Acetonitrile

Other names:	Acetonitril
	CH3CN
	Cyanomethane
	Cyanure de methyl
	ETHANENITIRILE
	Ethanenitrile
	Ethanonitrile
	Ethyl nitrile
	Methane, cyano-
	Methanecarbonitrile
	Methyl cyanide
	Methylkyanid
	NA 1648
	NCI-C60822
	Rcra waste number U003
	UN 1648
	USAF EK-488
Inchi:	InChI=1S/C2H3N/c1-2-3/h1H3
InchiKey:	WEVYAHXRMPXWCK-UHFFFAOYSA-N
Formula:	C2H3N
SMILES:	CC#N
Mol. weight [g/mol]:	41.05
CAS:	75-05-8

Physical Properties

Property code	Value	Unit	Source
af	0.3270		KDB
affp	787.40 ± 5.90	kJ/mol	NIST Webbook
affp	779.20	kJ/mol	NIST Webbook
aigt	797.04	K	KDB
basg	748.00	kJ/mol	NIST Webbook
chl	-1270.00	kJ/mol	NIST Webbook
chl	-1247.20 ± 7.20	kJ/mol	NIST Webbook
chl	-1256.33 ± 0.30	kJ/mol	NIST Webbook

cpl	92.36	J/mol×K	Volumes, Heat Capacities, and Compressibilities of the Mixtures of Acetonitrile with N,N-Dimethylacetamide and Propylene Carbonate
cpl	90.93	J/mol×K	Heat capacities of the mixtures of ionic liquids with acetonitrile
dm	3.50	debye	KDB
ea	0.01	eV	NIST Webbook
ea	0.00 ± 0.01	eV	NIST Webbook
ea	0.01	eV	NIST Webbook
fll	4.40	% in Air	KDB
flu	16.00	% in Air	KDB
fpc	278.71	K	KDB
gf	105.70	kJ/mol	KDB
gyrad	1.8210		KDB
hf	87.92	kJ/mol	KDB
hf	74.04 ± 0.37	kJ/mol	NIST Webbook
hf	65.86	kJ/mol	NIST Webbook
hfl	31.40	kJ/mol	NIST Webbook
hfl	40.56 ± 0.40	kJ/mol	NIST Webbook
hfus	2.44	kJ/mol	Joback Method
hvap	33.40	kJ/mol	NIST Webbook
hvap	32.94 ± 0.06	kJ/mol	NIST Webbook
hvap	33.45 ± 0.21	kJ/mol	NIST Webbook
hvap	33.00	kJ/mol	NIST Webbook
hvap	33.00	kJ/mol	NIST Webbook
ie	12.20 ± 0.01	eV	NIST Webbook
ie	12.20 ± 0.01	eV	NIST Webbook
ie	13.14	eV	NIST Webbook
ie	12.21	eV	NIST Webbook
ie	12.46	eV	NIST Webbook
ie	15.11	eV	NIST Webbook
ie	12.19 ± 0.01	eV	NIST Webbook
ie	12.33 ± 0.08	eV	NIST Webbook
ie	12.30 ± 0.25	eV	NIST Webbook
ie	13.11	eV	NIST Webbook
ie	15.12	eV	NIST Webbook
ie	12.20	eV	NIST Webbook
ie	12.38 ± 0.04	eV	NIST Webbook
ie	16.98	eV	NIST Webbook
ie	12.23 ± 0.05	eV	NIST Webbook
ie	12.21 ± 0.00	eV	NIST Webbook
ie	12.20 ± 0.00	eV	NIST Webbook
ie	12.22 ± 0.01	eV	NIST Webbook

ie	12.12	eV	NIST Webbook
ie	12.19 ± 0.01	eV	NIST Webbook
log10ws	0.26		Estimated Solubility Method
log10ws	0.26		Aqueous Solubility Prediction Method
logp	0.530		Crippen Method
mcvol	40.420	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
nfpah	%!d(float64=2)		KDB
nfpas	%!d(float64=1)		KDB
pc	4850.00	kPa	KDB
pc	4833.20 ± 81.06	kPa	NIST Webbook
pc	4830.00 ± 81.06	kPa	NIST Webbook
pc	4934.00 ± 3.00	kPa	NIST Webbook
pc	4934.00 ± 20.00	kPa	NIST Webbook
pc	4830.00 ± 20.00	kPa	NIST Webbook
pc	4833.20 ± 81.06	kPa	NIST Webbook
pc	4934.00 ± 3.00	kPa	NIST Webbook
pc	4830.00 ± 20.00	kPa	NIST Webbook
pc	4890.00 ± 10.00	kPa	NIST Webbook
rinpol	450.00		NIST Webbook
rinpol	453.90		NIST Webbook
rinpol	439.00		NIST Webbook
rinpol	452.92		NIST Webbook
rinpol	452.71		NIST Webbook
rinpol	452.50		NIST Webbook
rinpol	454.52		NIST Webbook
rinpol	452.35		NIST Webbook
rinpol	456.90		NIST Webbook
rinpol	432.00		NIST Webbook
rinpol	455.45		NIST Webbook
rinpol	457.67		NIST Webbook
rinpol	456.69		NIST Webbook
rinpol	455.74		NIST Webbook
rinpol	455.25		NIST Webbook
rinpol	454.45		NIST Webbook
rinpol	453.70		NIST Webbook
rinpol	467.00		NIST Webbook
rinpol	439.00		NIST Webbook
rinpol	455.00		NIST Webbook
rinpol	447.00		NIST Webbook
rinpol	455.00		NIST Webbook
rinpol	464.00		NIST Webbook
rinpol	490.00		NIST Webbook

rinpol	455.00		NIST Webbook
rinpol	443.00		NIST Webbook
rinpol	452.72		NIST Webbook
rinpol	452.00		NIST Webbook
rinpol	443.00		NIST Webbook
rinpol	447.00		NIST Webbook
rinpol	470.00		NIST Webbook
rinpol	456.00		NIST Webbook
rinpol	500.00		NIST Webbook
rinpol	446.00		NIST Webbook
rinpol	439.00		NIST Webbook
rinpol	453.18		NIST Webbook
rinpol	444.00		NIST Webbook
rinpol	460.00		NIST Webbook
rinpol	440.00		NIST Webbook
rinpol	453.32		NIST Webbook
rinpol	425.00		NIST Webbook
rinpol	464.00		NIST Webbook
rinpol	460.00		NIST Webbook
rinpol	452.53		NIST Webbook
rinpol	447.00		NIST Webbook
rinpol	452.90		NIST Webbook
rinpol	442.00		NIST Webbook
ripol	1025.00		NIST Webbook
ripol	1012.00		NIST Webbook
ripol	1026.00		NIST Webbook
ripol	1002.00		NIST Webbook
ripol	1026.00		NIST Webbook
ripol	1030.00		NIST Webbook
ripol	1003.00		NIST Webbook
ripol	1013.00		NIST Webbook
ripol	988.00		NIST Webbook
ripol	1013.00		NIST Webbook
ripol	1002.00		NIST Webbook
ripol	1011.00		NIST Webbook
ripol	1045.00		NIST Webbook
ripol	1012.00		NIST Webbook
ripol	1026.00		NIST Webbook
ripol	1045.00		NIST Webbook
ripol	1045.00		NIST Webbook
ripol	1010.00		NIST Webbook
sl	149.62	J/mol×K	NIST Webbook
tb	355.00 ± 0.50	K	NIST Webbook
tb	354.75 ± 0.50	K	NIST Webbook

tb	354.70 ± 2.00	K	NIST Webbook
tb	354.70	K	NIST Webbook
tb	355.15 ± 1.50	K	NIST Webbook
tb	354.80 ± 0.50	K	NIST Webbook
tb	354.00 ± 2.00	K	NIST Webbook
tb	354.90 ± 0.30	K	NIST Webbook
tb	355.00 ± 2.00	K	NIST Webbook
tb	354.90 ± 0.70	K	NIST Webbook
tb	353.00 ± 2.00	K	NIST Webbook
tb	355.00 ± 2.00	K	NIST Webbook
tb	370.55 ± 0.10	K	NIST Webbook
tb	354.25 ± 0.50	K	NIST Webbook
tb	354.60 ± 0.30	K	NIST Webbook
tb	355.00	K	NIST Webbook
tb	354.25 ± 0.30	K	NIST Webbook
tb	354.80	K	NIST Webbook
tb	354.75 ± 0.20	K	NIST Webbook
tb	354.75	K	Isobaric Vapor Liquid Equilibrium for the Acetonitrile + Water System Containing Different Ionic Liquids at Atmospheric Pressure
tb	354.65	K	Vapor-Liquid Equilibria for the Ternary System Acetonitrile + 1-Propanol + Dimethyl Sulfoxide and the Corresponding Binary Systems at 101.3 kPa
tb	354.74	K	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure
tb	354.70	K	Acetonitrile Dehydration via Extractive Distillation Using Low Transition Temperature Mixtures as Entrainers
tb	354.25	K	Vapor Liquid Equilibria Measurement of (Methanol + Ethanenitrile + Bis(trifluoromethylsulfonyl) Imide)-Based Ionic Liquids at 101.3 kPa
tb	355.00 ± 2.00	K	NIST Webbook
tb	354.72 ± 0.10	K	NIST Webbook
tb	354.68	K	Measurement and correlation of isobaric vapour-liquid equilibrium for the (acetonitrile + water) system containing different ionic liquids at atmospheric pressure

tb	354.65	K	(Vapour + liquid) equilibria in the ternary system (acetonitrile + n-propanol + ethylene glycol) and corresponding binary systems at 101.3 kPa
tb	353.90 ± 1.50	K	NIST Webbook
tb	354.80 ± 0.30	K	NIST Webbook
tb	354.80 ± 0.30	K	NIST Webbook
tb	354.90 ± 0.20	K	NIST Webbook
tb	354.80 ± 0.30	K	NIST Webbook
tb	354.80 ± 0.60	K	NIST Webbook
tb	354.80 ± 0.40	K	NIST Webbook
tb	354.75 ± 0.30	K	NIST Webbook
tb	354.65	K	The isobaric vapor liquid equilibria of ethyl acetate p acetonitrile p bis(trifluoromethylsulfonyl)imide-based ionic liquids at 101.3 kPa
tb	354.95 ± 1.50	K	NIST Webbook
tb	354.75 ± 0.30	K	NIST Webbook
tb	354.69 ± 0.50	K	NIST Webbook
tb	354.80	K	KDB
tb	354.90 ± 0.60	K	NIST Webbook
tb	354.69 ± 0.30	K	NIST Webbook
tb	355.00 ± 1.50	K	NIST Webbook
tb	355.00 ± 1.50	K	NIST Webbook
tb	354.70 ± 2.00	K	NIST Webbook
tb	354.35 ± 0.50	K	NIST Webbook
tb	354.71 ± 0.20	K	NIST Webbook
tb	354.70	K	NIST Webbook
tb	351.00 ± 3.00	K	NIST Webbook
tb	355.00 ± 0.50	K	NIST Webbook
tb	354.80 ± 0.30	K	NIST Webbook
tb	354.65 ± 0.50	K	NIST Webbook
tb	354.75 ± 0.20	K	NIST Webbook
tb	354.45 ± 0.50	K	NIST Webbook
tb	354.75	K	NIST Webbook
tb	354.80 ± 0.30	K	NIST Webbook
tb	354.75 ± 0.20	K	NIST Webbook
tb	354.80 ± 0.50	K	NIST Webbook
tb	354.70 ± 0.40	K	NIST Webbook
tb	354.80 ± 0.40	K	NIST Webbook
tc	545.50 ± 0.20	K	NIST Webbook
tc	545.50	K	KDB
tc	543.15 ± 1.00	K	NIST Webbook
tc	543.45 ± 1.00	K	NIST Webbook

tc	543.20 ± 2.00	K	NIST Webbook
tc	544.99 ± 0.20	K	NIST Webbook
tc	544.99 ± 0.20	K	NIST Webbook
tc	544.99 ± 0.20	K	NIST Webbook
tc	545.50 ± 0.30	K	NIST Webbook
tc	545.50 ± 0.30	K	NIST Webbook
tf	227.45	K	Aqueous Solubility Prediction Method
tf	229.07	K	Efficient determination of crystallisation and melting points at low cooling and heating rates with novel computer controlled equipment
tf	230.42	K	Experimental (Solid + Liquid) and (Liquid + Liquid) Equilibria and Excess Molar Volume of Alkanol + Acetonitrile, Propanenitrile, and Butanenitrile Mixtures
tf	229.32	K	KDB
tt	229.32 ± 0.02	K	NIST Webbook
vc	0.173	m3/kmol	KDB
zc	0.1849930		KDB
zra	0.20		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	62.33	J/mol×K	411.95	Joback Method
cpg	72.87	J/mol×K	541.37	Joback Method
cpg	70.38	J/mol×K	509.01	Joback Method
cpg	67.79	J/mol×K	476.66	Joback Method
cpg	56.48	J/mol×K	347.24	Joback Method
cpg	59.46	J/mol×K	379.59	Joback Method
cpg	65.11	J/mol×K	444.30	Joback Method
cpl	91.70	J/mol×K	298.15	NIST Webbook
cpl	91.70	J/mol×K	298.15	NIST Webbook
cpl	81.80	J/mol×K	303.15	NIST Webbook
cpl	77.40	J/mol×K	298.15	NIST Webbook
cpl	91.69	J/mol×K	298.15	NIST Webbook

cpl	92.82	J/mol×K	313.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	92.40	J/mol×K	308.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	92.21	J/mol×K	303.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	91.69	J/mol×K	298.15	Volumetric Properties, Viscosities, and Isobaric Heat Capacities of Imidazolium Octanoate Protic Ionic Liquid in Molecular Solvents
cpl	82.50	J/mol×K	297.00	NIST Webbook
cpl	91.46	J/mol×K	298.15	NIST Webbook
cpl	91.70	J/mol×K	298.15	NIST Webbook
dvisc	0.0003020	Pa×s	323.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003431	Pa×s	298.15	Volumetric properties of ionic liquid 1,3-dimethylimidazolium methyl sulfate + molecular solvents at T = (298.15 - 328.15) K

dvisc	0.0003000	Pa×s	323.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0002800	Pa×s	333.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0002600	Pa×s	343.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003443	Pa×s	298.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques
dvisc	0.0003124	Pa×s	308.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques

dvisc	0.0002890	Pa×s	318.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques
dvisc	0.0002891	Pa×s	318.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques
dvisc	0.0003125	Pa×s	308.15	Electrical Conductances of 1-Butyl-3-propylimidazolium Bromide and 1-Butyl-3-propylbenzimidazolium Bromide in Water, Methanol, and Acetonitrile at (308, 313, and 318) K at 0.1 MPa
dvisc	0.0003042	Pa×s	313.15	Electrical Conductances of 1-Butyl-3-propylimidazolium Bromide and 1-Butyl-3-propylbenzimidazolium Bromide in Water, Methanol, and Acetonitrile at (308, 313, and 318) K at 0.1 MPa
dvisc	0.0002903	Pa×s	318.15	Electrical Conductances of 1-Butyl-3-propylimidazolium Bromide and 1-Butyl-3-propylbenzimidazolium Bromide in Water, Methanol, and Acetonitrile at (308, 313, and 318) K at 0.1 MPa

dvisc	0.0003700	Pa×s	298.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0002918	Pa×s	318.15	Volumetric properties of ionic liquid 1,3-dimethylimidazolium methyl sulfate + molecular solvents at T = (298.15 - 328.15) K
dvisc	0.0003540	Pa×s	303.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003270	Pa×s	313.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0002680	Pa×s	328.15	Volumetric properties of ionic liquid 1,3-dimethylimidazolium methyl sulfate + molecular solvents at T = (298.15 - 328.15) K
dvisc	0.0002810	Pa×s	333.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile

dvisc	0.0002620	Pa×s	343.15	Properties of Pure 1-Butyl-2,3-dimethylimidazolium Tetrafluoroborate Ionic Liquid and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003440	Pa×s	298.15	Viscosity, Density, and Speed of Sound for the Binary Mixtures of Formamide with 2-Methoxyethanol, Acetophenone, Acetonitrile, 1,2-Dimethoxyethane, and Dimethylsulfoxide at Different Temperatures
dvisc	0.0003130	Pa×s	308.15	Viscosity, Density, and Speed of Sound for the Binary Mixtures of Formamide with 2-Methoxyethanol, Acetophenone, Acetonitrile, 1,2-Dimethoxyethane, and Dimethylsulfoxide at Different Temperatures
dvisc	0.0002890	Pa×s	318.15	Viscosity, Density, and Speed of Sound for the Binary Mixtures of Formamide with 2-Methoxyethanol, Acetophenone, Acetonitrile, 1,2-Dimethoxyethane, and Dimethylsulfoxide at Different Temperatures
dvisc	0.0003417	Pa×s	298.15	Density and Viscosity of Anhydrous Mixtures of Dimethylsulfoxide with Acetonitrile in the Range (298.15 to 318.15) K

dvisc	0.0003280	Pa×s	303.15	Density and Viscosity of Anhydrous Mixtures of Dimethylsulfoxide with Acetonitrile in the Range (298.15 to 318.15) K
dvisc	0.0003129	Pa×s	308.15	Density and Viscosity of Anhydrous Mixtures of Dimethylsulfoxide with Acetonitrile in the Range (298.15 to 318.15) K
dvisc	0.0003009	Pa×s	313.15	Density and Viscosity of Anhydrous Mixtures of Dimethylsulfoxide with Acetonitrile in the Range (298.15 to 318.15) K
dvisc	0.0002899	Pa×s	318.15	Density and Viscosity of Anhydrous Mixtures of Dimethylsulfoxide with Acetonitrile in the Range (298.15 to 318.15) K
dvisc	0.0003081	Pa×s	293.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures
dvisc	0.0002857	Pa×s	298.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures
dvisc	0.0002642	Pa×s	303.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures

