Trichloromethane

Other names:	CHCl3
	Chloroform
	Chloroforme
	Cloroformio
	F 20
	Formyl trichloride
	Freon 20
	Methane trichloride
	Methane, trichloro-
	Methenyl trichloride
	Methyl trichloride
	NCI-C02686
	NSC 77361
	R 20
	R 20 (refrigerant)
	R-20
	Rcra waste number U044
	Trichloormethaan
	Trichlormethan
	Trichloroform
	Triclorometano
	UN 1888
Inchi:	InChI=1S/CHCl3/c2-1(3)4/h1H
InchiKey:	HEDRZPFGACZZDS-UHFFFAOYSA-N
Formula:	CHCl3
SMILES:	ClC(Cl)Cl
Mol. weight [g/mol]:	119.38
CAS:	67-66-3

Physical Properties

Property code	Value	Unit	Source
af	0.2180		KDB
chl	-474.00 ± 8.40	kJ/mol	NIST Webbook
chl	-473.21	kJ/mol	NIST Webbook
dm	1.10	debye	KDB
ea	0.62 ± 0.16	eV	NIST Webbook

ea	1.76 ± 0.05	eV	NIST Webbook
gf	-68.58	kJ/mol	KDB
gyrad	3.1780		KDB
hf	-101.30	kJ/mol	KDB
hf	-102.90 ± 2.50	kJ/mol	NIST Webbook
hfl	-134.30	kJ/mol	NIST Webbook
hfl	-134.10 ± 2.50	kJ/mol	NIST Webbook
hfus	7.41	kJ/mol	Joback Method
hvap	31.40	kJ/mol	NIST Webbook
hvap	31.32 ± 0.08	kJ/mol	NIST Webbook
hvap	30.50 ± 0.42	kJ/mol	NIST Webbook
hvap	31.10	kJ/mol	NIST Webbook
ie	11.48	eV	NIST Webbook
ie	11.42 ± 0.03	eV	NIST Webbook
ie	11.50	eV	NIST Webbook
ie	11.40	eV	NIST Webbook
ie	11.37 ± 0.02	eV	NIST Webbook
ie	11.48	eV	NIST Webbook
ie	11.30	eV	NIST Webbook
ie	11.41 ± 0.02	eV	NIST Webbook
ie	11.37 ± 0.02	eV	NIST Webbook
ie	11.50	eV	NIST Webbook
log10ws	-1.17		Aqueous Solubility Prediction Method
log10ws	-1.17		Estimated Solubility Method
logp	1.986		Crippen Method
mcvol	61.670	ml/mol	McGowan Method
nfpah	%!d(float64=2)		KDB
pc	5328.68 ± 10.13	kPa	NIST Webbook
pc	5470.00	kPa	KDB
pc	5328.68 ± 6.07	kPa	NIST Webbook
rhoc	490.64 ± 3.58	kg/m3	NIST Webbook
rhoc	495.42 ± 2.39	kg/m3	NIST Webbook
rhoc	458.41 ± 2.39	kg/m3	NIST Webbook
rinpol	607.90		NIST Webbook
rinpol	621.10		NIST Webbook
rinpol	610.00		NIST Webbook
rinpol	614.80		NIST Webbook
rinpol	625.80		NIST Webbook
rinpol	604.00		NIST Webbook
rinpol	606.00		NIST Webbook
rinpol	604.00		NIST Webbook
rinpol	606.80		NIST Webbook
rinpol	610.00		NIST Webbook

rinpol	604.00	NIST Webbook
rinpol	608.00	NIST Webbook
rinpol	602.00	NIST Webbook
rinpol	582.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	568.00	NIST Webbook
rinpol	595.20	NIST Webbook
rinpol	616.00	NIST Webbook
rinpol	628.40	NIST Webbook
rinpol	587.00	NIST Webbook
rinpol	622.00	NIST Webbook
rinpol	622.00	NIST Webbook
rinpol	613.00	NIST Webbook
rinpol	598.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	616.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	621.70	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	590.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	600.00	NIST Webbook
rinpol	561.00	NIST Webbook
rinpol	569.00	NIST Webbook
rinpol	630.00	NIST Webbook
rinpol	605.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	600.00	NIST Webbook
rinpol	588.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	616.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	617.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	611.00	NIST Webbook
rinpol	629.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	603.00	NIST Webbook

