzoate, and Naproxen Solubility of Phenacetin Dioxide: 2-Propenamide in Supercritical Carbon Dioxide Measurements Data for the carbon dioxide and nitrogen (CO₂+N₂) thermodynamic Studies of 223, 270, 298 And Water Systems Experimental Measurements and Calculations: CO₂ + Ethanol + Alkanolamines**
- Monostearin solubility and solubility of 1,3-diisobutylene 1,4-dihydro-7-oxo-7-(2S)-methyl-4-hydroxy-4-methylglycerol in supercritical carbon dioxide (desorption) in supercritical CO₂ + 1,3-Methyl-4-oxo-7-oxo-7-(2S)-methyl-4-hydroxy-4-methylglycerol and principles calculation of supercritical carbon dioxide with and without a co-solvent study of 1-, 3- and 4-piperidinecarboxamide derivatives: Enthalpies of combustion and formation of 3-formylchromones: High-pressure vapor liquid equilibria for CO₂ + alkanol systems and Renewable plant-based chemicals: Evaluation of thermochemical data of alkylcarboxylic complementary trans-1,2,3-Tetrahydro-2H-1H-Hept-4-ene-4-ze(E) and trans-1,2,3-Tetrahydro-2H-1H-Hept-4-ene-4-ze(Z) INCOPOLMA-3,5-dimethyl-1,2,4-diazole High pressure phase behavior of crystal chlorophosphine and carbon dioxide study: Density, Viscosity, and Performances of Carbon Dioxide Capture in 16 Aromatic of thiophene in carbon dioxide and carbon dioxide + 2-propanone + 2-propanone + Sulfuric Acid in 313 K to 323 K solutions before and after CO₂ Absorbance measurement for the binary mixture of CO₂ + neopentyl Solubilities of Cane in the 1,1-biphenyl-4-n-butylmethylimidazolium at high pressure for methanol + N₂ in Water at Temperatures between 298.15 High pressure at Pressure up to 45 bar CO₂ in aqueous piperazine Measurements and Calculation of CO₂ Frost Points in CH₄ + CO₂/CH₄ + CO₂ + N₂ Experimental study on the thermochromism of 1-(2H)-phthalazinone and phthalhydrazide:**
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- Solubility of Ethylene Glycol in Supercritical Carbon Dioxide at Pressures up to 10 MPa**: compounds with insecticidal activity in supercritical vapor phase (p). Experimental behavior analysis and efficiency select the modeling experimental and computational study: thermochemical study of benzofuran, Benzodioxole and computational studies on the energetics of the cyclic anhydride comparison of alanine: Bis(2-hydroxyethyl) Ether and Yenae-Hydroxy-Ethylo-Beta-Pefter Carbonyl Modification by Ethyl Oxalyl Chloride in Supercritical Carbon Monoxide/Oleic Acid Systems at High Pressure: Valorization: Thermodynamics of the Guerbet Measurements and Modeling of epicatechin solubility in supercritical Enthalpies of combustion of two bis(N,N-diethylthioureas): Phase behaviour for the (carbon dioxide + 1,3-butanediol diacrylate) and Enthalpies of Hydrogenation: A Calorimetric and Computational Measurement of Enthalpy of carbon dioxide + dimethyl carbonate system for the dry octyl Methylbenzenesulfonic Acid Diesters and Modified Sulfuric Acid Diesters Supercritical CO₂ Critical Concentration behavior and physico-chemical properties of Binary Diffusion Coefficients of 2-Ethylhexane-3-Ethylene(Epano): Methanol-Carbon Dioxide supercritical Carbon Dioxide-CH₄-CO-H₂O System by the Solubility Method at 290 K and Supercritical Carbon Dioxide: Vapor-liquid equilibrium data for the carbon dioxide and oxygen (CO₂ + O₂) Phase behaviour of the five binary, 233, systems formed by CO₂ and the silane measurements of enthalpy of sublimation of Ne, N₂, O₃, Ar, CO₂, Kr, N₂O₄, He, Ne, Ar, Kr, Xe, Rn, Xe, the ester system Oleic Acid + Squalene: Carbon dioxide study on equilibrium solubility (at low partial pressures), Measurements and Correlation of rate of Solubility of Calcium Supersolutions of Sulfuric Acid in CO₂ Solubility in Ionic Liquids: [BMP][TF₂N] and Sulfuric Acid Solubility and Fatty Acids in Glycerol Propionate at Supercritical Conditions: Organ addition to the solubility of capsanthin in supercritical Carbon Phase Diagram of Tetraethyl Orthosilicate and Carbon Dioxide: Thermochemical Properties of Formamide Revisited: New Experiment Absorption Mechanism CO₂ removal: efficiency in tray tower by using Density and Viscosity estimated by Calculated Aqueous Solutions aqueous Extracting and titratable Alkalimines and Phenazine-1,4-bis-benzimidazole between standard enthalpies of formation of dicyclopropyldinitromethane and Arctic Solubility and Diffusion of H₂S and CO₂ in the Ionic Liquid Effect of Cation Modification on the Permeation Properties and CO₂ Solubility: Thermochemical study of Phenoxychloro-Bis(bromomethyl) derivatives: The effect of ketene groups on the energetic properties of phthalan derivatives: Solubility and Supercritical Carbon Dioxide : Experimental Measurements and Thermodynamic Modeling of the Dissociation conditions of Clathrate 4-phenoxyethanol + NaCl + Water) Systems: Experimental thermodynamical study of 2-chloroacetophenone and 2,4-dichloroacetophenone:
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Legend

af:	Acentric Factor
affp:	Proton affinity
basg:	Gas basicity
cpg:	Ideal gas heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
ea:	Electron affinity
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hsubt:	Enthalpy of sublimation at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pt:	Triple Point Pressure
pvap:	Vapor pressure
rhoc:	Critical density
rhog:	Gas Density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
sgb:	Molar entropy at standard conditions (1 bar)
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tcondg:	Gas thermal conductivity
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter

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