dvisc	0.0002444	Pa×s	308.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures
dvisc	0.0002374	Pa×s	310.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures
dvisc	0.0002267	Pa×s	313.15	Thermodynamic Study of Phenyl Salicylate Solutions in Aprotic Solvents at Different Temperatures
dvisc	0.0003417	Pa×s	298.15	The (water + acetonitrile) mixture revisited: A new approach for calculating partial molar volumes
dvisc	0.0003280	Pa×s	303.15	The (water + acetonitrile) mixture revisited: A new approach for calculating partial molar volumes
dvisc	0.0003171	Pa×s	308.15	Volumetric properties of ionic liquid 1,3-dimethylimidazolium methyl sulfate + molecular solvents at T = (298.15 - 328.15) K
dvisc	0.0003129	Pa×s	308.15	The (water + acetonitrile) mixture revisited: A new approach for calculating partial molar volumes
dvisc	0.0003009	Pa×s	313.15	The (water + acetonitrile) mixture revisited: A new approach for calculating partial molar volumes

dvisc	0.0002899	Pa×s	318.15	The (water + acetonitrile) mixture revisited: A new approach for calculating partial molar volumes
dvisc	0.0003860	Pa×s	293.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0003700	Pa×s	298.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0003540	Pa×s	303.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0003270	Pa×s	313.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0003020	Pa×s	323.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0002810	Pa×s	333.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents
dvisc	0.0002620	Pa×s	343.15	Viscosity of binary mixtures of 1-ethyl-3-methylimidazolium tetrafluoroborate ionic liquid with four organic solvents

dvisc	0.0003696	Pa×s	298.15	Molecular interaction studies and theoretical estimation of ultrasonic speeds using scaled particle theory in binary mixtures of toluene with homologous nitriles at different temperatures
dvisc	0.0003554	Pa×s	303.15	Molecular interaction studies and theoretical estimation of ultrasonic speeds using scaled particle theory in binary mixtures of toluene with homologous nitriles at different temperatures
dvisc	0.0003414	Pa×s	308.15	Molecular interaction studies and theoretical estimation of ultrasonic speeds using scaled particle theory in binary mixtures of toluene with homologous nitriles at different temperatures
dvisc	0.0003900	Pa×s	293.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003700	Pa×s	298.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile

dvisc	0.0003300	Pa×s	313.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
dvisc	0.0003500	Pa×s	303.15	Viscosities of 1-Hexyl-3-methylimidazolium Tetrafluoroborate and Its Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile
hfust	8.17	kJ/mol	229.30	NIST Webbook
hvapt	31.38	kJ/mol	352.80	KDB
hvapt	33.23	kJ/mol	298.15	NIST Webbook
hvapt	29.75	kJ/mol	354.70	NIST Webbook
hvapt	33.30	kJ/mol	334.50	NIST Webbook
hvapt	33.80	kJ/mol	325.00	NIST Webbook
hvapt	33.90	kJ/mol	290.00	NIST Webbook
hvapt	34.80	kJ/mol	321.00	NIST Webbook
hvapt	34.20	kJ/mol	298.00	NIST Webbook
pvap	122.63	kPa	360.94	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	36.39	kPa	325.02	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	31.31	kPa	321.12	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	24.97	kPa	315.43	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	19.64	kPa	309.64	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	15.27	kPa	303.81	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry

pvap	13.37	kPa	300.82	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	10.60	kPa	295.78	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	9.18	kPa	292.75	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	8.30	kPa	290.65	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	7.17	kPa	287.67	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	6.38	kPa	285.38	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	5.25	kPa	281.57	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	4.49	kPa	278.62	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	4.32	kPa	277.92	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	96.34	kPa	353.10	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	82.37	kPa	348.25	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol

pvap	69.62	kPa	343.18	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	58.58	kPa	338.15	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	54.01	kPa	335.84	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	49.06	kPa	333.15	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	45.62	kPa	331.15	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	41.38	kPa	328.50	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	37.90	kPa	326.16	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	33.80	kPa	323.16	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	30.64	kPa	320.65	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol

pvap	27.78	kPa	318.18	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	25.72	kPa	316.27	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	22.65	kPa	313.15	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	21.13	kPa	311.49	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	17.52	kPa	307.09	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	14.15	kPa	302.25	Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 1-Pentanamine + 1-Methoxy-2-propanol
pvap	131.21	kPa	363.15	Vapor-Liquid Equilibria for Four Binary Systems at 363.15 K: N-Methylformamide + Hexane, + Benzene, + Chlorobenzene, and + Acetonitrile
pvap	101.30	kPa	354.74	Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + 1-Propanol + Ionic Liquids at an Atmospheric Pressure

pvap	101.32	kPa	354.70	Acetonitrile Dehydration via Extractive Distillation Using Low Transition Temperature Mixtures as Entrainers
pvap	10.00	kPa	295.40	Phase Behavior of Binary Mixtures Containing Succinic Acid or Its Esters
pvap	10.00	kPa	295.25	Phase Behavior of Binary Mixtures Containing Succinic Acid or Its Esters
pvap	101.32	kPa	354.25	Vapor Liquid Equilibria Measurement of (Methanol + Ethanenitrile + Bis(trifluoromethylsulfonyl) Imide)-Based Ionic Liquids at 101.3 kPa
pvap	37.00	kPa	325.58	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	35.65	kPa	324.59	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	34.49	kPa	323.64	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	33.19	kPa	322.64	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	32.94	kPa	322.56	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	32.95	kPa	322.54	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	31.73	kPa	321.65	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	30.52	kPa	320.67	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	29.39	kPa	319.69	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	28.27	kPa	318.66	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	27.21	kPa	317.72	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	26.11	kPa	316.63	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	125.13	kPa	361.61	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	25.10	kPa	315.71	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	24.17	kPa	314.79	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	23.42	kPa	313.77	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	22.50	kPa	312.81	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	21.55	kPa	311.88	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	21.38	kPa	311.79	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	20.49	kPa	310.71	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	19.85	kPa	309.92	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	18.22	kPa	307.95	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	17.53	kPa	306.90	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	16.96	kPa	306.04	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	16.10	kPa	305.01	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	15.57	kPa	304.16	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	14.18	kPa	302.14	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	13.61	kPa	301.19	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	12.61	kPa	299.42	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	12.17	kPa	298.67	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	12.05	kPa	298.30	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	11.60	kPa	297.48	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	11.02	kPa	296.44	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	10.77	kPa	295.94	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	10.43	kPa	295.33	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	10.03	kPa	294.40	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	9.82	kPa	293.92	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	9.41	kPa	293.21	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	9.16	kPa	292.48	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	8.60	kPa	291.36	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	8.30	kPa	290.60	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	8.03	kPa	290.09	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	7.80	kPa	289.39	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	7.59	kPa	288.83	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	7.56	kPa	288.80	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	7.50	kPa	288.72	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	6.75	kPa	286.61	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	6.59	kPa	286.17	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	4174.61	kPa	535.03	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	6.10	kPa	284.56	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	5.92	kPa	283.81	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	5.84	kPa	283.56	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	5.64	kPa	283.17	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	5.34	kPa	281.91	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	4.86	kPa	280.27	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	4.73	kPa	279.81	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	4.51	kPa	278.98	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	4.54	kPa	278.88	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	4.44	kPa	278.72	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	96.17	kPa	353.13	Experimental determination and prediction of gas solubility data for oxygen in acetonitrile
pvap	33.71	kPa	323.12	Experimental determination and prediction of gas solubility data for oxygen in acetonitrile
pvap	16.47	kPa	305.51	Experimental determination and prediction of gas solubility data for oxygen in acetonitrile
pvap	101.30	kPa	354.65	The isobaric vapor liquid equilibria of ethyl acetate p acetonitrile p bis(trifluoromethylsulfonyl)imide-based ionic liquids at 101.3 kPa
pvap	155.33	kPa	368.97	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	168.12	kPa	371.74	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	175.04	kPa	373.17	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry

pvap	204.59	kPa	378.81	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1923.88	kPa	484.77	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	45.91	kPa	331.27	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	55.17	kPa	336.41	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	65.07	kPa	341.18	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	71.24	kPa	343.94	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	49.76	kPa	333.57	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile

pvap	18.59	kPa	308.32	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	27.19	kPa	317.53	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	37.80	kPa	326.03	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	43.71	kPa	329.94	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	49.76	kPa	333.51	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile

pvap	55.46	kPa	336.56	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	60.76	kPa	339.20	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	64.81	kPa	341.07	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	71.24	kPa	343.86	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile

pvap	75.40	kPa	345.58	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	80.49	kPa	347.55	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	86.34	kPa	349.69	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	92.16	kPa	351.76	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile
pvap	99.40	kPa	354.12	Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Acetonitrile or Propanenitrile

pvap	101.30	kPa	354.75	Isobaric Vapor Liquid Equilibrium for the Acetonitrile + Water System Containing Different Ionic Liquids at Atmospheric Pressure
pvap	75.44	kPa	345.57	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1096.90	kPa	453.15	Vapor-Liquid Equilibria on Seven Binary Systems: Ethylene Oxide + 2-Methylpropane; Acetophenone + Phenol; cis-1,3-Dichloropropene + 1,2-Dichloropropane; 1,5-Hexadiene + Allyl Chloride; Isopropyl Acetate + Acetonitrile; Vinyl Chloride + Methyl Chloride; and 1,4-Butanediol + c-Butyrolactone
pvap	33.86	kPa	323.15	Vapor Pressures for the Acetonitrile + Tetrabutylammonium Bromide, Water + Tetrabutylammonium Bromide, and Acetonitrile + Water + Tetrabutylammonium Bromide Systems

pvap	297.37	kPa	393.15	Vapor-Liquid Equilibria on Seven Binary Systems: Ethylene Oxide + 2-Methylpropane; Acetophenone + Phenol; cis-1,3-Dichloropropene + 1,2-Dichloropropane; 1,5-Hexadiene + Allyl Chloride; Isopropyl Acetate + Acetonitrile; Vinyl Chloride + Methyl Chloride; and 1,4-Butanediol + c-Butyrolactone
pvap	85.31	kPa	349.33	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	95.59	kPa	352.88	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	100.75	kPa	354.55	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	101.12	kPa	354.67	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	110.61	kPa	357.56	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	235.79	kPa	384.11	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	306.28	kPa	394.29	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	399.50	kPa	405.24	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry

pvap	502.67	kPa	415.21	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	605.60	kPa	423.68	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	708.99	kPa	431.12	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	804.86	kPa	437.30	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	910.82	kPa	443.50	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1025.47	kPa	449.60	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1151.97	kPa	455.74	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1290.20	kPa	461.87	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1450.28	kPa	468.37	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1602.63	kPa	474.05	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	1749.84	kPa	479.15	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	2110.77	kPa	490.37	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry

pvap	2303.51	kPa	495.75	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	2523.66	kPa	501.48	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	2747.95	kPa	506.92	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	2999.22	kPa	512.61	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	3254.08	kPa	518.01	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	3512.89	kPa	523.15	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	3760.37	kPa	527.79	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	4001.46	kPa	532.08	Vapor Pressures of Acetonitrile Determined by Comparative Ebulliometry
pvap	6.47	kPa	285.70	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	49.16	kPa	333.15	Vapor-Liquid Equilibria on Seven Binary Systems: Ethylene Oxide + 2-Methylpropane; Acetophenone + Phenol; cis-1,3-Dichloropropene + 1,2-Dichloropropane; 1,5-Hexadiene + Allyl Chloride; Isopropyl Acetate + Acetonitrile; Vinyl Chloride + Methyl Chloride; and 1,4-Butanediol + c-Butyrolactone
pvap	25.15	kPa	315.79	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
rfi	1.33960		303.15	Liquid liquid equilibria measurements of ternary systems (acetonitrile + a carboxylic acid + dodecane) at 303.15 K
rfi	1.34160		298.15	Excess Gibbs free energies of the binary mixtures of acetonitrile with butanols at 94.83 kPa
rfi	1.34160		298.15	(Vapor + liquid) equilibrium of the binary mixtures formed by acetonitrile with selected compounds at 95.5 kPa

rfi	1.34410	293.15	Experimental study and ERAS modeling of the excess molar enthalpy of (acetonitrile + 1-heptanol or 1-octanol) mixtures at (298.15, 313.15, and 323.15) K and atmospheric pressure	
rfi	1.35900	298.15	Volumetric properties, viscosity and refractive index of the protic ionic liquid, pyrrolidinium octanoate, in molecular solvents	
rfi	1.34390	293.15	Experimental solubility for betulin and estrone in various solvents within the temperature range T = (293.2 to 328.2) K	
rfi	1.34420	293.15	ERAS modeling of the excess molar enthalpies of binary liquid mixtures of 1-pentanol and 1-hexanol with acetonitrile at atmospheric pressure and 288, 298, 313 and 323K	
rfi	1.34110	298.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory	

rfi	1.33900	303.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rfi	1.33710	308.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rfi	1.34180	298.15	Physics and Chemistry of Lithium Halides in 1,3-Dioxolane and Its Binary Mixtures with Acetonitrile probed by Conductometric, Volumetric, Viscometric, Refractometric and Acoustic Study
rfi	1.34180	298.15	Vapor-Liquid Equilibrium Data for the Binary Methyl Esters (Butyrate, Pentanoate, and Hexanoate) (1) + Acetonitrile (2) Systems at 93.32 kPa
rfi	1.34300	293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents

rfi	1.34300	293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.34250	293.15	Isobaric Vapor Liquid Equilibrium for the Extractive Distillation of Acetonitrile + Water Mixtures Using Dimethyl Sulfoxide at 101.3 kPa
rfi	1.34160	298.15	Acoustic and Volumetric Properties of Binary Mixtures of Ionic Liquid 1-Butyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide with Acetonitrile and Tetrahydrofuran
rfi	1.34386	293.15	Ternary Liquid Liquid Equilibrium Data for the Water + Acetonitrile + {Butan-1-ol or 2-Methylpropan-1-ol} Systems at (303.2, 323.2, 343.2) K and 1 atm
rfi	1.34190	298.15	Viscosity, Density, Speed of Sound, and Refractive Index of Binary Mixtures of Organic Solvent + Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate at 298.15 K
rfi	1.34410	295.10	Ternary Liquid-Liquid Equilibria of Acetonitrile and Water with Heptanoic Acid and Nonanol at 323.15 K and 1 atm
rfi	1.34139	298.15	Isothermal Vapor-Liquid Equilibrium of Binary Mixtures Containing 1-Chlorobutane, Ethanol, or Acetonitrile

rfi	1.34390		293.15	A novel static analytical apparatus for phase equilibrium measurements
rfi	1.34360		293.15	Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory
rhol	760.16	kg/m3	313.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	776.82	kg/m3	298.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K
rhol	782.00	kg/m3	293.00	KDB
rhol	776.85	kg/m3	298.15	Compressibility Studies of Some Copper(I), Silver(I), and Tetrabutylammonium Salts in Acetonitrile + Adiponitrile Binary Mixtures

rhol	754.73	kg/m3	318.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures
rhol	760.25	kg/m3	313.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures
rhol	765.75	kg/m3	308.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures
rhol	771.21	kg/m3	303.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures
rhol	776.64	kg/m3	298.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures
rhol	782.04	kg/m3	293.15	Volumetric Properties of Binary Mixtures of 1-Butyl-3-methylimidazolium Chloride + Water or Hydrophilic Solvents at Different Temperatures

rhol	726.94	kg/m3	343.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	738.19	kg/m3	333.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	749.73	kg/m3	323.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	760.68	kg/m3	313.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	765.97	kg/m3	308.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K

rhol	771.44	kg/m3	303.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	776.93	kg/m3	298.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	782.30	kg/m3	293.15	Densities of Ionic Liquids, 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate, with Benzene, Acetonitrile, and 1-Propanol at T = (293.15 to 343.15) K
rhol	755.34	kg/m3	313.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K
rhol	763.20	kg/m3	308.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K

rhol	752.52	kg/m3	313.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K	
rhol	761.40	kg/m3	308.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K	
rhol	752.55	kg/m3	313.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K	
rhol	760.00	kg/m3	308.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K	
rhol	768.14	kg/m3	303.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K	

rhol	755.36	kg/m3	313.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K
rhol	761.45	kg/m3	308.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K
rhol	771.43	kg/m3	303.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K
rhol	776.65	kg/m3	298.15	Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K
rhol	754.85	kg/m3	318.15	Density, Viscosity, Speed of Sound, and Refractive Index of a Ternary Solution of Aspirin, 1-Butyl-3-methylimidazolium Bromide, and Acetonitrile at Different Temperatures T = (288.15 to 318.15) K

rhol	765.81	kg/m3	308.15	Density, Viscosity, Speed of Sound, and Refractive Index of a Ternary Solution of Aspirin, 1-Butyl-3-methylimidazolium Bromide, and Acetonitrile at Different Temperatures T = (288.15 to 318.15) K
rhol	776.62	kg/m3	298.15	Density, Viscosity, Speed of Sound, and Refractive Index of a Ternary Solution of Aspirin, 1-Butyl-3-methylimidazolium Bromide, and Acetonitrile at Different Temperatures T = (288.15 to 318.15) K
rhol	787.39	kg/m3	288.15	Density, Viscosity, Speed of Sound, and Refractive Index of a Ternary Solution of Aspirin, 1-Butyl-3-methylimidazolium Bromide, and Acetonitrile at Different Temperatures T = (288.15 to 318.15) K
rhol	726.43	kg/m3	343.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents
rhol	737.86	kg/m3	333.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents

rhol	749.11	kg/m3	323.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents
rhol	760.18	kg/m3	313.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents
rhol	771.13	kg/m3	303.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents
rhol	771.10	kg/m3	303.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	776.55	kg/m3	298.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents
rhol	781.95	kg/m3	293.15	Volumetric Properties of Binary Mixtures of Two 1-Alkyl-3-Methylimidazolium Tetrafluoroborate Ionic Liquids with Molecular Solvents