rinpol	620.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	603.00	NIST Webbook
rinpol	619.00	NIST Webbook
rinpol	620.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	611.00	NIST Webbook
rinpol	601.00	NIST Webbook
rinpol	606.00	NIST Webbook
rinpol	603.00	NIST Webbook
rinpol	609.00	NIST Webbook
rinpol	616.00	NIST Webbook
rinpol	595.00	NIST Webbook
rinpol	605.00	NIST Webbook
rinpol	605.00	NIST Webbook
rinpol	598.00	NIST Webbook
rinpol	617.20	NIST Webbook
rinpol	605.00	NIST Webbook
rinpol	618.20	NIST Webbook
ripol	1024.00	NIST Webbook
ripol	1037.00	NIST Webbook
ripol	1034.00	NIST Webbook
ripol	1030.11	NIST Webbook
ripol	1028.95	NIST Webbook
ripol	1027.69	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1024.00	NIST Webbook
ripol	1021.00	NIST Webbook
ripol	1020.00	NIST Webbook
ripol	1022.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1026.00	NIST Webbook
ripol	1013.00	NIST Webbook
ripol	1028.00	NIST Webbook
ripol	1010.00	NIST Webbook
ripol	1020.00	NIST Webbook
ripol	1021.00	NIST Webbook
ripol	1022.00	NIST Webbook
ripol	1022.00	NIST Webbook
ripol	1018.00	NIST Webbook

ripol	1020.00		NIST Webbook
ripol	1023.00		NIST Webbook
ripol	1022.00		NIST Webbook
ripol	999.00		NIST Webbook
ripol	1038.00		NIST Webbook
ripol	1010.00		NIST Webbook
ripol	1045.00		NIST Webbook
ripol	1007.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1017.00		NIST Webbook
ripol	1025.00		NIST Webbook
ripol	1015.00		NIST Webbook
ripol	1016.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1018.00		NIST Webbook
ripol	1020.00		NIST Webbook
ripol	992.00		NIST Webbook
ripol	1014.00		NIST Webbook
ripol	1014.00		NIST Webbook
ripol	1000.00		NIST Webbook
ripol	1000.00		NIST Webbook
ripol	1024.00		NIST Webbook
ripol	1037.00		NIST Webbook
ripol	1027.69		NIST Webbook
ripol	1010.00		NIST Webbook
ripol	1037.00		NIST Webbook
tb	334.30 ± 0.50	K	NIST Webbook
tb	334.35	K	Separation of azeotrope (2,2,3,3-tetrafluoro-1-propanol + water): Isobaric vapour-liquid phase equilibrium measurements and azeotropic distillation
tb	334.45 ± 0.60	K	NIST Webbook
tb	334.35 ± 0.20	K	NIST Webbook
tb	334.32	K	KDB
tb	335.00	K	The role of organic diluents in the aspects of equilibrium, kinetics and thermodynamic model for silver ion extraction using an extractant D2EHPA

