

Benzene, 1,3,5-trimethoxy-

Other names: 1,3,5-Trimethoxybenzene; 1,3,5-Trimethoxybenzene;
O,O,O-1,3,5-trimethylresorcinol; Phloroglucinol trimethyl ether.

InChI: InChI=1S/C9H12O3/c1-10-7-4-8(11-2)6-9(5-7)12-3/h4-6H,1-3H3

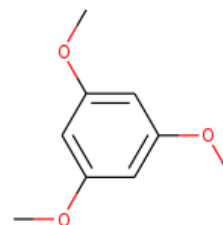
InChI Key: LKUDPHPHKOZXCD-UHFFFAOYSA-N

Formula: C₉H₁₂O₃

SMILES: COc1cc(OC)cc(OC)c1

Molecular Weight: 168.19

CAS: 621-23-8



Physical Properties

Property	Value	Unit	Source
PAff	926.70	kJ/mol	NIST Webbook
BasG	898.20	kJ/mol	NIST Webbook
$\Delta_f G^\circ$	-196.95	kJ/mol	Joback Method
$\Delta_f H^\circ_{\text{gas}}$	-412.16	kJ/mol	Joback Method
$\Delta_{\text{fus}} H^\circ$	15.89	kJ/mol	Joback Method
$\Delta_{\text{sub}} H^\circ$	100.60 ± 1.90	kJ/mol	NIST Webbook
$\Delta_{\text{vap}} H^\circ$	68.20 ± 2.00	kJ/mol	NIST Webbook
IE	7.96	eV	NIST Webbook
$\log P_{\text{oct/wat}}$	1.712		Crippen Method
P_c	2979.54	kPa	Joback Method
T_{boil}	528.70	K	NIST Webbook
T_c	714.26	K	Joback Method
T_{fus}	309.34	K	Joback Method
V_c	0.485	m ³ /kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	285.91	J/mol×K	509.22	Joback Method
η	0.0001357	Paxs	509.22	Joback Method
$\Delta_{sub}H$	116.00 ± 1.90	kJ/mol	375.0	NIST Webbook

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method

NIST Webbook:

[http://webbook.nist.gov/cgi/inchi/InChI=1S/C9H12O3/c1-10-7-4-8\(11-2\)6-9\(5-7\)12-3/h4-6H,1-3H3](http://webbook.nist.gov/cgi/inchi/InChI=1S/C9H12O3/c1-10-7-4-8(11-2)6-9(5-7)12-3/h4-6H,1-3H3)

Crippen Method: <http://pubs.acs.org/doi/abs/10.1021/ci9903071>

Legend

PAff: Proton affinity (kJ/mol).

BasG: Gas basicity (kJ/mol).

$C_{p,gas}$: Ideal gas heat capacity (J/mol×K).

η : Dynamic viscosity (Paxs).

$\Delta_f G^\circ$: Standard Gibbs free energy of formation (kJ/mol).

$\Delta_f H^\circ_{gas}$: Enthalpy of formation at standard conditions (kJ/mol).

$\Delta_{fus} H^\circ$: Enthalpy of fusion at standard conditions (kJ/mol).

$\Delta_{sub} H^\circ$: Enthalpy of sublimation at standard conditions (kJ/mol).

$\Delta_{sub} H$: Enthalpy of sublimation at a given temperature (kJ/mol).

$\Delta_{vap} H^\circ$: Enthalpy of vaporization at standard conditions (kJ/mol).

IE: Ionization energy (eV).

$\log P_{oct/wat}$: Octanol/Water partition coefficient .

P_c : Critical Pressure (kPa).

T_{boil} : Normal Boiling Point Temperature (K).

T_c : Critical Temperature (K).

T_{fus} : Normal melting (fusion) point (K).

V_c : Critical Volume (m³/kg-mol).

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