rhol	776.62	kg/m3	298.15	Calorimetric Study of Nitrile Group-Solvent Interactions and Comparison with Dispersive Quasi-Chemical (DISQUAC) Predictions
rhol	760.20	kg/m3	313.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	765.68	kg/m3	308.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	771.12	kg/m3	303.13	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	776.52	kg/m3	298.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	781.89	kg/m3	293.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	749.07	kg/m3	323.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile

rhol	749.07	kg/m3	323.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	749.11	kg/m3	323.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	749.07	kg/m3	323.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	760.16	kg/m3	313.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	760.15	kg/m3	313.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	

rhol	760.18	kg/m3	313.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	776.53	kg/m3	298.15	Volumetric Properties of the Ionic Liquid, 1-Butyl-3-methylimidazolium Tetrafluoroborate, in Organic Solvents at T = 298.15 K
rhol	765.65	kg/m3	308.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	765.66	kg/m3	308.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	765.64	kg/m3	308.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile

rhol	771.11	kg/m3	303.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	771.11	kg/m3	303.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	771.10	kg/m3	303.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	776.53	kg/m3	298.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	776.53	kg/m3	298.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	

rhol	776.53	kg/m3	298.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	776.53	kg/m3	298.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	781.93	kg/m3	293.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	781.93	kg/m3	293.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	
rhol	781.93	kg/m3	293.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile	

rhol	781.93	kg/m3	293.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	787.31	kg/m3	288.15	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	787.31	kg/m3	288.13	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	787.31	kg/m3	288.13	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile
rhol	787.31	kg/m3	288.13	Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile

rhol	749.25	kg/m3	323.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide
rhol	754.82	kg/m3	318.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide
rhol	760.34	kg/m3	313.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide
rhol	765.83	kg/m3	308.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide
rhol	771.29	kg/m3	303.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide

rhol	776.71	kg/m3	298.15	Effect of Solvents and Temperature on Interactions in Binary and Ternary Mixtures of 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate with Acetonitrile or/and N,N-Dimethylformamide
rhol	785.71	kg/m3	298.15	Liquid-Liquid Equilibrium for Ternary Systems, Water + 5-Hydroxymethylfurfural + (1-Butanol, Isobutanol, Methyl Isobutyl Ketone), at 313.15, 323.15, and 333.15 K
rhol	749.44	kg/m3	323.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	755.14	kg/m3	318.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	760.36	kg/m3	313.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile

rhol	765.83	kg/m3	308.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	771.12	kg/m3	303.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	776.52	kg/m3	298.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	781.89	kg/m3	293.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile
rhol	787.58	kg/m3	288.15	The Thermodynamic and Excess Properties of Trialkyl-Substituted Imidazolium-Based Ionic Liquids with Thiocyanate and Its Binary Systems with Acetonitrile

rhol	749.16	kg/m3	323.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	754.72	kg/m3	318.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	760.25	kg/m3	313.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	765.73	kg/m3	308.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile

rhol	771.18	kg/m3	303.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	776.61	kg/m3	298.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	781.96	kg/m3	293.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile
rhol	787.33	kg/m3	288.15	Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile

rhol	749.25	kg/m3	323.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa
rhol	754.82	kg/m3	318.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa
rhol	760.34	kg/m3	313.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa
rhol	765.83	kg/m3	308.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa

rhol	771.29	kg/m3	303.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa
rhol	776.71	kg/m3	298.15	Excess Molar Volumes, Excess Molar Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters for 1-Ethyl-3-methyl-imidazolium Trifluoro-methanesulfonate + Dimethyl Sulfoxide and/or Acetonitrile at T = 298.15 to 323.15 K and P = 0.1 MPa
rhol	748.99	kg/m3	323.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model
rhol	754.56	kg/m3	318.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model
rhol	760.10	kg/m3	313.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model

rhol	765.59	kg/m3	308.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model
rhol	771.06	kg/m3	303.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model
rhol	776.51	kg/m3	298.15	Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model
rhol	776.68	kg/m3	298.15	Exploration of Solvation Consequence of Ionic Liquid [Bu4PCH3SO3] in Various Solvent Systems by Conductance and FTIR Study
rhol	765.77	kg/m3	308.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	773.30	kg/m3	303.15	Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Molecular Compression Data

rhol	776.66	kg/m3	298.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	782.06	kg/m3	293.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	787.44	kg/m3	288.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	792.79	kg/m3	283.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	798.12	kg/m3	278.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	755.59	kg/m3	318.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K
rhol	760.90	kg/m3	313.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K

rhol	766.21	kg/m3	308.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K
rhol	771.52	kg/m3	303.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K
rhol	765.98	kg/m3	308.10	Synthesis of 1,3-Dimethylimidazolium Chloride and Volumetric Property Investigations of Its Aqueous Solution
rhol	782.12	kg/m3	293.15	Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of acetonitrile with some alkyl methacrylates at temperatures from 293.15 K to 318.15 K
rhol	760.00	kg/m3	313.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures

rhol	771.00	kg/m3	303.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	782.00	kg/m3	293.15	Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures
rhol	771.00	kg/m3	303.15	Thermodynamic and transport properties of acetonitrile + alkanediols liquid mixtures at different temperatures, experimental measurements and modeling
rhol	776.00	kg/m3	298.15	Thermodynamic and transport properties of acetonitrile + alkanediols liquid mixtures at different temperatures, experimental measurements and modeling
rhol	782.00	kg/m3	293.15	Thermodynamic and transport properties of acetonitrile + alkanediols liquid mixtures at different temperatures, experimental measurements and modeling

rhol	755.59	kg/m3	318.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K
rhol	760.90	kg/m3	313.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K
rhol	766.21	kg/m3	308.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K
rhol	771.52	kg/m3	303.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K
rhol	776.82	kg/m3	298.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K
rhol	782.12	kg/m3	293.15	Densities and volumetric properties of (acetonitrile + alkyl acrylate monomer) binary mixtures at temperatures from 293.15 K to 318.15 K

rhol	782.10	kg/m3	293.15	Measurement and correlation of the vapor-liquid equilibrium for methanol + acetonitrile + imidazolium-based ionic liquids at 101.3 kPa
rhol	777.00	kg/m3	298.15	Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K
rhol	777.10	kg/m3	298.15	Solubility and solution thermodynamics of thymol in six pure organic solvents
rhol	765.75	kg/m3	308.15	Volumetric properties of binary mixtures of (acetonitrile + amines) at several temperatures with application of the ERAS model
rhol	771.20	kg/m3	303.15	Volumetric properties of binary mixtures of (acetonitrile + amines) at several temperatures with application of the ERAS model
rhol	776.61	kg/m3	298.15	Volumetric properties of binary mixtures of (acetonitrile + amines) at several temperatures with application of the ERAS model

rhol	782.01	kg/m3	293.15	Volumetric properties of binary mixtures of (acetonitrile + amines) at several temperatures with application of the ERAS model
rhol	776.30	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents
rhol	776.62	kg/m3	298.15	Ionic molar volumes in methanol mixtures with acetonitrile, N,N-dimethylformamide and propylene carbonate at T = 298.15 K
rhol	759.55	kg/m3	313.15	Solution thermodynamics of iron(III)-N,N'-ethylene-bis(salicylideneiminato)-chloride in binary mixtures of N,N-dimethylformamide and acetonitrile at T = (298.15, 303.15, 308.15 and 313.15) K
rhol	765.26	kg/m3	308.15	Solution thermodynamics of iron(III)-N,N'-ethylene-bis(salicylideneiminato)-chloride in binary mixtures of N,N-dimethylformamide and acetonitrile at T = (298.15, 303.15, 308.15 and 313.15) K
rhol	771.56	kg/m3	303.15	Solution thermodynamics of iron(III)-N,N'-ethylene-bis(salicylideneiminato)-chloride in binary mixtures of N,N-dimethylformamide and acetonitrile at T = (298.15, 303.15, 308.15 and 313.15) K

rhol	776.47	kg/m3	298.15	Solution thermodynamics of iron(III)-N,N'-ethylene-bis(salicylideneiminato)-chloride in binary mixtures of N,N-dimethylformamide and acetonitrile at T = (298.15, 303.15, 308.15 and 313.15) K
rhol	776.74	kg/m3	298.15	Probing subsistence of ion-pair and triple-ion of an ionic salt in liquid environments by means of conductometric contrivance
rhol	771.19	kg/m3	303.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory
rhol	776.62	kg/m3	298.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory
rhol	782.01	kg/m3	293.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory
rhol	787.38	kg/m3	288.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory

rhol	776.60	kg/m3	298.15	Density and speed of sound of lithium bromide with organic solvents: Measurement and correlation
rhol	755.70	kg/m3	318.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K
rhol	761.00	kg/m3	313.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K
rhol	766.30	kg/m3	308.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K
rhol	771.50	kg/m3	303.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K
rhol	776.80	kg/m3	298.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K
rhol	782.10	kg/m3	293.15	Densities and volumetric properties of (acetonitrile + an amide) binary mixtures at temperatures between 293.15 K and 318.15 K

rhol	776.53	kg/m3	298.15	Volumetric and compressibility behaviour of ionic liquid, 1-n-butyl-3-methylimidazolium hexafluorophosphate and tetrabutylammonium hexafluorophosphate in organic solvents at T = 298.15 K
rhol	776.69	kg/m3	298.15	Standard partial molar volumes of some electrolytes in ethylene carbonate based mixtures
rhol	782.10	kg/m3	293.15	Effect of imidazolium-based ionic liquid on vapor-liquid equilibria of 2-propanol + acetonitrile binary system at 101.3 kPa
rhol	776.64	kg/m3	298.15	Liquid liquid equilibria in the ternary systems {hexadecane + BTX aromatics + 2-methoxyethanol or acetonitrile} at 298.15 K
rhol	776.55	kg/m3	298.15	Liquid-liquid equilibrium data for ternary mixtures composed of n-hexane, benzene and acetonitrile at (298.15, 308.15, and 318.15) K
rhol	783.38	kg/m3	293.15	Isobaric vapor liquid equilibrium for binary system of methanol and acetonitrile
rhol	786.50	kg/m3	293.15	Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Molecular Compression Data

rhol	760.50	kg/m3	313.15	Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Molecular Compression Data
rhol	766.50	kg/m3	308.15	Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Molecular Compression Data
rhol	771.23	kg/m3	303.15	Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butylresorcin[4]arene in non-aqueous solvents
rhol	781.10	kg/m3	298.15	Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Molecular Compression Data
rhol	776.50	kg/m3	298.15	Bubble point measurements of binary mixtures formed by 1-hexanol with selected nitro-compounds and substituted benzenes at 95.6 kPa

speedsl	1278.62	m/s	298.15	Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile
speedsl	1299.01	m/s	293.15	Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bromide in Acetonitrile, Dimethylformamide, and Dimethylsulfoxide at T = (288.15 to 308.15) K
speedsl	1278.77	m/s	298.15	Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bromide in Acetonitrile, Dimethylformamide, and Dimethylsulfoxide at T = (288.15 to 308.15) K
speedsl	1258.73	m/s	303.15	Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile

speedsl	1238.66	m/s	308.15	Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile
speedsl	1239.00	m/s	308.15	Ultrasonic studies on binary mixtures of some aromatic ketones with acetonitrile at T = 308.15 K
speedsl	1319.14	m/s	288.15	Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bromide in Acetonitrile, Dimethylformamide, and Dimethylsulfoxide at T = (288.15 to 308.15) K
speedsl	1218.52	m/s	313.15	Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile
speedsl	1198.32	m/s	318.15	Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile

speedsl	1278.62	m/s	298.15	Volumetric and Speed of Sound of Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate with Acetonitrile and Methanol at T) (298.15 to 318.15) K
speedsl	1238.66	m/s	308.15	Volumetric and Speed of Sound of Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate with Acetonitrile and Methanol at T) (298.15 to 318.15) K
speedsl	1218.52	m/s	313.15	Volumetric and Speed of Sound of Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate with Acetonitrile and Methanol at T) (298.15 to 318.15) K
speedsl	1198.32	m/s	318.15	Volumetric and Speed of Sound of Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate with Acetonitrile and Methanol at T) (298.15 to 318.15) K
speedsl	1281.30	m/s	298.15	Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T) 298.15 K
speedsl	1278.28	m/s	298.15	Thermodynamic Properties of Inorganic Salts in Nonaqeous Solvents. IV. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Bromides and Chlorides in Acetonitrile

speedsl	1238.40	m/s	308.15	Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bromide in Acetonitrile, Dimethylformamide, and Dimethylsulfoxide at T = (288.15 to 308.15) K
speedsl	1258.51	m/s	303.15	Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bromide in Acetonitrile, Dimethylformamide, and Dimethylsulfoxide at T = (288.15 to 308.15) K
speedsl	1258.73	m/s	303.15	Volumetric and Speed of Sound of Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate with Acetonitrile and Methanol at T) (298.15 to 318.15) K
srf	0.03	N/m	298.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	298.15	Surface tension of non-ideal binary and ternary liquid mixtures at various temperatures and p = 81.5 kPa
srf	0.03	N/m	283.15	Surface tension of non-ideal binary and ternary liquid mixtures at various temperatures and p = 81.5 kPa

srf	0.03	N/m	308.15	Surface tension of non-ideal binary and ternary liquid mixtures at various temperatures and p = 81.5 kPa	
srf	0.03	N/m	308.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.03	N/m	303.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.03	N/m	293.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide	
srf	0.03	N/m	293.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.03	N/m	288.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	
srf	0.03	N/m	283.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile	

srf	0.03	N/m	278.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.03	N/m	293.20	KDB
srf	0.03	N/m	303.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	308.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	313.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	298.15	Physicochemical properties of two 1-alkyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide ionic liquids and of binary mixtures of 1-butyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide with methanol or acetonitrile
srf	0.03	N/m	288.15	Physicochemical properties of two 1-alkyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide ionic liquids and of binary mixtures of 1-butyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide with methanol or acetonitrile

srf	0.03	N/m	308.15	Physicochemical properties of two 1-alkyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide ionic liquids and of binary mixtures of 1-butyl-1-methylpyrrolidinium bis[(trifluoromethyl)sulfonyl]imide with methanol or acetonitrile
srf	0.03	N/m	313.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.03	N/m	293.15	Density and Surface Tension of Binary Mixtures of Acetonitrile + 1-Alkanol at 293.15 K
srf	0.03	N/m	298.15	Experimental Data and Correlation of Surface Tensions of the Binary and Ternary Systems of Water + Acetonitrile + 2-Propanol at 298.15 K and Atmospheric Pressure
srf	0.03	N/m	293.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
srf	0.03	N/m	308.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile

srf	0.03	N/m	313.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
srf	0.03	N/m	303.13	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
srf	0.03	N/m	298.15	Application of the Extended Langmuir Model for the Determination of Lyophobicity of 1-Propanol in Acetonitrile
srf	0.03	N/m	298.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
svapt	111.44	J/mol×K	298.15	NIST Webbook
tcondl	0.20	W/m×K	293.15	Liquid Thermal Conductivities of Acetonitrile, Diethyl Sulfide, Hexamethyleneimine, Tetrahydrothiophene, and Tetramethylethylenediamine
tcondl	0.21	W/m×K	273.15	Liquid Thermal Conductivities of Acetonitrile, Diethyl Sulfide, Hexamethyleneimine, Tetrahydrothiophene, and Tetramethylethylenediamine

tcondl	0.18	W/m×K	313.15	Liquid Thermal Conductivities of Acetonitrile, Diethyl Sulfide, Hexamethyleneimine, Tetrahydrothiophene, and Tetramethylethylenediamine

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.53562e+01
Coeff. B	-3.64591e+03
Coeff. C	-1.52630e+01
Temperature range (K), min.	229.32
Temperature range (K), max.	545.50

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	4.10590e+01
Coeff. B	-4.99962e+03
Coeff. C	-3.88171e+00
Coeff. D	3.51596e-06
Temperature range (K), min.	229.32
Temperature range (K), max.	545.50

Datasets

Mass density, kg/m3

Pressure, kPa - Liquid	Temperature, K - Liquid	Mass density, kg/m3 - Liquid
200.00	313.15	755.0
200.00	353.15	710.0

600.00	313.15	755.0
600.00	353.15	711.0
600.00	393.15	664.0
1000.00	313.15	756.0
1000.00	353.15	711.0
1000.00	393.15	664.0
1000.00	433.15	611.0
1400.00	313.15	756.0
1400.00	353.15	711.0
1400.00	393.15	665.0
1400.00	433.15	612.0
1800.00	313.15	757.0
1800.00	353.15	712.0
1800.00	393.15	666.0
1800.00	433.15	614.0
1800.00	473.15	550.0
2000.00	313.15	757.0
2000.00	353.15	713.0
2000.00	393.15	666.0
2000.00	433.15	614.0
2000.00	473.15	552.0

Reference *https://www.doi.org/10.1021/je900048r*

Refractive index (Na D-line)

Pressure, kPa - Liquid	Temperature, K - Liquid	Refractive index (Na D-line) - Liquid
81.50	298.15	1.3417

Reference *https://www.doi.org/10.1021/je700645p*

Sources

Vapor-Liquid Equilibria for Four Binary Systems at 363.15 K:	https://www.doi.org/10.1021/je020130i
Solid-Liquid Equilibrium of N-Methylformamide + Hexane, + Benzene, + Chlorobenzene, and + Acetonitrile at T = 273.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b01193
Solubility determination and modelling for ophthalmide in mixed solvents of 1,4-dioxane + methanol from 278.15 to 313.15 K:	https://www.doi.org/10.1016/j.jct.2016.11.023
Experimental Determination and Thermodynamic Modeling of Diffusion Coefficients of Liquids in Water, Methanol, and Acetonitrile at Infinite Dilution:	https://www.doi.org/10.1021/je800817b
Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution:	https://www.doi.org/10.1016/j.fluid.2018.09.024