tb	334.10	K	Phase diagrams of (vapour + liquid) equilibrium for binary mixtures of a,a,a-trifluorotoluene with ethanol, or benzene, or chloroform at pressure 101.4 kPa
tb	334.40 ± 0.30	K	NIST Webbook
tb	334.12	K	Isobaric Vapor-Liquid Phase Equilibrium Measurements for Allyl Alcohol with Chloroform, Ethyl Acetate, and Methyl Propionate at 101.3 kPa
tb	334.30	K	NIST Webbook
tb	334.37 ± 0.15	K	NIST Webbook
tb	334.45 ± 0.50	K	NIST Webbook
tb	334.10 ± 0.30	K	NIST Webbook
tb	334.35 ± 0.06	K	NIST Webbook
tb	334.34 ± 0.25	K	NIST Webbook
tb	334.39 ± 0.30	K	NIST Webbook
tb	334.15 ± 0.50	K	NIST Webbook
tb	334.35 ± 0.10	K	NIST Webbook
tb	334.24 ± 0.15	K	NIST Webbook
tb	333.95 ± 0.40	K	NIST Webbook
tb	334.42 ± 0.35	K	NIST Webbook
tb	334.42 ± 0.30	K	NIST Webbook
tb	334.35 ± 0.30	K	NIST Webbook
tb	334.35 ± 0.30	K	NIST Webbook
tb	334.35 ± 0.30	K	NIST Webbook
tb	334.30 ± 0.20	K	NIST Webbook
tb	334.30 ± 0.50	K	NIST Webbook
tb	334.45 ± 0.50	K	NIST Webbook
tb	334.20 ± 1.00	K	NIST Webbook
tb	334.50 ± 0.20	K	NIST Webbook
tb	334.22 ± 0.50	K	NIST Webbook
tb	334.27 ± 0.30	K	NIST Webbook
tb	334.30 ± 0.25	K	NIST Webbook
tb	334.30 ± 0.50	K	NIST Webbook
tb	334.23 ± 0.06	K	NIST Webbook
tb	334.30 ± 0.30	K	NIST Webbook
tb	334.40 ± 0.50	K	NIST Webbook
tb	334.37 ± 0.20	K	NIST Webbook
tb	334.20 ± 0.30	K	NIST Webbook
tb	334.22 ± 0.20	K	NIST Webbook
tb	334.35 ± 0.30	K	NIST Webbook
tb	334.30 ± 0.20	K	NIST Webbook

tc	536.40	K	KDB
tf	209.64 ± 0.10	K	NIST Webbook
tf	210.46 ± 0.20	K	NIST Webbook
tf	210.00 ± 2.00	K	NIST Webbook
tf	209.50	K	KDB
tf	212.15 ± 1.00	K	NIST Webbook
tf	209.90 ± 0.30	K	NIST Webbook
tf	211.59 ± 0.50	K	NIST Webbook
tf	210.00 ± 2.00	K	NIST Webbook
tf	209.50 ± 0.70	K	NIST Webbook
tf	209.95	K	Aqueous Solubility Prediction Method
tf	211.05 ± 0.30	K	NIST Webbook
tf	209.65 ± 0.20	K	NIST Webbook
tf	209.73 ± 0.01	K	NIST Webbook
tf	210.15 ± 1.00	K	NIST Webbook
tf	209.65 ± 0.20	K	NIST Webbook
tf	209.61 ± 0.08	K	NIST Webbook
tt	209.61 ± 0.20	K	NIST Webbook
vc	0.239	m3/kmol	KDB
volm	8.07e-05	m3/mol	Excess Gibbs energies and volumes of the ternary system chloroform + tetrahydrofuran + cyclohexane at 298.15 K
zc	0.2931300		KDB
zra	0.28		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	75.56	J/mol×K	400.12	Joback Method
cpg	79.40	J/mol×K	466.12	Joback Method
cpg	81.13	J/mol×K	499.12	Joback Method
cpg	71.18	J/mol×K	334.13	Joback Method
cpg	82.74	J/mol×K	532.11	Joback Method
cpg	77.54	J/mol×K	433.12	Joback Method
cpg	73.44	J/mol×K	367.13	Joback Method
cpl	114.25	J/mol×K	298.15	NIST Webbook
cpl	114.35	J/mol×K	298.15	NIST Webbook
cpl	117.10	J/mol×K	298.00	NIST Webbook
cpl	113.20	J/mol×K	298.15	NIST Webbook
cpl	139.70	J/mol×K	303.60	NIST Webbook