Thermodynamic study of solubility for 2-amino-4-chloro-6-methoxypyrimidine in twelve organic solvents at temperatures from 278.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.10.020
p-Nitrophenylacetonitrile Solubility in Several Pure Solvents: Measurement, Correlation, and Solvent Effect	https://www.doi.org/10.1021/acs.jced.9b00243
Solubility Determination and Thermodynamic Mixing Properties of 5-Methyl-2-pyrazinecarboxylic Acid in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00406
Measurement of activity coefficients at infinite dilution for acetonitrile, water, benzene, tetrachloroethene and their binary pairs.	https://www.doi.org/10.1016/j.fluid.2008.09.011
Solubility Measurement and Thermodynamic Model Correlation of 5-Nitrosalicylaldehyde in Different Solvents.	https://www.doi.org/10.1021/acs.jced.9b00232
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-Hydroxyethyl)-3-methylimidazolium bis(trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1016/j.jct.2013.05.008
Phase diagrams of acetonitrile or (acetone + water + EPPS) buffer phase systems at T = 298.15 K and solubility of cilnidipine in seven pure organic solvents.	https://www.doi.org/10.1016/j.jct.2012.03.026
Measurement and correlation of solubility of 4-(Hydroxymethyl)-1-oxido-2,6,7-trioxa-1-phosphabicyclo[2.2.2]octane in Organic Solvents:	https://www.doi.org/10.1016/j.jct.2016.10.014
Isothermal Vapor–Liquid Equilibria for the Ternary Systems of 4-Methyl-1,3-dioxolan-2-one + 1,1,1-Trimethylbenzene + Decane, + Decane + Ethylbenzene, or + Decane + Propylbenzene at 313.15 K	https://www.doi.org/10.1021/je401056v
Densities of Ionic Liquid [BMIM][BF4] + Ethanol, + Benzene, and + Acetonitrile Binary Systems at (293.15 to 313.15) K	https://www.doi.org/10.1021/je900048r
Excess Enthalpies in Binary Systems of Isomeric C8 Aliphatic Monoethers with Acetonitrile and Their Description by the COSMO-SAC Model	https://www.doi.org/10.1021/acs.jced.5b00838
Diffusion Coefficients, Viscosities, and Conductivities of Some Aromatic Ketones with Acetonitrile at T = 308.15 K:	https://www.doi.org/10.1021/je101130e
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-Butyl-1-methylpyrrolidinium (trifluoromethylsulfonyl)imide.	https://www.doi.org/10.1016/j.jct.2006.01.009
Thermodynamic Study of Solubility for Imatinib Mesylate in Nine Organic Solvents from T = 278.15 K to T = 318.15 K and Mixing Properties of Solutions.	https://www.doi.org/10.1021/je4009816
Solubility of dibenzothiophene in nine organic solvents: Experimental measurement and thermodynamic modelling	https://www.doi.org/10.1016/j.jct.2013.07.004
Solubility of Camellianin A in Methanol, Acetonitrile, Acetone, and Water from (293.15 to 313.15) K.	https://www.doi.org/10.1021/je900554r
Abraham model linear free energy relationships for describing the partitioning and solubility behavior of [bmim][Tf2N] in dry and wet organic solvents.	https://www.doi.org/10.1016/j.fluid.2016.10.009
Thermodynamic Study of Solubility for Imatinib Mesylate in Nine Organic Solvents: Experimental Measurement and Thermodynamic Modelling	https://www.doi.org/10.1021/acs.jced.8b00551
Solubility of dibenzothiophene in pyridine at 298.15 K and solvent effect	https://www.doi.org/10.1016/j.jct.2018.09.017
Solubility determination and thermodynamic functions of 2-Cyanophenothiazine in nine organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions.	https://www.doi.org/10.1016/j.jct.2016.11.029
Apparent Molar Volumes of 1-Alkyl-(n = 2, 4, 6)-3-methylimidazolium Bromides in Water at (288.15 to 318.15) K: Measurement and Correlation.	https://www.doi.org/10.1021/je400473z
Volumetric properties, viscosity and refractive index of the binary mixtures of 1-butyl-3-methylimidazolium octanoate, in molecular solvents:	https://www.doi.org/10.1016/j.jct.2010.01.013
Methanol:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C75058&Units=SI
A model for excess volumes of salty water-acetonitrile mixtures at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2011.09.016
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00425
Solubility measurement and thermodynamic model correlation of salicyline in twelve pure solvents:	https://www.doi.org/10.1021/acs.jced.9b00693
Solubility and Thermodynamic Modeling of Dimethyl Terephthalate in Pure Solvents and the Evaluation of the Mixing Properties of the Solutions:	https://www.doi.org/10.1021/acs.jced.9b00658
Isobaric Vapor–Liquid Equilibrium for the Extractive Distillation of Hexafluoropropane with Acetonitrile and Methanol at Sub-Critical Conditions:	https://www.doi.org/10.1021/je050165t
Solubilities of Phosphorus-Containing Compounds in Selected Solvents:	https://www.doi.org/10.1021/je400531a
Infinite Dilution Activity Coefficients of Solutes Dissolved in Two Phosphonium Ionic Liquids:	https://www.doi.org/10.1021/je100341q
Solubility of Candesartan Cilexetil in Different Solvents at Various Temperatures:	https://www.doi.org/10.1021/je500050p
	https://www.doi.org/10.1021/je101020m
	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1386
Experimental solubility of diosgenin and estriol in various solvents between T = (293.2 328.2) K:	https://www.doi.org/10.1016/j.jct.2016.11.017
Determination and correlation of cefuroxime acid solubility in (acetonitrile + water) mixtures:	https://www.doi.org/10.1016/j.jct.2014.05.018
The Research and Measurement about the Solubility of l-Serine in Eight Common Pure Solvents and Four Binary Mixed Solvents for	https://www.doi.org/10.1021/acs.jced.9b00460
ERAS modeling of the excess molar enthalpies of binary liquid mixtures of acetonitrile and 1-alkanol with (2 to 6) carbon atoms at 298.15 K.	https://www.doi.org/10.1016/j.tca.2006.08.007
Solubilities of succinic acid in different pure solvents and binary mixtures at atmospheric pressure and 293, 298, 318 and 323K.	https://www.doi.org/10.1016/j.tca.2012.03.007
Liquid–liquid equilibrium for ternary systems. Water + 5-Hydroxymethylfurfural + (1-Butanol,	https://www.doi.org/10.1021/acs.jced.8b00120
Conductometric study of some alkali metal halides in dimethyl sulfoxide + acetonitrile at 25 degrees C.	https://www.doi.org/10.1016/j.jct.2009.03.005
Volumetric, Viscometric, and Speed of Sound Properties of Binary Mixtures of Propionic Acid with Acetonitrile at Temperatures (293.15, 298.15, 303.15, and 308.15) K	https://www.doi.org/10.1021/je1005483
Compressibilities of the Mixtures of Acetonitrile with	https://www.doi.org/10.1021/acs.jced.9b00360
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of 1,3-Dimethylacetamide and Propylene Carbonate	https://www.doi.org/10.1021/je200616t
Electrical Conductivity of Solutions of Ionic Liquids with Methanol, Ethanol, Acetonitrile, and Propylene Carbonate:	

Experimental and theoretical study of interaction between organic compounds and 1-(2-Hydroxyethyl)-3-methylimidazolium based ionic liquids:	https://www.doi.org/10.1016/j.fluid.2014.06.021
Thermodynamics of fenofibrate and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.029
Solubility of 2-(2,4,6-Trichlorophenoxy)ethyl Bromide in Methanol, Ethanol, 1-Propanol, Isopropanol, Acetonitrile, 1,4-Dioxane, and Acetone:	https://www.doi.org/10.1021/je200074c
Salting out effect of sodium potassium carbonate, sulfite, tartrate and Phase Equilibrium Measurement and Phase Equilibria of Aqueous Two-Phase Systems of 1-Methyl-2-pyrrolidone + Liquid Salt and Basic Physicochemical Properties for Determination of Solid Liquid Equilibrium of Some Aprotic Solvents:	https://www.doi.org/10.1016/j.fluid.2013.09.028
Solubility Measurement and Phase Equilibria of Aqueous Two-Phase Systems of 1-Methyl-2-pyrrolidone + Liquid Salt and Basic Physicochemical Properties for Determination of Solid Liquid Equilibrium of Some Aprotic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Activity coefficients at infinite dilution measurements for alkanols or alkanones + 1-methyl-2-pyrrolidone: A liquid-liquid equilibrium study	https://www.doi.org/10.1016/j.jct.2012.05.017
Solubility and physicochemical properties for organic solutes and water in the ionic liquid 1-Butyl-3-methylimidazolium Dibutylphosphate:	https://www.doi.org/10.1016/j.jct.2016.12.002
Thermodynamic Models for Determination of the Solubility of 1,10-Phenanthroline in Different Solvents at Temperatures from (278.15 to 328.15) K	https://www.doi.org/10.1021/je500437m
Solubility of Benzophenone in Different Solvents (acetonitrile + alkyl acrylate monomer) Binary Mixtures at Temperatures from (293.15 to 323.15) K	https://www.doi.org/10.1016/j.jct.2016.08.026
Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents:	https://www.doi.org/10.1021/je100315g
Vapor Liquid Equilibrium, Excess Molar Enthalpies, and Excess Molar Volumes of Binary Mixtures Containing 2-Ethoxy-2-methylpropane or 2-Ethoxy-2-methylbutane and Alcohols	https://www.doi.org/10.1021/je301238f
Isothermal compressibility and internal pressure studies of some Ionic Liquid Mixtures at Different Temperatures:	https://www.doi.org/10.1016/j.jct.2006.08.007
Solubility of Acetaminophen in Different Solvents: 2-Ethoxy-2-methylbutane and 2-Methoxy-2-methylbutane + Ethanol Binary Solvent Mixtures at Different Temperatures between (284.15 and 323.15) K and 1-Butyl-3-methylimidazolium Hexafluorophosphate and 1-Butyl-3-methylimidazolium Tetrafluoroborate:	https://www.doi.org/10.1021/je100125x
Efficient determination of crystallization and melting points at high pressure with a novel optical device	https://www.doi.org/10.1021/je700266n
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Ethyl-3-methylimidazolium Ethylsulfate and 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1016/j.jct.2008.05.012
Measurement and correlation of solubility of 1-Ethyl-3-methylimidazolium tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je800218g
Volumetric Properties of Binary Mixtures of Ionic Liquid 1-Ethyl-3-methylimidazolium Tetrafluoroborate + Dimethyl Sulfoxide + Water from (293.15 to 343.15) K	https://www.doi.org/10.1016/j.fluid.2013.06.037
Liquid-liquid equilibrium data for mixtures of ionic liquid 1-ethyl-3-methylimidazolium tetrafluoroborate + (ethyl acetate + ethanol) at 298.15 K	https://www.doi.org/10.1021/je300316a
Solubility and mixing thermodynamic properties of (2,4,6-trimethylbenzoyl) diphenylphosphine oxide in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2016.03.008
Solubilities of N-[4-Bromo-3,5-difluorine)-phenyl]maleimide in Ethanol, Acetone, Tetrahydrofuran, 1,3,2-Dioxaphosphinane 2-methanol-a,a-5,5-tetramethyl 2-oxide, and Trichloromethane at Temperature between 278.15 and 323.15 K and Thermodynamic Modeling of the Solubility of 1-Ethyl-3-methylimidazolium Ethylsulfate in Pure Solvents	https://www.doi.org/10.1016/j.fluid.2018.01.015
Solubility and Solution Thermodynamics of Cefixime Trihydrate in Different Pure Solvents	https://www.doi.org/10.1021/je100819f
Solubility and Solution Thermodynamics of Cefixime Trihydrate in Different Pure Solvents	https://www.doi.org/10.1021/je4003114
Activity coefficients at infinite dilution measurements for alkanols or alkanones + 1-methyl-2-pyrrolidone: A liquid-liquid equilibrium study	https://www.doi.org/10.1016/j.jct.2012.08.016
Solubility and physicochemical properties for organic solutes and water in the ionic liquid 1-Ethyl-3-methylimidazolium Dicyanamide at Different Temperatures	https://www.doi.org/10.1021/acs.jced.6b00163
Modeling of the Solubility of Different Polymorphs of Cefixime Trihydrate in Different Pure Solvents	https://www.doi.org/10.1016/j.jct.2016.03.019
Solubility and Solution Thermodynamic Properties of Mixtures Containing Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate and Ethanol at Different Temperatures	https://www.doi.org/10.1021/je0602723
Determination and correlation of solubility and solution thermodynamics of 1-Ethyl-3-methylimidazolium Ethylsulfate in Different Pure Solvents	https://www.doi.org/10.1016/j.fluid.2016.08.019
Measurements of activity coefficients at infinite dilution for organic solutes and water in 1-hexyl-3-methylimidazolium tris(pentafluoroethyl)trifluorophosphate: Effect of the anion fluorination on the solvent properties	https://www.doi.org/10.1016/j.jct.2013.02.004
Thermodynamic functions for solubility of 2-mercaptobenzothiazole in eleven pure solvents at different temperatures	https://www.doi.org/10.1016/j.jct.2017.05.013
Liquid-Liquid Equilibria for Ternary Systems of FC3283 + HFE7300 + Organic Solvents at Temperatures from 273.15 K to 328.15 K and Mixing Properties of Solutions: 1-Octene + Octane, and Toluene + Octane at 298.15 K	https://www.doi.org/10.1021/je500107y
Solubility and Solvent Effect in the Liquid Phase Equilibria for the Ternary System Ethyl Acetate + Ethanol + 1-Butyl-3-methylimidazolium Acetate at 278.15 K to 298.15 K:	https://www.doi.org/10.1021/acs.jced.8b01101
Separation of binary mixtures based on binary activity coefficients data using specific ammonium based ionic liquid	https://www.doi.org/10.1016/j.fluid.2017.12.029
Liquid-Liquid Equilibria on Seven Binary Systems Ethylene Oxide + Determination and Correlation of Phenol Liquid Equilibria for nine binary 1,2-Dichloropropane + Apparent Molar Volume and Isentropic Compressibility of 1-Butyl-3-methylimidazolium Tetrafluoroborate in Methanol, Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, and 2-Butanol at Different Temperatures	https://www.doi.org/10.1021/je050317k
Solubility and modeling of thermodynamic functions of 1,4-Dihydroxyanthraquinone + Ethanol + Ethyl Acetate + 1,4-Dioxane, 1,3-Hexadiene + Cyclohexane, Toluene + Ethyl Acetate, and Ethanol + 1,4-Butanediol + Cyclohexane	https://www.doi.org/10.1016/j.fluid.2010.10.013
Solubility of Tetracycline hydrochloride in Some Organic Solvents:	https://www.doi.org/10.1021/je900888e
Solubilities of Benzenemethanamine, N-(phenylmethyl)-, Diphenyl Ester in Selected Solvents:	https://www.doi.org/10.1016/j.fluid.2018.06.017
Separation of binary mixtures based on gamma infinity data using [OMIM][BF4] ionic liquid and Viscosities of Acetate or Chloride Based Ionic Liquids and Some Binary Mixtures with Water or Other Common Solvents:	https://www.doi.org/10.1016/j.jct.2012.08.007
Solubility of Ranolazine Dihydrochloride Monohydrate in Different Solvents and Thermodynamic Modeling and Correlation	https://www.doi.org/10.1021/acs.jced.5b00007
Solubility and Solution Thermodynamics of Gibberellin A4 in Different Organic Solvents from 278.15 K to 338.15 K:	https://www.doi.org/10.1016/j.jct.2018.09.003
Density, Electrical Conductivity, Dynamic Viscosity, Excess Properties, and Molecular Interactions of Ionic Liquid 1-Cyanopropyl-3-methylimidazolium Tetrafluoroborate and Binary System with Acetonitrile:	https://www.doi.org/10.1021/je1007235
Solubility and Thermodynamic Functions of Isonicotinamide in Different Pure Solvents and Prediction in Binary Solvent Mixtures	https://www.doi.org/10.1021/acs.jced.8b01181
Determination and Model Correlation of Solubility and Solution Thermodynamics of Lobetyolin in Different Pure Solvents	https://www.doi.org/10.1021/acs.jced.5b00190
Thermodynamics of Gibberellin A4 in Different Organic Solvents from 278.15 K to 338.15 K:	https://www.doi.org/10.1021/acs.jced.7b00898

Thermodynamic Models for Determination of the Solubility of the (1:1) Complex of (Urea + Malic Acid) in Solvent Mixtures:	https://www.doi.org/10.1021/je500991b
Thermodynamic modelling for solubility of 4-nitrobenzaldehyde in binary solvent mixtures at temperature range from 278.15 to 318.15) K and Thermodynamic Study of Some Carbon Dioxide + Organic Binary Systems:	https://www.doi.org/10.1016/j.jct.2016.09.012
Isobaric vapor-liquid equilibria of ethyl acetate + acetonitrile + [HMIM][NTf2] and mixing behaviors of ternary mixtures at 101.3 kPa.	https://www.doi.org/10.1016/j.fluid.2016.06.022
Determination and Correlation of Solubility of 1-Ethyl-3-methylimidazolium Ethylsulfate in Different Solvents from 288.15 to 328.15 K:	https://www.doi.org/10.1021/je0498560
Experimental study and ERAS modeling of the excess molar enthalpy of (acetonitrile + 1-heptanol or + 1-octanol) mixtures at (298.15, 313.15, and 323.15) K and atmospheric pressure.	https://www.doi.org/10.1016/j.jct.2008.01.017
Solubilities of 1-methylimidazolium Bis(2,4,4-trimethylpentyl) Phosphinate in Eight Solvents from 288.20 to 332.15 K.	https://www.doi.org/10.1016/j.jct.2017.10.003
Thermodynamic evidence for nano-heterogeneity in solutions of the macrocycle C-butyrocalix[4]arene in ionic liquids.	https://www.doi.org/10.1021/acs.jced.9b00381
Measurement and correlation of isobaric vapor-liquid equilibrium for the (acetonitrile + water) system at different pressures.	https://www.doi.org/10.1021/je1003115
Effect of Organic Cosolvents on the Solubility of Ionic Liquids in Supercritical CO2:	https://www.doi.org/10.1016/j.jct.2018.06.006
Measurement and Correlation of Solubility of Theobromine, Theophylline, and Caffeine in Water at Different Temperatures:	https://www.doi.org/10.1016/j.jct.2019.07.006
Modeling of Vapor Liquid Equilibrium Solubility of Urea in Acetonitrile + Water Mixtures and Liquid Phase Separation of Urea-Saturated Acetonitrile + Water Mixtures:	https://www.doi.org/10.1016/j.jct.2009.08.002
Solid-Liquid Equilibrium for Esomeprazole Sodium in Monosolvents and in the Ethanol + Ethyl Acetate Binary Solvent Mixtures:	https://www.doi.org/10.1021/je034286+
Solubility and Calorimetric Determination and Modeling:	https://www.doi.org/10.1021/acs.jced.7b00065
Solid Liquid Phase Equilibrium of 3-chloro-N-phenylphthalimide in Ethanol, Acetone, and Acetonitrile:	https://www.doi.org/10.1021/acs.jced.7b00178
Phosphoramidic acid, N,N'-1,2-Ethanediylbis[P,P,P',P'-tetraphenyl- Ester in Selected Solvents:	https://www.doi.org/10.1021/je200122b
Conductivities of Binary Mixtures of Ionic Liquids with Polar Solvents:	https://www.doi.org/10.1021/acs.jced.6b01044
Determination of Activity Coefficients at Infinite Dilution of Organic Solutes in the Ionic Liquid 1-(2-Methoxyethyl)-3-methylimidazolium Hexafluorophosphate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1016/j.jct.2016.01.003
EXCESS MOLAR ENTHALPIES OF DIBROMOMETHANE + 1-(2-Methoxyethyl)-3-methylimidazolium Hexafluorophosphate or 1-Butanesulfonate Using Flow Calorimetry:	https://www.doi.org/10.1021/acs.jced.5b00075
Measurements of Activity Coefficients at Infinite Dilution for Organic Solutes in two Quaternary Ammonium Based Ionic Liquids [DEME][BF4] and [DEME][CF3SO3]:	https://www.doi.org/10.1021/je5010565
Aqueous Solubility Prediction Methodology Based on Partial Solubility Parameters with the Polar Component Subdivided:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Solubility of itaconic acid in different organic solvents: Experimental measurement and thermodynamic modeling:	https://www.doi.org/10.1021/je800468h
Thermodynamic Properties of Binary Mixtures of 1-Ethyl-3-methylimidazolium Ethylsulfate and Organic Solvents from T = (278.15 to 318.15) K and thermodynamic properties of dissolution:	https://www.doi.org/10.1021/acs.jced.8b00080
Solubility and solvation of 1-Butyl-3-methylimidazolium Chloride + Water or Alkanols or Ethanol + Propan-1-ol Binary Mixtures at Different Temperatures:	https://www.doi.org/10.1016/j.tca.2010.01.005
Solubility of citric, malonic, and malic acids in different solvents from 303.2 to 323.2 K and Correlation by the Apelblat Equation:	https://www.doi.org/10.1016/j.fluid.2018.11.011
Activity Coefficients at Infinite Dilution of Organic Compounds in 1-(2-Methoxyethyl)-1-methylpyrrolidinium Hexafluorophosphate and 1-(2-Methoxyethyl)-1-methylpyrrolidinium Tetrafluoroborate Ionic Liquids at Temperatures from 298.15 to 328.15 K:	https://www.doi.org/10.1021/je1006289
Solubility and solvation of 1-Butyl-3-methylimidazolium Chloride + Water or Alkanols or Ethanol + Propan-1-ol Binary Mixtures at Different Temperatures:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1386
Solubility of 1-(2-hydroxy-2-oxoethyl)-2-thiophenecarboxylate in Selected Solvents at Temperatures from (278.15 to 323.15) K:	https://www.doi.org/10.1016/j.fluid.2011.09.027
Solubility of Cyanuric Acid in Different Pure Solvents at Evaluated Temperatures:	https://www.doi.org/10.1016/j.jct.2016.07.001
Activity Coefficients at Infinite Dilution of Organic Compounds in Trihexyl(tetradecyl)phosphonium Tris(pentafluoroethyl)trifluorophosphate Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je900733j
Activity coefficients at infinite dilution and physicochemical properties for organic solvents and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium Tetrafluoroborate:	https://www.doi.org/10.1016/j.jct.2015.11.025
Solid-Liquid Phase Equilibrium of N-(Phosphonomethyl)iminodiacetic Acid + Water + Ethanol Binary Solvent Mixtures:	https://www.doi.org/10.1016/j.fluid.2011.09.033
Solubility of 1-(2-Methoxyethyl)-1-methylpyrrolidinium Hexafluorophosphate in Selected Solvents from T = (283.15 to 323.15) K:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Solubility and Thermodynamic Functions of Isatin in Pure Solvents:	https://www.doi.org/10.1021/je800754w
Liquid Thermal Conductivities of Acetonitrile, Diethyl Sulfide, Hexamethyleneimine, Tetrahydrothiophene, and Tetramethylene Sulfone:	https://www.doi.org/10.1021/je500396b
Activity coefficients at infinite dilution and physicochemical properties for organic solvents and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium Hexafluorophosphate:	https://www.doi.org/10.1021/je0498661
Isobaric Vapor-Liquid Equilibrium for Ethyl Acetate + Acetonitrile + 1-(2-Methoxyethyl)-1-methylpyrrolidinium Tetrafluoroborate Ionic Liquid:	https://www.doi.org/10.1016/j.jct.2013.01.007
Solid-Liquid Phase Equilibrium of N-(Phosphonomethyl)iminodiacetic Acid + Water + Propan-1-ol Binary Solvent Mixtures:	https://www.doi.org/10.1021/je301204w
Solubility of Bis(2,4,4-trimethylpentyl)phosphinic Acid in Selected Solvents from T = (288.15 to 328.15) K:	https://www.doi.org/10.1021/acs.jced.7b00206
Solubility of 1-Ethyl-3-methylimidazolium Bis(2,4,4-trimethylpentyl) Phosphinate in Eight Solvents at Temperatures from (248.2 to 273.2) K:	https://www.doi.org/10.1021/je300401c
Measurement and Correlation of the Solubility of Enrofloxacin in Different Solvents from (308.15 to 321.05) K:	https://www.doi.org/10.1021/je5002158
(Vapor + liquid) equilibrium of the binary mixtures formed by acetonitrile with selected compounds at 98.4 kPa:	https://www.doi.org/10.1016/j.jct.2007.03.007
The use of ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements: Application of Prigogine Flory Patterson theory to excess molar volume and speed of sound of 1-n-butyl-3-methylimidazolium hexafluorophosphate or 1-n-butyl-3-methylimidazolium tetrafluoroborate in methanol and acetonitrile:	https://www.doi.org/10.1016/j.jct.2018.07.024
	https://www.doi.org/10.1016/j.jct.2006.01.016

Interpretation of Association Behavior and Molecular Interactions in Binary Mixtures from Thermoacoustics and Physics and Chemistry of Lithium Halides in 1,3-Dioxolane and Its Binary Mixtures with Acetonitrile probed by Conductometric, Volumetric, Viscometric and Ultrasonic Data.	https://www.doi.org/10.1007/s10765-016-2096-3
Solubility and solution thermodynamics of Dinaphthalenediimide and Perylenediimide with organic solvents: Measurement and Correlation.	https://www.doi.org/10.1016/j.tca.2012.08.009
Solubility determination and thermodynamic modelling of allisartan isoproxil in different binary solvent mixtures from T = (278.15 to 318.15) K	https://www.doi.org/10.1016/j.jct.2016.10.037
Solubility determination and thermodynamic modelling of 4-nitro-3-phenylamine in Binary Mixtures with Dimethyl Sulfoxide and Acetonitrile.	https://www.doi.org/10.1016/j.jct.2007.04.006
Solubility of Sodium Acetate in Binary Mixtures of Methanol, 1-Propanol, Acetonitrile, and Water at 298.2 K:	https://www.doi.org/10.1016/j.jct.2017.03.004
Rebaudioside A in Pure and Binary Mixed Solvents at T = (288.15 to 328.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.035
Liquid-Liquid Equilibrium for the Water + Acetonitrile + Limonene System at different temperatures.	https://www.doi.org/10.1021/je3007474
Solubility determination and thermodynamic modeling of paclobutrazol in nine organic solvents from T = (278.15 to 318.15) K and Determination of the Solubility of Pharmaceutical Solvates at Different Temperatures.	https://www.doi.org/10.1016/j.jct.2016.11.014
Thermodynamic Models for Determination of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions.	https://www.doi.org/10.1021/je8003595
Measurement and Correlation of Solubility of 2-Cyanoacetamide in 14 Pure Solvents from 278.15 to 328.15) K.	https://www.doi.org/10.1016/j.jct.2013.03.018
Phase Equilibrium Properties of Mixtures Containing Ionic Liquids. 8. Activity Coefficients at Infinite Dilution of Organic Solutes in Aprotic Solvents and Alkanes: Ketones, Esters, and Alkylbenzenes.	https://www.doi.org/10.1021/je400625f
Solubility and Solution Thermodynamics of Diphenoxylate in Different Pure Solvents.	https://www.doi.org/10.1021/acs.jced.8b00780
Solubility of Isomeric C8 Alkanes in Liquid 1-Octanol: Influence of Solute Molecular Flexibility and Solvent Cohesion Based on the COSMO-SAC Model.	http://link.springer.com/article/10.1007/BF02311772
Solubility of 2,4,6,8,10,12-hexanitrostilbene in binary mixtures of isomeric dimethylsulfoxide + propanol solvents at different temperatures.	https://www.doi.org/10.1021/je050066f
Thermodynamic models for determination of the solubility of 1,3,5,7-tetranitro-1,3,5,7-tetrazocine in (methanol + acetonitrile) binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2016.09.038
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-butyl-3-methylimidazolium methyl sulfate using inverse gas chromatography.	https://www.doi.org/10.1021/je500867u
Investigation of the Solubility of Octahydro-1,3,5,7-tetranitro-1,3,5,7-tetrazocine and 1,3,5-Triacetyl-hexahydro-s-triazine:	https://www.doi.org/10.1021/acs.jced.8b01205
Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran	https://www.doi.org/10.1021/je0503554
Thermodynamic Evaluation of Solubility and Solution Apparent Volume of Propanolol Hydrochloride in Different Solvent Systems.	https://www.doi.org/10.1021/je200822w
Isobaric Vapor-Liquid Equilibrium for Ethyl Acetate + Acetonitrile, Model Correlation and Solvent Effect Analysis.	https://www.doi.org/10.1021/je5010033
Solubility determination and correlation for o-phenylenediamine in methanol, acetonitrile and water) and Densities, ultrasonic speeds, excess and partial molar properties of binary mixtures of (acetonitrile + amines) at 298.15 K.	https://www.doi.org/10.1016/j.jct.2014.12.025
Solubility Parameter Dependent Model for Predicting the Solubility Profile and Application of Thermodynamic Properties of Acetonitrile Mixtures at 298.2 K:	https://www.doi.org/10.1016/j.jct.2016.06.028
Solubility and thermodynamic properties of dirithromycin form A and form B in pure solvents and binary solvent mixture.	https://www.doi.org/10.1016/j.jct.2015.12.004
Thermodynamic Functions for the Solubility of 3-Nitrobenzonitrile in 12 Organic Solvents from T/K = (278.15 to 318.15).	https://www.doi.org/10.1016/j.jct.2014.08.012
Solubility Measurement and the Correlation of 1-Naphthaleneacetic Acid in Pure and Methanol + Water Binary Solvents from T = (278.25 to 323.55) K:	https://www.doi.org/10.1016/j.fluid.2014.01.008
	https://www.doi.org/10.1016/j.jct.2014.04.020
	https://www.doi.org/10.1021/je301237t
	https://en.wikipedia.org/wiki/Joback_method
	https://www.doi.org/10.1016/j.jct.2011.09.022
	https://www.doi.org/10.1021/acs.jced.9b00562
	https://www.doi.org/10.1021/acs.jced.5b00167
	https://www.doi.org/10.1016/j.fluid.2017.09.007
	https://www.doi.org/10.1021/acs.jced.9b00458
	https://www.doi.org/10.1021/acs.jced.5b00995
	https://www.doi.org/10.1016/j.jct.2016.10.018
	https://www.doi.org/10.1016/j.jct.2018.03.013
	https://www.doi.org/10.1016/j.jct.2015.09.002
	https://www.doi.org/10.1016/j.jct.2018.12.035
	https://www.doi.org/10.1021/je020216b
	https://www.doi.org/10.1016/j.jct.2018.12.044
	https://www.doi.org/10.1016/j.jct.2016.11.030
	https://www.doi.org/10.1021/acs.jced.7b00615
	https://www.doi.org/10.1021/acs.jced.6b00816

Solubility of Trimethoprim in Selected Pure Solvents and (Water + Ethanol/2-Propanol) Mixed-Solvent Systems:	https://www.doi.org/10.1021/acs.jced.5b00616
Solubility Measurement and Thermodynamic Modeling of 4-Nitrophthalimide in Twelve Pure Solvents at Elevated Temperatures Ranging from (273.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.6b00230
Measurement and correlation of liquid liquid equilibria for acetonitrile +	https://www.doi.org/10.1016/j.fluid.2007.01.024
Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate:	https://www.doi.org/10.1016/j.jct.2012.09.017
Pressure-Drop Method for Detecting Bubble and Dew Points of Multicomponent Mixtures at Temperatures up to 573 K:	https://www.doi.org/10.1021/acs.jced.7b00755
Solubility Modeling and Mixing Thermodynamics of Thiamphenicol in Water and Twelve Neat Organic Solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.7b00542
Solubility determination and thermodynamic modelling for 2-amino-4-chlorobenzoic acid in eleven organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.019
Solubility Measurement and Thermodynamic Modeling of N-(4-Methylphenyl)-Z-3-chloro-2-(phenylthio)propenamide in Twelve Pure Solvents at Temperatures Ranging from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b01011
Solubilities of phosphoric acid P,P'-1,4-phenylene [illegible] tetraphenyl ester in the organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.12.034
Solubility Increment and Thermodynamic Analysis of Bioactive Antofloxacin Hydrochloride in Aqueous ChCl/PTS Deep Eutectic Solvent and Cosolvent Mixtures:	https://www.doi.org/10.1021/acs.jced.9b00717
Activity Coefficients at Infinite Dilution of Organic Compounds in Four New Imidazolium-Based Ionic Liquids:	https://www.doi.org/10.1021/je200195q
Molar heat capacities of the (water + acetonitrile) mixtures at T = (283.15, 298.15, 313.15, and 328.15) K:	https://www.doi.org/10.1016/j.jct.2010.04.019
Thermodynamics and activity coefficients at infinite dilution for organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
Experiment and Computation of Solubility and Dissolution Properties for Enalapril Maleate and Its Intermediate in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00333
Experimental Data and Correlation of Surface Tensions of the Binary and Ternary Systems of Water + Acetonitrile + 2-Propanol at 298.15 K and Atmospheric Pressure:	https://www.doi.org/10.1021/je1001329
Solubility of Gastrodin in Pure and Mixed Solvents at 273.15-313.15 K and its Correlation with Different Thermodynamic Models:	https://www.doi.org/10.1021/acs.jced.9b00286
Solubility and Melting Properties of Salicylamide:	https://www.doi.org/10.1021/je060178m
Effect of imidazolium-based ionic liquid on vapor-liquid equilibria of 2-propanol + acetonitrile binary system at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2015.10.035
Equilibrium Solubility of Biapenem in Different Neat and Binary Solvents: Experimental Determination and Model Correlation:	https://www.doi.org/10.1021/acs.jced.8b01051
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity Coefficients at Infinite Dilution of Hydrocarbons, Alcohols, Esters, and Aldehydes in Trimethyl-butylammonium Bis(trifluoromethylsulfonyl) Imide Using Gas-Liquid Chromatography and Static Method:	https://www.doi.org/10.1021/je050440b
Thermodynamic properties of ternary mixture {[C4mim][SCN] + acetic or propionic acid + acetonitrile} over the temperature range of (293.15-313.15) K:	https://www.doi.org/10.1016/j.jct.2019.07.002
Solubility of Flavonoids in Organic Solvents:	https://www.doi.org/10.1021/je7001094
Thermodynamic study of solubility for pyrazinamide in ten solvents from T = (283.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2017.04.014
Solubility, solid-liquid equilibrium and mixing properties of 2-Cyano-4'-methylbiphenyl in pure solvents:	https://www.doi.org/10.1016/j.jct.2016.07.050
Activity coefficients at infinite dilution of organic solvents and water in 1-butyl-3-methylimidazolium dicyanamide. A literature review of hexane/hex-1-ene separation:	https://www.doi.org/10.1016/j.fluid.2016.02.004
Solubility of Thiotriazinone in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, Acetonitrile, Acetone, Ethyl Acetate, and Water from 293 K to 348 K:	https://www.doi.org/10.1021/je900177h
Thermostabilities and solubilities of flame retardant bisphenol S bis (5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy phosphate ester) as a function of temperature:	https://www.doi.org/10.1016/j.fluid.2014.04.006
Solubility of (Z)-2-Amino-α-(methoxyimino)-4-thiazole-ethanethioic Acid S-2-Benzothiazolyl Ester in Pure Solvents and a Mixture of Acetonitrile and Dichloromethane:	https://www.doi.org/10.1021/je900784v
Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents:	https://www.doi.org/10.1016/j.jct.2015.02.004
Ionic molar volumes in methanol mixtures with acetonitrile, N,N-dimethylformamide and propylene carbonate at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2014.02.021
Solvation parameter model and thermodynamic parameters in a dicationic ionic liquid based on pyrrolidinium:	https://www.doi.org/10.1016/j.jct.2014.12.023
Henry's constants and activity coefficients of some organic solutes in 1-butyl,3-methylimidazolium hydrogen sulfate and in 1-methyl,3-trimethylsilylmethylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2012.08.022
Liquid-liquid equilibria measurements of ternary systems (acetonitrile + a carboxylic acid + dodecane) at 303.15 K:	https://www.doi.org/10.1016/j.fluid.2014.12.013
Thermodynamic study of molecular interaction-selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Activity Coefficients at Infinite Dilution for Organic Compounds Dissolved in 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids Having Six-, Eight-, and Ten-Carbon Alkyl Chains:	https://www.doi.org/10.1021/je300692s
Dortmund Data Bank Vapor-Liquid Equilibrium Data:	http://www.ddbst.com/en/EED/VLE/VLE%20Acetonitrile%3B1,2-Ethanediol.php
Isothermal Vapor-Liquid Equilibrium of Binary Mixtures Containing 1-Chlorobutane, Ethanol, or Acetonitrile:	https://www.doi.org/10.1021/je800535c
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Solubility of Acetoguanamine in Twelve Neat Solvents from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00593
Solubility of 3,3'-(Oxidi-2,1-ethanediyl)bis[1-methyl-imidazolium] Dihexafluorophosphate in Water, Methanol, Ethanol, Acetone, and Acetonitrile and Binary Mixtures (Water + DMF) from (283.00 to 323.00) K:	https://www.doi.org/10.1021/je501026m
Thermodynamic functions for solubility of 3-nitro-o-toluic acid in nine organic solvents from T = (283.15 to 318.15) K and apparent thermodynamic properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.017
Evaluation of the Performance of Trigeminal Tricationic Ionic Liquids for Separation Problems:	https://www.doi.org/10.1021/je201129y