cpl	113.00	J/mol×K	298.10	NIST Webbook
cpl	113.73	J/mol×K	298.15	NIST Webbook
cpl	114.26	J/mol×K	298.15	NIST Webbook
cpl	116.20	J/mol×K	293.00	NIST Webbook
cpl	113.40	J/mol×K	300.00	NIST Webbook
cpl	114.32	J/mol×K	298.15	NIST Webbook
cpl	115.50	J/mol×K	298.15	NIST Webbook
cpl	115.50	J/mol×K	293.20	NIST Webbook
cpl	114.00	J/mol×K	303.20	NIST Webbook
cpl	116.70	J/mol×K	303.00	NIST Webbook
cpl	114.18	J/mol×K	298.00	NIST Webbook
cpl	115.50	J/mol×K	298.15	NIST Webbook
cpl	113.85	J/mol×K	298.15	NIST Webbook
dvisc	0.0007130	Pa×s	273.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]
dvisc	0.0005200	Pa×s	303.15	Studies of viscosities of dilute solutions of alkylamine in non-electrolyte solvents. II. Haloalkanes and other polar solvents
dvisc	0.0004710	Pa×s	313.15	Studies of viscosities of dilute solutions of alkylamine in non-electrolyte solvents. II. Haloalkanes and other polar solvents
dvisc	0.0004910	Pa×s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K

dvisc	0.0005340	Pa×s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Octan-1-ol and Decan-1-ol at (303.15 and 313.15) K	
dvisc	0.0004910	Pa×s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Octan-1-ol and Decan-1-ol at (303.15 and 313.15) K	
dvisc	0.0005130	Pa×s	308.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K	
dvisc	0.0006000	Pa×s	293.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	
dvisc	0.0005760	Pa×s	298.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	
dvisc	0.0005520	Pa×s	303.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	

dvisc	0.0005090	Pa×s	313.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	
dvisc	0.0004910	Pa×s	318.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	
dvisc	0.0005560	Pa×s	298.15	Densities, viscosities, and ultrasonic velocity studies of binary mixtures of trichloromethane with methanol, ethanol, propan-1-ol, and butan-1-ol at T=(298.15 and 308.15) K	
dvisc	0.0005340	Pa×s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Propan-1-ol and Butan-1-ol at (303.15 and 313.15) K	
dvisc	0.0006510	Pa×s	283.15	Density, Viscosity, Vapor-Liquid Equilibrium, and Excess Molar Enthalpy of [Chloroform + Methyl tert-Butyl Ether]	
hfust	8.80	kJ/mol	209.60	NIST Webbook	
hfust	8.80	kJ/mol	209.60	NIST Webbook	
hvapt	30.40	kJ/mol	374.50	NIST Webbook	
hvapt	28.90	kJ/mol	445.50	NIST Webbook	
hvapt	30.10	kJ/mol	501.00	NIST Webbook	
hvapt	30.90	kJ/mol	320.50	NIST Webbook	
hvapt	29.71	kJ/mol	334.30	KDB	
hvapt	32.50	kJ/mol	296.50	NIST Webbook	
hvapt	35.00	kJ/mol	274.50	NIST Webbook	

hvapt	29.24	kJ/mol	334.30	NIST Webbook
hvapt	30.80	kJ/mol	366.50	NIST Webbook
hvapt	31.80	kJ/mol	248.00	NIST Webbook
kvisc	0.0000003	m2/s	313.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Pentan-1-ol, Hexan-1-ol, and Heptan-1-ol at (303.15 and 313.15) K
kvisc	0.0000004	m2/s	303.15	Densities, Viscosities, and Ultrasonic Velocity Studies of Binary Mixtures of Chloroform with Pentan-1-ol, Hexan-1-ol, and Heptan-1-ol at (303.15 and 313.15) K
pvap	61.85	kPa	320.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	32.19	kPa	303.15	Density, viscosity, isothermal (vapour + liquid) equilibrium, excess molar volume, viscosity deviation, and their correlations for chloroform + methyl isobutyl ketone binary system
pvap	101.30	kPa	334.35	Separation of azeotrope (2,2,3,3-tetrafluoro-1-propanol + water): Isobaric vapour-liquid phase equilibrium measurements and azeotropic distillation

pvap	32.41	kPa	303.15	Total Vapor Pressure Measurements for 2-Ethoxyethanol with Carbon Tetrachloride, Chloroform, and Dichloromethane at 303.15 K
pvap	32.19	kPa	303.15	Density, Viscosity, Vapor-Liquid Equilibrium, Excess Molar Volume, Viscosity Deviation, and Their Correlations for the Chloroform + 2-Butanone Binary System
pvap	34.81	kPa	305.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	42.47	kPa	310.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	51.43	kPa	315.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	129.28	kPa	342.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa

pvap	103.50	kPa	335.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	87.71	kPa	330.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	73.88	kPa	325.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	121.45	kPa	340.00	Isobaric Vapor-Liquid and Liquid-Liquid Equilibria for Chloroform + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	101.30	kPa	334.12	Isobaric Vapor-Liquid Phase Equilibrium Measurements for Allyl Alcohol with Chloroform, Ethyl Acetate, and Methyl Propionate at 101.3 kPa
rfi	1.44100		303.15	Densities, speeds of sound, isentropic compressibilities, refractive indexes, and viscosities of tetrahydrofuran with haloalkane or alkyl ethanoate at T = 303.15 K
rfi	1.44760		293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents

rfi	1.45000	283.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.44390	293.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.44280	298.15	Vapour pressure and excess Gibbs energy of binary 1,2-dichloroethane + cyclohexanone, chloroform + cyclopentanone and chloroform + cyclohexanone mixtures at temperatures from 298.15 to 318.15 K
rfi	1.44200	298.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rfi	1.44300	298.15	Liquid-Liquid Equilibrium, Solid-Liquid Equilibrium, Densities, and Refractivity of a Water, Chloroform, and Acetylacetone Mixture
rfi	1.44290	298.15	Study of Vapor-Liquid Equilibrium for Binary Mixtures (Chloroform + 2,2,2-Trifluoroethanol) and (r,r,r-Trifluorotoluene + 2,2,2-Trifluoroethanol) at Pressure 102 kPa

rfi	1.44710		288.15	Partial Molar Volumes of N,N'-1,2-Ethyl-bis(salicyladimine) Schiff Base (Salen) in Organic Solvents at T = (283.15 to 318.15) K
rhol	1479.21	kg/m3	298.15	Volumetric behaviour of binary mixtures of (trichloromethane + amines) at temperatures between T = (288.15 and 303.15) K at p = 0.1 MPa
rhol	1450.41	kg/m3	313.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)
rhol	1483.02	kg/m3	293.20	Isobaric vapour liquid equilibria for binary mixtures of 1,2-dibromoethane with 1,2-dichloroethane, trichloromethane, and 1,1,2,2-tetrachloroethane at atmospheric pressure
rhol	1498.08	kg/m3	288.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS

rhol	1488.64	kg/m3	293.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS	
rhol	1479.15	kg/m3	298.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS	
rhol	1469.61	kg/m3	303.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS	
rhol	1460.03	kg/m3	308.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS	
rhol	1450.41	kg/m3	313.15	Volumetric properties of the ternary system ethanol + chloroform + benzene at temperature range (288.15 313.15) K: Experimental data, correlation and prediction by cubic EOS	

rhol	1480.34	kg/m3	298.20	The choice of solvent and liquid liquid equilibrium for ternary water + 2-methylaziridine + chloroform system: Experimental data and modeling	
rhol	1478.75	kg/m3	298.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	1469.16	kg/m3	303.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	1460.01	kg/m3	308.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	1450.39	kg/m3	313.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	