Solubility in Binary Solvent Mixtures: Pyrene Dissolved in Alcohol + Acetonitrile Mixtures at 293.2 K:	https://www.doi.org/10.1021/je020226c
Surface tension of non-ideal binary and ternary liquid mixtures at various temperatures and p = 81.5 kPa:	https://www.doi.org/10.1016/j.jct.2010.09.003
Solubility and thermodynamic functions of tebuconazole in nine organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.032
Solubilities of Three Flavonoids in Different Natural Deep Eutectic Solvents at T = (288.15 to 328.15) K:	https://www.doi.org/10.1021/acs.jced.6b00552
Viscosity, Density, and Refractive Index of Binary Mixtures of Organic Solvent + Ionic Liquid, 2-Chloronicotinic Acid and Analysis of Thermodynamic Properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide:	https://www.doi.org/10.1021/je700338t
Solubility Determination of 2-Chloronicotinic Acid in Ethyl 2-Methyl Imidazolium Hexafluorophosphate at 298.15 K:	https://www.doi.org/10.1021/acs.jced.9b00661
Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide:	https://www.doi.org/10.1016/j.jct.2013.02.021
Solubility and solution thermodynamics of 2-Ethyl-6-nitroaniline in different organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.05.027
Determination and correlation thermodynamic models for solid liquid equilibrium of the Nifedipine in pure and binary solvent systems at 313.15-363.15 K:	https://www.doi.org/10.1016/j.jct.2016.07.035
Solubility of 2-Fluoronicotinic Acid in Eight Pure Solvents and Two Binary Solvent Systems at 313.15-363.15 K:	https://www.doi.org/10.1021/acs.jced.7b00927
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in Tributylethylphosphonium Diethylphosphate Using Gas Liquid Chromatography, Thermodynamic Properties of Mixtures Containing Ionic Liquids.	https://www.doi.org/10.1021/acs.jced.5b00980
Solubility of daidzin in different organic solvents and (ethyl alcohol + water) Measurement and correlation of the excess molar enthalpy of (1-propanol or 2-propanol or 2-methyl-2-propanol + acetonitrile) mixtures at 298.15 K and thermodynamic and thermodynamic properties of 1,3-Dimethylimidazolium Dimethylphosphate: Solubility, Activity Coefficients at Infinite Dilution, Synthesis of 1,3-Dimethylimidazolium Methylsulfate and Viscosity, Density, and Refractive Index of Binary Mixtures of Acetonitrile and Alkyl Acetates at Atmospheric Pressure:	https://www.doi.org/10.1016/j.jct.2015.04.025
Solubility of Calcium Diphosphate and 4,7-Dichloroquinoline in Water-Ethanol Cosolvent Mixtures at 101.3 kPa: Correlation, Dissolution Thermodynamic Properties and Preferential Solvation:	https://www.doi.org/10.1016/j.jct.2011.09.011
Solubility of Caffeine in Pure and Binary Solvent Mixtures of Ethanol + Dimethyl Sulfoxide and Ethanol + Acetonitrile: Measurement, Correlation, Molecular Simulations and Thermodynamic Analysis:	https://www.doi.org/10.1016/j.jct.2015.05.022
Measurement and Correlation of Solubility of 1-(4-Methoxyphenyl)-4-methyl-1,2,3,4-tetrahydroisoquinoline in Different Pure Solvents and Binary Solvent Mixtures of Ethanol + Ethyl Acetate from 278.15 to 323.15 K:	https://www.doi.org/10.1007/s10765-009-0651-x
Solubility of Caffeine and Theophylline in Binary Solvent Mixtures of Acetonitrile and Alkyl Acetates:	https://www.doi.org/10.1021/acs.jced.9b00526
Solubility of Tris(2-methylphenyl)phosphate: Determination and Correlation in Two-Phase System Equilibrium Data	https://www.doi.org/10.1021/acs.jced.8b00176
Interactions in Binary and Ternary Mixtures of Ionic Liquids: Thermodynamic Study of (vapour + liquid) equilibria in the ternary system (acetonitrile + 1-butyl-3-methylimidazolium trifluoromethanesulfonate) and its constituent binary mixtures at T = 348.15 K	https://www.doi.org/10.1016/j.jct.2013.08.011
Solubility of Chloroquine Diphosphate and 4,7-Dichloroquinoline in Water-Ethanol Cosolvent Mixtures at 101.3 kPa: Correlation, Dissolution Thermodynamic Properties and Preferential Solvation:	https://www.doi.org/10.1021/je8007099
Solubility of Ethyl Acetate, Acetonitrile, and Acetone from (298.2 to 393.2) K: Model Correlation of Data:	https://www.doi.org/10.1021/acs.jced.9b00331
N-[(4-Bromo-3,5-difluorine)phenyl]acrylamide Calorimetric Study of Nitrile, in Benzene, Methanol, Pyridine, Chain Solvent Interactions and Dipolar Liquid Equilibria for the Ternary Systems Acetonitrile (DMSO) Propanol, Dimethyl Sulfoxide and the Corresponding Binary Systems at 101.3 kPa:	https://www.doi.org/10.1021/je900609f
Solubility of Phenylbutazone in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/je100489z
Liquid 1-Butyl-3-methylimidazolium Tetrafluoroborate, in Organic Solvents at T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/je100033s
Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from T = (278.15 to 318.15) K	https://www.doi.org/10.1016/j.fluid.2010.08.010
Thermodynamic modeling of activity coefficients at infinite dilution of organic solutes and water in 1-hexyl-3-methylimidazolium tetrafluoroborate ionic liquid	https://www.doi.org/10.1021/acs.jced.6b00911
Solubility of Phenylbutazone in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1007/s10765-008-0395-z
Liquid 1-Butyl-3-methylimidazolium Tetrafluoroborate, in Organic Solvents	https://www.doi.org/10.1016/j.jct.2016.07.009
Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from T = (278.15 to 318.15) K	https://www.doi.org/10.1016/j.jct.2016.07.023
Thermodynamic modeling of activity coefficients at infinite dilution of organic solutes and water in 1-hexyl-3-methylimidazolium tetrafluoroborate ionic liquid	https://www.doi.org/10.1016/j.jct.2013.09.007
Thermodynamics and selectivity of aqueous biphasic systems with phosphonium-based ionic liquids at 298.15 K and 310.15 K and binding properties of proteins	https://www.doi.org/10.1016/j.jct.2018.05.017
Solubility and Solution Thermodynamics of Phosphaspirocyclic Compounds in Different Solvents:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
1-Hexyl-3-methylimidazolium Tetrafluoroborate and its Binary Mixtures with Water and Alcohols:	https://www.doi.org/10.1021/acs.jced.9b00696
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium ethylsulfate and modeling of the solubility of 1-hexyl-1,4-diaza-2,2,2-bicyclooctanium chloride in aqueous solution thermodynamics of 1,6-dimethylurea in different solvents from (288.15 to 328.15) K:	https://www.doi.org/10.1021/je060138i
Solubility of Some Phosphaspirocyclic Compounds in Different Solvents:	https://www.doi.org/10.1021/acs.jced.8b00684
Thermodynamics and activity coefficients at infinite dilution for measurement of activity coefficients at infinite dilution of organic solutes in 1-ethyl-3-methylpyrrolidinium ethylsulfate ionic liquid and modeling of the data	https://www.doi.org/10.1016/j.fluid.2018.06.013
Measurement and correlation of activity coefficients at infinite dilution of organic solutes in 1-hexyl-1,4-diaza-2,2,2-bicyclooctanium chloride ionic liquid and thermodynamics of solution	https://www.doi.org/10.1016/j.jct.2013.10.026
The Water Solubility Enhancement of Coumarin Drug by Novel Ionic Liquids: Experimental Measurement and Computational Modeling	https://www.doi.org/10.1016/j.jct.2016.07.043
Solubility and solution thermodynamics of 1,6-dimethylurea in different solvents from (288.15 to 328.15) K: A new approach for calculating partial molar volumes:	https://www.doi.org/10.1016/j.jct.2015.12.032
Solution thermodynamics of iron(III)-N,N'-ethylene-bis(salicylideneiminato)-chloride in binary mixtures of N,N-dimethylformamide and Organometallic Derivatives of Some Copper(I), Silver(I), and Gold(I) Salts in Acetonitrile at Several Temperatures:	https://www.doi.org/10.1016/j.jct.2005.08.009
Determination and Modeling of Solubility of 4-Aminobenzamide in Acetonitrile at Several Temperatures from (278.15 to 313.15) K	https://www.doi.org/10.1016/j.jct.2014.02.019
Separation of aliphatic from aromatic hydrocarbons and sulphur compounds from fuel based on measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid N,N-diethyl-N-methyl-N-(2-methoxy-ethyl)ammonium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1021/acs.jced.8b01144
	https://www.doi.org/10.1021/je900915p
	https://www.doi.org/10.1016/j.jct.2016.07.017

Solubility of 2-nitro-p-phenylenediamine in nine pure solvents and mixture of (methanol + ethyl acetate) from (283.15 to 318.15) K: Determination and modelling.	https://www.doi.org/10.1016/j.jct.2017.01.006
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 283.15 to 318.15 K.	https://www.doi.org/10.1021/acs.jced.9b00490
Solubility Measurement and Correlation for 2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Five Organic Solvents at Temperatures between 288.15 and 328.15 K.	https://www.doi.org/10.1021/acs.jced.6b00664
Solubility modelling and thermodynamic dissolution functions of 2,4,6,8,10,12-hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Pure Organic Solvents: experimental determination and prediction of gas solubility data for Ethyl Acetate + Ethanol Binary Solvent Mixtures at Temperatures between 283.15 and 323.15 K.	https://www.doi.org/10.1016/j.jct.2015.10.024
Solubility of 3-Methylpyridazine in Pure and Binary Solvent Mixtures and Investigation of Intermolecular Interaction.	https://www.doi.org/10.1016/j.jct.2004.07.014
Solubility Measurement and Modeling of 4,4′-Dihydroxydiphenyl Sulfone in Nine Organic Solvents at Temperatures from 278.15 to 313.15 K and Thermodynamic Properties of Dissolution.	https://www.doi.org/10.1021/acs.jced.7b01134
Solubility of Sulfamethoxypyridazine in Some Binary Solvents at Various Temperatures and Correlation with Jouyban–Acree Model.	https://www.doi.org/10.1021/je5007218
Solubility Measurement and Modeling of 4,4′-Dihydroxydiphenyl Sulfone in Nine Organic Solvents at Temperatures from 278.15 to 313.15 K and Thermodynamic Properties of Dissolution.	https://www.doi.org/10.1021/acs.jced.5b00714
Solubility of Olaparib in Selected Organic Solvents and Thermodynamic Properties of Solution.	https://www.doi.org/10.1016/j.jct.2014.12.027
Thermodynamic Properties of (Tetrahydrofuran + MTBE) + Isopropyl Alcohol Ternary Solvent Mixtures: Experimental Solid–Liquid and Liquid–Liquid Equilibria and Excess Molar Volume of Alkanol + Acetonitrile, 1-Octanol + Butan-1-ol Binary Mixtures.	https://www.doi.org/10.1016/j.jct.2016.02.006
Vapor–Liquid Equilibrium for the Binary Methyl Esters (Butyrate, Pentanoate, or Hexanoate) + Acetonitrile (2) Systems at 99.32 kPa.	https://www.doi.org/10.1016/j.fluid.2018.07.028
Crystal Structures, Reaction Rates, and Selected Physical Properties of Tetrabutylammonium Bromide Ionic Liquid Solvates.	https://www.doi.org/10.1021/acs.jced.7b00489
Measurement and Modeling of Solid–Liquid Equilibria of Glutamic Acid in Pure Solvents and Aqueous Binary Mixtures.	https://www.doi.org/10.1021/acs.jced.8b00196
Solubility determination and thermodynamic modelling of 5-phenyltetrazole in thirteen mono solvents and liquid mixtures of (methanol + ethyl acetate) at related temperatures.	https://www.doi.org/10.1021/je050262m
Activity Coefficients at Infinite Dilution of Polar Solutes in Ionic Liquids by Inverse Gas–Liquid Chromatography.	https://www.doi.org/10.1016/j.fluid.2007.03.019
Interactions of volatile organic compounds with the ionic liquid 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Determined by Comparative Ebulliometry.	https://www.doi.org/10.1021/je060001s
Solubility determination and thermodynamic modelling of 3-amino-1,2,4-triazole in ten organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions.	https://www.doi.org/10.1021/je800557h
Activity Coefficients at Infinite Dilution for Organic Solutes Dissolved in Three 1-Alkyl-1-methylpyrrolidinium Bis(trifluoromethylsulfonyl)imide Ionic Liquids.	https://www.doi.org/10.1021/je3005112
Solubility determination and thermodynamic modelling of 3-amino-1,2,4-triazole in ten organic solvents from T = (273.15 to 318.15) K.	https://www.doi.org/10.1021/acs.jced.8b01084
Liquid–liquid equilibria in the ternary systems {hexadecane + BTX aromatics + 2-methoxyethanol or acetonitrile} at 298.15 K.	https://www.doi.org/10.1016/j.jct.2017.04.019
Solution Thermodynamics of Benzotriazole in Different Pure and Binary Mixtures of Ionic Liquid 1-Butyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide with Acetonitrile and Tetrahydrofuran.	https://www.doi.org/10.1016/j.jct.2012.09.033
Activity Coefficients at Infinite Dilution of Polar Solutes in a Novel Superbase-Derived Protic Ionic Liquid: 2-Hydroxyethylammonium 2-Cyanopyrrolide.	https://www.doi.org/10.1021/je700640r
Solubility determination and thermodynamic modelling of 3-amino-1,2,4-triazole in ten organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions.	https://www.doi.org/10.1016/j.fluid.2012.11.008
Partial Molar Volumes of Butyltriethylammonium Iodide in Single Solvents and Binary Mixtures of Acetonitrile + Water at Temperatures from 288.15 to 318.15 K.	https://www.doi.org/10.1016/j.jct.2016.09.033
Solubility of Valsartan in Different Organic Solvents and Ethanol + Water Co-solvence phenomenon and thermodynamic properties of dissolution.	https://www.doi.org/10.1016/j.jct.2011.09.028
Volumetric and Isentropic Compressibility Behavior of Ionic Liquid, 1-Propyl-3-Methylimidazolium Bis(trifluoromethylsulfonyl)imide in Acetonitrile + Dimethylformamide and Dimethylsulfoxide at 298.15 K.	https://www.doi.org/10.1021/je4001894
Measurement and correlation of the solubility of genistin in eleven organic solvents from T = (288.2 to 323.2) K: Solubility of Danthron: Measurement of activity coefficients at infinite dilution in 1-hexadecyl-3-methylimidazolium tetrafluoroborate ionic liquid: 2,5-thiophenedicarboxylic acid in different organic solvents:	https://www.doi.org/10.1021/je900542y
	https://www.doi.org/10.1021/je034144f
	https://www.doi.org/10.1016/j.jct.2016.07.013
	https://www.doi.org/10.1021/acs.jced.7b01085
	https://www.doi.org/10.1016/j.fluid.2015.01.002
	https://www.doi.org/10.1021/je400781b
	https://www.doi.org/10.1016/j.jct.2011.04.010
	https://www.doi.org/10.1016/j.jct.2019.05.011
	https://www.doi.org/10.1016/j.jct.2006.07.014
	https://www.doi.org/10.1021/je8007815
	https://www.doi.org/10.1016/j.jct.2019.06.018
	https://www.doi.org/10.1007/s10765-009-0648-5
	https://www.doi.org/10.1021/je200252c
	https://www.doi.org/10.1016/j.jct.2015.05.019
	https://www.doi.org/10.1021/acs.jced.5b00192
	https://www.doi.org/10.1016/j.jct.2007.01.004
	https://www.doi.org/10.1016/j.fluid.2014.05.003