rhol	1440.03	kg/m3	318.15	Densities, viscosities, and refractive indices of binary and ternary mixtures of methanol, acetone, and chloroform at temperatures from (298.15-318.15) K and ambient pressure	
rhol	1498.14	kg/m3	288.15	Volumetric behaviour of binary mixtures of (trichloromethane + amines) at temperatures between T = (288.15 and 303.15) K at p = 0.1 MPa	
rhol	1488.70	kg/m3	293.15	Volumetric behaviour of binary mixtures of (trichloromethane + amines) at temperatures between T = (288.15 and 303.15) K at p = 0.1 MPa	
rhol	1460.03	kg/m3	308.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	1469.63	kg/m3	303.15	Volumetric behaviour of binary mixtures of (trichloromethane + amines) at temperatures between T = (288.15 and 303.15) K at p = 0.1 MPa	

rhol	1475.30	kg/m3	303.00	Thermo-physical properties of 1,3,4-oxadiazole derivatives in pure solvents	
rhol	1470.80	kg/m3	308.00	Thermo-physical properties of 1,3,4-oxadiazole derivatives in pure solvents	
rhol	1465.80	kg/m3	313.00	Thermo-physical properties of 1,3,4-oxadiazole derivatives in pure solvents	
rhol	1475.30	kg/m3	303.00	Thermo-acoustical studies of 1,3,4-oxadiazole as binary mixture at three different temperatures	
rhol	1470.80	kg/m3	308.00	Thermo-acoustical studies of 1,3,4-oxadiazole as binary mixture at three different temperatures	
rhol	1465.80	kg/m3	313.00	Thermo-acoustical studies of 1,3,4-oxadiazole as binary mixture at three different temperatures	
rhol	1488.70	kg/m3	293.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	

rhol	1469.69	kg/m3	303.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	
rhol	1450.39	kg/m3	313.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	
rhol	1429.89	kg/m3	323.15	Excess volumes, Viscosities, and Excess Gibbs Energy of Activation for Viscous Flow, for binary and ternary mixtures 1- propanol + N-N dimethylformamid + chloroform at different temperatures	
rhol	1470.50	kg/m3	303.15	Viscosity and Density for Binary Mixtures of Carbon Tetrachloride + Chloroform, Carbon Tetrachloride + Dichloromethane, and Chloroform + Dichloromethane and One Ternary Mixture of Chloroform + 1:1 (Carbon Tetrachloride + Dichloromethane) at 303.15 K	

rhol	1498.08	kg/m3	288.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	
rhol	1488.64	kg/m3	293.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	
rhol	1479.15	kg/m3	298.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	
rhol	1469.61	kg/m3	303.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	
rhol	1460.03	kg/m3	308.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	

rhol	1450.41	kg/m3	313.15	Volumetric Properties of the Binary Methanol + Chloroform and Ternary Methanol + Chloroform + Benzene Mixtures at (288.15, 293.15, 298.15, 303.15, 308.15, and 313.15) K	
rhol	1473.42	kg/m3	298.15	Volumetric Study for the Binary Nitromethane with Chloroalkane Mixtures at Temperatures in the Range (298.15 to 318.15) K	
rhol	1454.33	kg/m3	308.15	Volumetric Study for the Binary Nitromethane with Chloroalkane Mixtures at Temperatures in the Range (298.15 to 318.15) K	
rhol	1434.98	kg/m3	318.15	Volumetric Study for the Binary Nitromethane with Chloroalkane Mixtures at Temperatures in the Range (298.15 to 318.15) K	
rhol	1487.63	kg/m3	293.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform	
rhol	1478.16	kg/m3	298.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform	

rhol	1468.61	kg/m3	303.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	1469.61	kg/m3	303.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)
rhol	1498.08	kg/m3	288.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K
rhol	1488.64	kg/m3	293.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K