Volumetric and compressibility behaviour of ionic liquid, 1-H butyl-3-methylimidazolium tetrafluoroborate and 1-butyl-3-methylimidazolium hexafluorophosphate at T=(288.15 to 318.15) K	https://www.doi.org/10.1016/j.jct.2005.07.018
Exploration of Solvation Consequence of Ionic Liquid [Bmim][HSO3] in Several Aprotic Solvents by Volumetric and Viscometric Approach	https://www.doi.org/10.1021/acs.jced.5b00670
Research on Dissolution Capability of Tetrabutylammonium Salts, Hexafluorophosphate Ionic Liquids, and Solvents at T = 298.15 K	https://www.doi.org/10.1021/acs.jced.8b00288
Vapor-Liquid Equilibrium for Acetonitrile + Propanenitrile and 2-Methylpropanenitrile + Propanenitrile	https://www.doi.org/10.1021/je025660t
Solubility Modeling of 4-(Methylsulfonyl)benzaldehyde in Nine Monosolvents and Two Binary Mixtures of 1-Methoxy-2-propanol + Organic Solvents at Elevated Temperatures	https://www.doi.org/10.1021/acs.jced.5b01053
Solubility Measurement and Thermodynamic Modeling for o-Toluenesulfonamide in 16 Solvents from 278.15 to 323.15 K	https://www.doi.org/10.1021/acs.jced.9b00445
Thermodynamic Solubility and Mixing Properties of Phenazopyridine in 14 Pure Solvents at Temperatures Ranging from 278.15 to 323.15 K	https://www.doi.org/10.1021/acs.jced.9b00844
Solubility Measurement and Modeling of 2-Amino-4,6-dichloropyrimidine in Some Pure Solvents and Ethanol + Ethyl Acetate Solvent Mixtures.	https://www.doi.org/10.1021/acs.jced.8b00292
Isopiestic vapor-liquid equilibrium of ethanenitrile + water + 1,2-ethanediol + 1-ethyl-3-methylimidazolium chloride: thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic solvents from (278.15 to 323.15) K.	https://www.doi.org/10.1016/j.fluid.2014.06.025
Solubility determination and thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic solvents from (278.15 to 323.15) K.	https://www.doi.org/10.1016/j.jct.2016.04.018
Viscosity of binary mixtures of 1-ethyl-3-methylimazolium ethylsulfate and 1-butyl-3-methylimidazolium ethylsulfate at (293.15 and 313.15) K	https://www.doi.org/10.1016/j.jct.2016.06.015
Separation of acetonitrile and ethanol azeotropic mixture using 1-ethyl-3-methylimidazolium ethylsulfate as entrainer.	https://www.doi.org/10.1016/j.fluid.2018.08.009
Viscosity of binary mixtures of imidazolium-based ionic liquids as solvents in cyclohexane/cyclohexene and hex-1-ene/hex-1-ene,	https://www.doi.org/10.1016/j.jct.2017.11.017
Solubility Determination of 1,5-Naphthalenediamine and Solubility Determination and Modeling for 4,4′-Dihydroxydiphenyl Sulfone in Several Pure Solvents and Binary Solvent Mixtures	https://www.doi.org/10.1021/acs.jced.8b01214
Solubility Determination and Modeling of Acetone, Ethyl Acetate, and Their Binary Solvent Mixtures at Different Temperatures	https://www.doi.org/10.1021/acs.jced.6b00421
Solubility of Quercetin in Different Monosolvents of Acetone, Ethyl Acetate, 1-Propanol, Acetonitrile, and Dimethyl Sulfoxide and Their Binary Mixtures at 313.15 K	https://www.doi.org/10.1016/j.fluid.2013.06.038
Solubility, Viscosity, and Density of Acetonitrile + Ethanol from (278.15 to 318.15) K and Their Binary Mixtures with Dimethylsulfoxide at 298.15 K	https://www.doi.org/10.1021/je8001909
Solubility Determination and Thermodynamic Modeling of Acetonitrile in the Range (288.15 to 318.15) K	https://www.doi.org/10.1021/acs.jced.8b01099
Ibuprofen in Different Solvents and Gibbs correlations for ternary excess/deviation properties of [(x)2mim][TfSO4] (x1) + acetic or propionic acid (x2) + acetonitrile (x3) systems at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.12.014
Thermodynamic Dissolution Properties of Ibuprofen in Some Organic Solvents at Different Temperatures	https://www.doi.org/10.1021/je9003178
Solubility and Solution Thermodynamics of 1-Hydroxybenzotriazole in Nine Monosolvents at Temperatures from (278.15 to 323.15) K and Mixing Properties of Solutions.	https://www.doi.org/10.1021/acs.jced.7b00316
Thermodynamic equilibrium of hydroxyacetic acid in pure and binary solvent mixtures from 283.15 to 323.15 K.	https://www.doi.org/10.1021/acs.jced.9b00778
Solubility and Melting Properties of Salicylic Acid:	https://www.doi.org/10.1016/j.jct.2017.01.004
Activity coefficients at infinite dilution of organic solutes in	https://www.doi.org/10.1021/je060134d
Solubility of Imipramine Hydrochloride in Different Pure Solvents	https://www.doi.org/10.1016/j.fluid.2012.05.006
Supercritical CO2 Containing Different Cosolvents	https://www.doi.org/10.1021/acs.jced.7b01108
Solubility of Tiamulin Hydrogen Fumarate in Acetone, Acetonitrile, Ethyl Acetate, Ethyl Formate, and Butyl Acetate from (288.2 to 328.2) K	https://www.doi.org/10.1021/je800838w
Solubilities of Bis(2,4,6-trimethylbenzoyl)phenylphosphine Oxide in Different Organic Solvents at Several Temperatures	https://www.doi.org/10.1021/je300343y
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.03.015
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2018.01.003
Solubility Modeling and Solvent Effect Analysis of Aripiprazole in Neat Solvents and Binary Solvent Mixtures at 278.15 K to 323.15 K	https://www.doi.org/10.1021/acs.jced.8b00430
Solubility determination and thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide in gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Molecular interaction studies and theoretical estimation of ultrasonic speeds using scaled particle theory in binary mixtures of toluene with isomeric heptanones at different temperatures	https://www.doi.org/10.1016/j.tca.2009.10.001
Solid–Liquid Equilibrium of 5-Chloro-8-hydroxyquinoline and 5,7-Dichloro-8-hydroxyquinoline in Different Solvents and Mixing Properties of Solutions	https://www.doi.org/10.1021/je901024t
Solubility Determination, Correlation, and Thermodynamic Analysis of Nitrofurantoin in Different Pure Solvents and Binary Solvent Mixtures at 278.15 K to 323.15 K	https://www.doi.org/10.1021/acs.jced.9b00028
Solubility Determination, Correlation, and Thermodynamic Dissolution Functions of 1,3-Diphenylguanidine in nine Pure Solvents and Ethanol + Ethyl Acetate Binary Solvent Mixtures at Various Temperatures	https://www.doi.org/10.1016/j.jct.2019.06.007
Determination of Solid–Liquid Equilibria and Thermodynamic Properties of Ibuprofen in Pure and Binary Solvents	https://www.doi.org/10.1021/je800404f
Vapor–Liquid Equilibria Measurement of (Methanol + Ethanenitrile + Bis(trifluoromethylsulfonyl) imide)-Based Ionic Liquids at 101.3 kPa	https://www.doi.org/10.1016/j.jct.2016.03.011
Thermodynamics of Flufenamic Acid Dissolution in Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.9b00172
Solubility of Tetramethylpyrazine in Nine Monosolvents and Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.5b00796
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1021/acs.jced.9b00278
	https://www.doi.org/10.1021/acs.jced.8b00888
	https://www.doi.org/10.1016/j.jct.2011.11.021

Determination and correlation of solubilities of 1,2-bis(2-oxo-5,5-dimethyl-1,3,2-dioxyphosphacyclohexyl-2-imine)	https://www.doi.org/10.1016/j.fluid.2013.09.058
Isobaric Vapor-Liquid Equilibrium of the Acetonitrile + Propanol + Ionic Liquid 1-Ethyl-3-methylimidazolium Ethyl Sulfate System at 101.3 kPa	https://www.doi.org/10.1021/acs.jced.9b00015
Solubility Determination and Modeling of p-Nitrobenzamide Dissolved in Twelve Neat Solvents from 283.15 to 328.15 K	https://www.doi.org/10.1021/acs.jced.9b00065
Solubility of N-phenylanthranilic acid in nine organic solvents from T = (283.15 to 313.15) K: Determination and Modeling.	https://www.doi.org/10.1016/j.jct.2016.08.017
Solubility of 2-Chloronaphthalene and 2-Chloroanthraquinone in 12 Pure Solvents from T = (273.15 to 313.15) K.	https://www.doi.org/10.1021/acs.jced.8b00139
(Liquid + liquid and liquid + liquid equilibrium for ternary system water + acetonitrile + cyclohexene at 298.15 K.	https://www.doi.org/10.1016/j.fluid.2015.06.039
Solubility Measurements at 296 and 310 K and Physicochemical Characterization of Abiraterone and Abiraterone Acetate.	https://www.doi.org/10.1021/acs.jced.8b00566
Solubility and Solvent Effect of 1-(2-Bromo-phenyl)-pyrrole-2,5-dione in Mono-Solvents from 278.15 to 323.15 K	https://www.doi.org/10.1021/acs.jced.9b00560
Measurement and correlation for the Solubility of Dimethyl 1,4-Cyclohexanedione-2,5-dicarboxylate in 12 Pure Solvents and Some Binary Solvent Mixtures at Different Temperatures	https://www.doi.org/10.1021/je2000292
Solubility and solution thermodynamics of baymol in pure and binary solvents at 298.15 K.	https://www.doi.org/10.1016/j.jct.2015.09.010
Determination and modelling of troxerutin solubility in eleven mono solvents and (1,4-dioxane + 2-propanol) binary solvents at temperatures from 278.15 to 323.15 K.	https://www.doi.org/10.1016/j.jct.2016.09.015
The Thermodynamic and Excess Properties of Trialkyl Substituted Imidazolium-Based Ionic Liquids with Tetrafluoroborate Anion in Binary Mixtures with Acetonitrile.	https://www.doi.org/10.1021/acs.jced.7b01006
Solubility and dissolution thermodynamic properties of ethyl 2-(4-phenoxyphenoxy) propionate in pure and binary solvent systems at 278.15–323.15 K	https://www.doi.org/10.1016/j.fluid.2013.05.038
Experimental determination and thermodynamic Models for Solubility Behavior and Data Modeling of L-Proline in Different Neat and Binary Mixtures of Ethanol + Water	https://www.doi.org/10.1021/acs.jced.8b00165
Solid-Liquid Equilibrium of 4-Dimethylaminopyridine-2-carboxamide in Pure and Binary Solvent Mixtures for T = (278.15-328.15) K.	https://www.doi.org/10.1021/acs.jced.9b00802
Assessment of Pyrrolidinium-Based Ionic Liquid for the Separation of Binary Mixtures Based on Activity Coefficients at Infinite Dilution.	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1386
Solubilities of six lithium salts in five nonaqueous solvents and in a few of their binary mixtures:	https://www.doi.org/10.1021/acs.jced.9b00341
Measurement and Correlation of Isothermal Vapor-Liquid Equilibrium Data for the System Acetonitrile + Water + Tetrapropylammonium Bromide	https://www.doi.org/10.1016/j.fluid.2017.12.034
Activity Coefficients at Infinite Dilution of Organic Compounds in 1-Ethyl-3-methylimidazolium Ethyl Sulfate	https://www.doi.org/10.1021/je0600855
Solubility of Fluconazole in Different Solvents:	https://www.doi.org/10.1021/je800658v
Thermodynamic Analysis of Solubility of Fluconazole in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00854
Densities and Excess Molar Volumes for Binary Mixtures of Ionic Liquid 1-Ethyl-3-methylimidazolium Ethylsulfate with Solvents:	https://www.doi.org/10.1016/j.fluid.2005.02.016
Measurement and correlation of the solubilities of 1,3,2-Dioxaphosphorinane, 2-methanol-a,5,5-trimethyl-a-phenyl-2-oxide in different pure solvents!	https://www.doi.org/10.1021/je1002237
Isothermodynamic and transport properties of acetonitrile + alkanediols liquid mixtures at different temperatures, experimental measurements and modeling.	https://www.doi.org/10.1016/j.fluid.2014.01.043
Partition Coefficients of Organic Compounds in Four New Tetraalkylammonium 5-(Methoxymethyl)furfuryl Phosphate Ionic Liquids Using Inverse GC	https://www.doi.org/10.1021/je500806u
Solid–Liquid Phase Equilibria between Imidazolium-Based Ionic Liquid 1-Butyl-3-methylimidazolium Octanoate + Protic Ionic Liquid in Molecular Solvents.	https://www.doi.org/10.1016/j.jct.2017.06.004
Solubility of Benzocaine in Five Organic Solvents:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Electrical Conductances of 1-Butyl-3-propylimidazolium Bromide and 1-Butyl-3-propylbenzimidazolium Bromide in Water, Methanol, and Acetonitrile at Several Temperatures.	https://www.doi.org/10.1021/je200637v
Isobaric vapor–liquid equilibrium for binary system of methanol and acetonitrile.	https://www.doi.org/10.1021/acs.jced.7b00585
Thermodynamic models for determination of the solid-liquid equilibrium of oxiracetam in acetonitrile + acetone binary solvent mixtures:	https://www.doi.org/10.1021/je4010917
Density, Viscosity, Speed of Sound, and Refractive Index of a Ternary Solution of Aspirin, Synthesized Trifunctional Ionic Liquid, and Acetonitrile at Different Temperatures	https://www.doi.org/10.1021/je100671v
Measurement of the solubility of Diloxanide in Pure and Binary Solvent Mixtures of Acetonitrile and Methanol at Different Temperatures and Discussion of the Effect of the Cosolvent.	https://www.doi.org/10.1016/j.fluid.2014.01.033
Determination and correlation for 1,8-dinitronaphthalene in (acetone + methanol), (toluene + methanol) and acetonitrile + methanol mixed solvents	https://www.doi.org/10.1021/je8006088
Solubility of carbazole in eight solvents within the temperature range (288.15 to 318.15) K.	https://www.doi.org/10.1021/je300339q
Measurement and correlation of solubility and dissolution properties of flunixin meglumine in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2013.08.022
	https://www.doi.org/10.1016/j.fluid.2012.12.003
	https://www.doi.org/10.1016/j.fluid.2015.08.009
	https://www.doi.org/10.1021/je5008372
	https://www.doi.org/10.1021/acs.jced.8b00601
	https://www.doi.org/10.1021/acs.jced.9b00353
	https://www.doi.org/10.1016/j.jct.2015.10.015
	https://www.doi.org/10.1016/j.jct.2016.09.011
	https://www.doi.org/10.1016/j.fluid.2015.06.026