rhol	1479.15	kg/m3	298.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K	
rhol	1469.61	kg/m3	303.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K	
rhol	1460.03	kg/m3	308.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K	
rhol	1479.15	kg/m3	298.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	

rhol	1491.86	kg/m3	288.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclohexanone with Chloroalkanes at Temperatures between (288.15 and 318.15) K
rhol	1473.16	kg/m3	298.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclohexanone with Chloroalkanes at Temperatures between (288.15 and 318.15) K
rhol	1454.07	kg/m3	308.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclohexanone with Chloroalkanes at Temperatures between (288.15 and 318.15) K
rhol	1434.72	kg/m3	318.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclohexanone with Chloroalkanes at Temperatures between (288.15 and 318.15) K
rhol	1492.31	kg/m3	288.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclopentanone with Chloroalkanes at T = (288.15, 298.15, 308.15, and 318.15) K

rhol	1473.43	kg/m3	298.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclopentanone with Chloroalkanes at T = (288.15, 298.15, 308.15, and 318.15) K	
rhol	1454.33	kg/m3	308.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclopentanone with Chloroalkanes at T = (288.15, 298.15, 308.15, and 318.15) K	
rhol	1434.97	kg/m3	318.15	Densities and Excess Molar Volumes of the Binary Mixtures of Cyclopentanone with Chloroalkanes at T = (288.15, 298.15, 308.15, and 318.15) K	
rhol	1488.64	kg/m3	293.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	
rhol	1498.08	kg/m3	288.15	Temperature Dependence of Densities and Excess Molar Volumes of the Ternary Mixture (1-Butanol + Chloroform + Benzene) and its Binary Constituents (1-Butanol + Chloroform and 1-Butanol + Benzene)	

rhol	1459.00	kg/m3	308.15	Speed of sound, density and related thermodynamic excess properties of binary mixtures of butan-2-one with C1-C4 nalkanols and chloroform
rhol	1489.00	kg/m3	293.00	KDB
rhol	1450.41	kg/m3	313.15	Densities and Excess Molar Volumes of the Ternary Mixture 2-Butanol + Chloroform + Benzene and Binary Mixtures 2-Butanol + Chloroform, or + Benzene over the Temperature Range (288.15 to 313.15) K
speedsl	950.45	m/s	308.15	Speeds of sound, isentropic compressibilities and refractive indices for some binary mixtures of nitromethane with chloroalkane at temperatures from 298.15 to 318.15 K. Comparison with theories
speedsl	984.42	m/s	298.15	Speeds of sound, isentropic compressibilities and refractive indices for some binary mixtures of nitromethane with chloroalkane at temperatures from 298.15 to 318.15 K. Comparison with theories
speedsl	916.63	m/s	318.15	Speeds of sound, isentropic compressibilities and refractive indices for some binary mixtures of nitromethane with chloroalkane at temperatures from 298.15 to 318.15 K. Comparison with theories

srf	0.03	N/m	293.15	Investigation of Surface Properties and Solubility of 1-Vinyl-3-alkyl/Esterimidazolium Halide Ionic Liquids by Density Functional Methods
srf	0.03	N/m	298.20	KDB

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	334.25	K	96.60	Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.43630e+01
Coeff. B	-2.83926e+03
Coeff. C	-4.29370e+01
Temperature range (K), min.	244.65
Temperature range (K), max.	536.40

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	8.11439e+01
Coeff. B	-6.35114e+03
Coeff. C	-1.00709e+01
Coeff. D	9.12761e-06

Temperature range (K), min.	209.63
Temperature range (K), max.	536.40

Datasets

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
303.15	101.30	0.0005120

Reference *https://www.doi.org/10.1016/j.tca.2004.07.014*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
303.15	101.30	0.0005033

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Legend

af:	Acentric Factor
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
ea:	Electron affinity
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density

rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
volm:	Molar Volume
zc:	Critical Compressibility
zra:	Rackett Parameter