Equilibrium solubility of dinitolmide in several neat solvents and binary aqueous co-solvent mixtures:	https://www.doi.org/10.1016/j.jct.2019.01.013
Solvent effect on solubility and preferential solvation analysis of nifuroxazide in binary aqueous solvent mixtures of 1,4-dioxane and 2-propanol	https://www.doi.org/10.1016/j.jct.2019.06.019
Determination and correlation of solubility of candesartan cilexetil in binary solvent mixtures: 2-dioxaphosphorinane-5,5-dimethyl -2-acetonitrile and isopropanol	https://www.doi.org/10.1016/j.fluid.2012.09.027
Thermodynamic analysis and correlation of solubility of candesartan cilexetil in binary solvent mixtures: 2-dioxaphosphorinane-5,5-dimethyl -2-acetonitrile and isopropanol	https://www.doi.org/10.1016/j.fluid.2013.09.023
(2,4,6-Trihydroxyphenyl)-1-propanone in Selected Solvents	https://www.doi.org/10.1021/acs.jced.8b00235
Solubility Behavior and Thermodynamic Functions of Trimetozidine Hydrochloride in 12 Selected Solvents	https://www.doi.org/10.1021/acs.jced.8b01080
Effect of Solvent Properties and Composition on the Solubility of Acetal (Ethanal Diethyl Acetal) in Monosolvents and Solvent Mixtures	https://www.doi.org/10.1016/j.jct.2012.01.020
Measurement of the potential of head space gas chromatography for VLE measurements:	https://www.doi.org/10.1016/j.jct.2015.02.002
Quantitative Molecular Thermodynamic models for determination of the solubility of DL-malic acid in methanol + water mixtures	https://www.doi.org/10.1016/j.jct.2003.09.005
Standard partial molar volumes of some electrolytes in ethylene glycol + water mixtures at 298.15 K	https://www.doi.org/10.1021/je3009792
Isobaric Vapor–Liquid Equilibrium for the Acetonitrile + Water System Containing Different Ionic Liquids at Atmospheric Pressure	https://www.doi.org/10.1021/acs.jced.7b00429
Solubility of Isoquercitrin in Pure and Binary Solvents: Molar Volumes, Excess Molar Quantities, Isentropic Compressibilities, Viscosity Deviations, and Activation Parameters of Flow	https://www.doi.org/10.1016/j.jct.2011.11.005
Solubility of organic solutes in 1-ethyl-3-methylimidazolium acetate.	https://www.doi.org/10.1021/acs.jced.8b00635
Determination of Activity Coefficients at Infinite Dilution of Solutes in 1-Ethyl-3-methylimidazolium Ethylsulfate Using Inverse Gas Chromatography	https://www.doi.org/10.1016/j.jct.2018.11.025
Trimethylsulfonium Methanesulfonate + Dimethyl Sulfoxide Binary Solvent Mixtures at T = 298.15 to 328.15 K and p = 0.1 MPa: Solubility, Preferential Solvation, and Thermodynamic Properties	https://www.doi.org/10.1021/je800515w
Solubility of Artemisinin in Seven Different Pure Solvents from (283.15 to 323.15) K	https://www.doi.org/10.1021/je100890f
Liquid–Liquid Equilibrium of Ternary Systems Containing 1-Butyl-3-methylimidazolium Tetrafluoroborate Ionic Liquid and Two Alcohols at 298.15 K	https://www.doi.org/10.1021/je400790d
Solubility of Artemisinin in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Methanol from (283.15 to 323.15) K	https://www.doi.org/10.1016/j.jct.2013.12.009
Solubility of (a,Z)-2-Amino-a-(methoxyimino)-4-thiazoleethanethioic Acid 2-Benzothiazolyl Ester in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Methanol	https://www.doi.org/10.1016/j.fluid.2010.01.002
Apparent Molar Volumes and Expansivities of Ionic Liquids Based on Physicochemical properties of two 1-alkyl-3-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide ionic liquids	https://www.doi.org/10.1021/je700026d
Viscosity, Density, and Speed of Sound for the Binary Mixtures of 1-Butyl-3-methylimidazolium Tetrafluoroborate with Methanol, Acetonitrile, and Dimethylsulfoxide at Different Temperatures	https://www.doi.org/10.1016/j.jct.2015.05.014
Viscosity, Density, and Speed of Sound for the Binary Mixtures of Formamide with Methanol or Acetonitrile or Dimethylsulfoxide at Different Temperatures	https://www.doi.org/10.1016/j.jct.2013.05.035
Measurements of activity coefficients at infinite dilution of organic solutes in 1,2-dimethoxyethane and 2-ethoxyethanol at 298.15 K	https://www.doi.org/10.1021/je7001946
Thermodynamic Properties of Ionic Liquids Composed of the Trifluoromethanesulfonate Anion and a Series of 1-Alkyl-3-methylimidazolium Cations	https://www.doi.org/10.1016/j.jct.2011.11.025
Measurements of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-3-methylimidazolium tetrafluoroborate by inverse gas chromatography and thermodynamic modelling of the solute–ionic liquid interaction	https://www.doi.org/10.1016/j.jct.2017.02.011
Density and Surface Tension of Binary Mixtures of Acetone + 1-Alkanol at 298.15 K	https://www.doi.org/10.1021/je050519g
Activity Coefficients at Infinite Dilution and Mixing Properties of Solutions:	https://www.doi.org/10.1021/je060305e
Solubilities of 3-Chlorophthalic Anhydride and 4-Chlorophthalic Anhydride in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.5b00526
Isothermal Vapor-Liquid Equilibrium Data for the Systems 1,4-Dioxane + Water + Tetrabutylammonium Bromide:	https://www.doi.org/10.1021/je0201727
Determination of Henry's Law Constants Using Internal Standards	https://www.doi.org/10.1021/je0301287
Probing molecular interactions of ionic liquid in industrially important solvents by solvatochromic probe.	https://www.doi.org/10.1016/j.tca.2013.02.022
Solubility of Tetrabutylammonium Nitrate in Organic Solvent Mixtures: Alcohol + Acetonitrile Solvent Mixtures at 298.2 K	https://www.doi.org/10.1021/je3010535
Spectroscopic and Thermodynamic Investigations of 1-methyl-4-(methylsulfonyl)benzene with Benchmark Values:	https://www.doi.org/10.1016/j.jct.2016.08.021
Thermodynamic Properties of Mixtures Containing Ionic Liquids: Activity Coefficients at Infinite Dilution of Solutes in 1-Propyl Boronic Acid 3-Alkylimidazolium Bromide and Tetrafluoroborate Ionic Liquids at 101.3 kPa	https://www.doi.org/10.1021/je060033f
Solubility of organic solvents in 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1021/je500068b
Solubilities of Lithium Tetrafluoroborate in 1-Propyl 3-Methylimidazolium Bromide and 1-Butyl 3-Methylimidazolium Bromide from (278.15 to 313.15) K	https://www.doi.org/10.1016/j.jct.2004.03.007
Vapor pressures, osmotic and activity coefficients for (1,3-Dioxolane + acetonitrile) and (1,3-Dioxolane + acetone) at T = (298.15 and 308.15) K	https://www.doi.org/10.1016/j.jct.2014.04.024
Determination of Activity Coefficients at Infinite Dilution for Solutes Dissolved in Ionic Liquids Using Gas Chromatography:	https://www.doi.org/10.1016/j.jct.2015.02.024
Solubility of Benzothiophene and Dibenzothiophene in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Methanol at Different Temperatures	https://www.doi.org/10.1021/je700234e
Measurement and correlation of the isobaric vapor–liquid equilibrium for methanol + acetonitrile + imidazolium based ionic liquids at 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2016.05.011
Thermodynamic Functions for Dissolution of 4-Hydroxybenzoic Acid in Pure and Mixed Organic and Aqueous Solvents and Water at the Ionic Liquid	https://www.doi.org/10.1021/acs.jced.8b00271
Solubility Study and Mixing Property of 3-Nitro-2-methylbenzoic Acid in 13 Pure Solvents from 288.15 to 333.15 K	https://www.doi.org/10.1016/j.jct.2012.01.019
1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1021/acs.jced.9b00497

Thermodynamic Properties of Mixtures Containing Precursors of Vitamin B5:	https://www.doi.org/10.1021/je020060f
Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures	https://www.doi.org/10.1021/je060311a
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.07.021
Solubility Determination and Correlation of Phenacetin in 14 Pure Solvents from 278.15 to 323.15 K and at Ambient Pressure:	https://www.doi.org/10.1021/je700560s
Solubility Measurement and Modeling of 1,8-dinitronaphthalene in nine organic solvents from T = 278.15 to 323.15 K:	https://www.doi.org/10.1021/je700645p
Solubility of Resorcinol Bicyclic Solvates in Nine Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00717
Solubility Measurement and Modeling of 2,4-dinitrophenylhydrazine in Ten Pure Solvents from T = 278.15 to 318.15 K:	https://www.doi.org/10.1021/je900838a
Solubility determination and thermodynamic modelling of 4-(3-nitrophenyl)-2-(4-methylphenyl)-3H-pyrazol-3-one in pure organic solvents:	https://www.doi.org/10.1016/j.jct.2016.12.028
Solubility Measurement and Modeling of 4-(3-nitrophenyl)-2-(4-methylphenyl)-3H-pyrazol-3-one in Pure Organic Solvents from T = 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00258
Determination of the solubility of 2-(3-nitrophenyl)-4-(4-methylphenyl)-3H-pyrazol-3-one in pure organic solvents from T = 278.15 to 318.15 K:	https://www.doi.org/10.1016/j.jct.2015.07.010
Solubility Measurement and Modeling of 4-(3-nitrophenyl)-2-(4-methylphenyl)-3H-pyrazol-3-one in Pure Organic Solvents from T = 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b00695
Solubility Measurement and Modeling of 4-Hydroxy-2-nitrobenzoic Acid in Ten Pure Solvents and Ethanol + Ethyl Acetate Binary Solvent Mixtures from T = 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00564
Solubility determination and thermodynamic modelling of 4-nitrophthalimide in different pure solvents:	https://www.doi.org/10.1016/j.jct.2016.10.029
Solubility Measurement and Modeling of 4-Nitrophthalimide in Different Organic Solvents from T = 278.15 to 318.15 K and Mixing Properties of Solutions:	https://www.doi.org/10.1016/j.jct.2016.08.007
Solubility determination and thermodynamic modelling of 4-nitrophthalonitrile in pure solvents:	https://www.doi.org/10.1021/acs.jced.8b00192
Solubility Measurement and Modeling of 4-Nitrophthalonitrile in Pure Organic Solvents from T = 278.15 to 318.15 K and Thermodynamic Properties of Solutions:	https://www.doi.org/10.1016/j.jct.2015.11.005
Determination of the solubility of 5-nitroisophthalic acid in pure solvents and binary mixtures from T = 278.15 to 318.15 K and correlation with the modified Apelblat equation:	https://www.doi.org/10.1021/je800056h
Measurement and correlation solubility of 5-nitroisophthalic acid in water + ethanol mixtures from T = 278.15 to 338.15 K:	https://www.doi.org/10.1016/j.fluid.2013.10.011
Measurement and correlation solubility of 5-nitroisophthalic acid in water + propan-2-ol, acetonitrile, ethyl acetate, and 1,4-dioxane mixtures from T = 278.15 to 338.15 K:	https://www.doi.org/10.1016/j.jct.2011.11.007
Solubility Measurement and Modeling of 5-Nitroisophthalic Acid in Pure and Mixture Organic Solvents:	https://www.doi.org/10.1021/je101146f
Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of Alcohols and Polar Organic Compounds:	https://www.doi.org/10.1016/j.fluid.2013.05.002
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with GLC measurements:	https://www.doi.org/10.1016/j.jct.2013.08.030
Thermodynamic Model Correlation and Solubility Measurement of Acetohexamide in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00931
Thermodynamics and activity coefficients at infinite dilution for organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2013.10.038
Thermodynamics and limiting activity coefficients measurements for organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Influence of interfacial adsorption using gas liquid chromatography:	https://www.doi.org/10.1016/j.jct.2017.07.027
Thermodynamic modeling of solubility and dissolution behavior of ribavirin form I in (ethanol + water) binary solvent mixtures:	https://www.doi.org/10.1021/acs.jced.8b01001
Experimental Solubility and Density Data, Molecular Dynamics, and Thermodynamic Studies of Ethanol, Isopropanol, and Their Binary Solvent Mixtures (Water + 1,4-Dioxane)	https://www.doi.org/10.1016/j.jct.2016.07.049
Solubility of isoxazolin acid in different solvent systems:	https://www.doi.org/10.1016/j.jct.2017.12.012
Performance of Polarizable Continuum Model and Joback Acree Model:	https://www.doi.org/10.1021/acs.jced.9b00220
Separation of binary mixtures based on gamma infinity data using Monosolvent and Binary Solvent Chromatography:	https://www.doi.org/10.1021/acs.jced.9b00385
Measurement and Correlation of Solubility of Hydrochlorothiazide in 12 Monosolvents and Binary Solvent Mixtures at 278.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.08.034
Solubility Determination, Correlation, and Solute–Solvent Molecular Interactions of 5-Aminotetrazole in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.7b00005
Solubility measurement and correlation of the form A of ibrutinib in organic solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.7b00851
Phase Behavior of Binary Mixtures Containing Succinic Acid or Its Esters:	https://www.doi.org/10.1021/acs.jced.8b01226
Determination and Thermodynamic Analysis of the Solubility of Limonin in Eight Organic Solvents and Ethyl Acetate + Ethanol Binary Solvents	https://www.doi.org/10.1016/j.jct.2016.10.040
Solubility measurement and thermodynamic functions of 4-nitrobenzaldehyde in different pure solvents:	https://www.doi.org/10.1021/je101191e
Non-Newtonian Shear-Thickening Densities and Volumetric Properties of (acetonitrile + an amide) binary solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2006.01.015
Protic Ionic Liquids and Their Mixtures with Water and Acetonitrile: Between Solution and Solvent	https://www.doi.org/10.1016/j.jct.2016.09.036
Solubility measurement, modelling and thermodynamic analysis of rivaroxaban polymorphs in pure solvents:	https://www.doi.org/10.1021/je1009812
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1016/j.jct.2015.02.023
Study of extraction behavior of ionic liquids using crush and chromatography:	https://www.doi.org/10.1016/j.fluid.2014.11.020
Accelerating Dehydration via Extractive Distillation Using Low Transition Temperature Mixtures as Entrainers:	https://www.doi.org/10.1021/acs.jced.8b00228

Thermodynamic Properties of Inorganic Salts in Nonaqueous Solvents. II. Apparent Molar Volumes and Compressibilities of Divalent Transition-Metal Perchlorates in Acetonitrile:	https://www.doi.org/10.1021/je0496950
Measurement and Correlation of Solubility of Loratadine in Different Pure Solvents and Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00721
Liquid–liquid equilibria and partitioning in organic aqueous systems:	https://www.doi.org/10.1016/j.fluid.2007.01.003
Measurement and Correlation of Solubility of Gatifloxacin in 12 Pure Solvents from 273.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00902
Solubilities of Organic Semiconductors and Nonsteroidal Anti-inflammatory Drugs in Pure and Mixed Organic Solvents: Measurement and Modeling with Hansen Solubility Parameter:	https://www.doi.org/10.1021/acs.jced.8b00536
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je1005517
High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid:	https://www.doi.org/10.1016/j.fluid.2017.06.001
Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory:	https://www.doi.org/10.1016/j.jct.2007.09.007
Densities, viscosities, excess molar volumes, and refractive indices of acetonitrile and 2-alkanols binary mixtures at different temperatures: Experimental results and application of the Prigogine Flory Patterson theory:	https://www.doi.org/10.1016/j.tca.2009.06.015
The dissolution behavior and apparent thermodynamic analysis of temozolomide in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2018.11.026
Equilibrium study and diversified models of drug Norfloxacin in eight pure organic and binary solvents at T = (278.15 to 328.15) K:	https://www.doi.org/10.1016/j.fluid.2016.12.012
Equilibrium Solubility, Model Correlation and Solvent Effect of Indole-3-acetic Acid in Twelve Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01265
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je900704b
Solubility Measurement, Correlation, and Molecular Interactions of 3-Methyl-6-nitroindazole in Different Neat Solvents and Mixed Solvents from T = 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b01256
Solubility evaluation and thermodynamic modeling of beta-lapachone in water and ten organic solvents at different temperatures:	https://www.doi.org/10.1016/j.fluid.2018.05.005
Solubility of Tetranitrodimerglycoluril (TNDGU) in Different Solvents at Temperatures between 293.15 K and 313.15 K:	https://www.doi.org/10.1021/je500392g
Application of trihexyltetradecylphosphonium dicyanamide ionic liquid for various types of separations problems: Activity coefficients at infinite dilution measurements utilizing GLC method:	https://www.doi.org/10.1016/j.fluid.2019.03.023
Solid-liquid equilibrium of diphenyl anilinophosphonate in the different solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.036
Guaiacol and its mixtures: new data and predictive models. Part 2: Gibbs energy of solvation:	https://www.doi.org/10.1016/j.fluid.2018.04.004
Solubility Measurement and Thermodynamic Properties of Levetiracetam in Pure and Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.8b00720
Solubility of Captopril in 2-Propanol, Acetone, Acetonitrile, Methyl Acetate, Ethyl Acetate, and Butyl Acetate:	https://www.doi.org/10.1021/je900401z
Solubility and Mixing Thermodynamics Properties of p-Toluenesulfonamide and o-Toluenesulfonamide in Seven Monosolvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00714
Solubility Modeling and Mixing Properties for Benzoin in Different Monosolvents and Solvent Mixtures at the Temperature Range from 273.15 to 313.15 K:	https://www.doi.org/10.1021/acs.jced.7b00743
Solubility and Data Correlation of Isoniazid in Different Pure and Binary Mixed Solvent Systems from 283.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00785
Solubility, Solubility Modeling, and Antisolvent Precipitation of 1,3-Bis(9-carbazolyl)benzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00432
Thermophysical approach to understand the nature of molecular interactions and structural factor between methyl isobutyl ketone and organic solvents mixtures:	https://www.doi.org/10.1016/j.jct.2017.06.008
Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T = 298.15 K:	https://www.doi.org/10.1021/je0601098
Thermodynamic Modeling and Solubility Measurement of Cetirizine Hydrochloride and Deferiprone in Pure Solvents of Acetonitrile, Ethanol, Acetic Acid, Sulfolane, and Ethyl Acetate and Their Mixtures:	https://www.doi.org/10.1021/acs.jced.9b00620
Solubility and Data Correlation of b-Arbutin in Different Monosolvents from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00703
Densities and Viscosities of Ternary N,N'-Bis(2-pyridylmethylidene)-1,2-diiminoethane Schiff Base + Imidazolium Based Ionic Liquids + Acetonitrile Solutions at T = (298.15 to 313.15) K:	https://www.doi.org/10.1021/je501054c
Determination and thermodynamic modelling for 4-nitropyrazole solubility in (methanol + water), (ethanol + water) and (acetonitrile + water) binary solvent mixtures from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.023
Abraham model correlations for describing the thermodynamic properties of solute transfer into pentyl acetate based on headspace chromatographic and solubility measurements:	https://www.doi.org/10.1016/j.jct.2018.05.003
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.jct.2012.03.005
Solubility of 3,7-Dinitro-1,3,5,7-tetraazabicyclo[3.3.1]nonane in Ethanenitrile, Methanol, 1,1-Dichloroethane, Dimethyl Sulfoxide, Acetone, and Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.5b00033
Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the 1-butyl-3-methylimidazolium perchlorate:	https://www.doi.org/10.1016/j.jct.2017.07.012
Solubility of 1-methyl-4-nitropyrazole in seventeen pure solvents at temperatures from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2018.05.023
Solubility measurement and thermodynamic functions of dehydroepiandrosterone acetate in different solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.01.023
Application of the Extended Langmuir Model for the Determination of Solubility of 1-Propanol in Acetonitrile:	https://www.doi.org/10.1007/s10765-009-0581-7
Phase Equilibria in Systems of Morpholine, Acetonitrile, and n-Alkanes:	https://www.doi.org/10.1021/acs.jced.5b00175
High Pressure Phase Equilibria of the Related Substances in the Limonene Oxidation in Supercritical CO2:	https://www.doi.org/10.1021/je020150k
Solubility Determination and Thermodynamic Modeling of 2-Mercaptobenzimidazole in 12 Solvents from T = 278.15 K to T = 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00190
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
In situ measurement of liquid-liquid equilibria by medium field nuclear magnetic resonance:	https://www.doi.org/10.1016/j.fluid.2017.01.027

Solubility of 3,7,9,11-Tetraoxo-2,4,6,8,10-pentaaza[3.3.3] Propellane (TOPAP) in Different Pure Solvents at Temperatures between 278.15 and 313.15 K and Mixed Organic Solvents from (283.15 to 323.15) K: 1-Butyl-3-methylimidazolium Nitrate Using Gas–Liquid Chromatography: Dissolution behavior and thermodynamic properties of lapatinib ditosylate monohydrate in fourteen organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions: Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Nitrate Ionic Liquid at T = (313.15 to 343.15) K: Solvent Dependencies of Activity Coefficients at Infinite Dilution for Solutes Dissolved in 1-Butyl-3-methylimidazolium Nitrate	https://www.doi.org/10.1021/acs.jced.6b00349 https://www.doi.org/10.1016/j.fluid.2019.01.005 https://www.doi.org/10.1021/je200050q https://www.doi.org/10.1016/j.jct.2017.02.008

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
ea:	Electron affinity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
nfpas:	NFPA Safety Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices

sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
svapt:	Entropy of vaporization at a given temperature
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tcondl:	Liquid thermal conductivity
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter