

Sucrose

Other names:

(+)-Sucrose
(2R,3R,4S,5S,6R)-2-(((2S,3S,4S,5R)-3,4-dihydroxy-2,5-bis(hydroxymethyl)tetrahydrofuran-2-ylidene)-4,5-dihydroxy-2-methyl-6-oxo-1,3-dioxane-6-carboxylate)
(«alpha»-D-Glucosido)-«beta»-D-fructofuranoside
(Â«alphaÂ»-D-Glucosido)-Â«betaÂ»-D-fructofuranoside
.alpha.-trehalose
.beta.-D-fructofuranosyl .alpha.-D-glucopyranoside
4-O-.beta.-D-galactopyranosyl-D-glucose
Amerfond
Beet sugar
Cane sugar
Confectioner's sugar
D-(+)-Saccharose
D-(+)-Sucrose
D-(+)-lactose
D-Sucrose
D-trehalose
Fructofuranoside, «alpha»-D-glucopyranosyl, «beta»-D-glucopyranosyl
Fructofuranoside, Â«alphaÂ»-D-glucopyranosyl, Â«betaÂ»-D-glucopyranosyl
Glucopyranoside, «beta»-D-fructofuranosyl, «alpha»-D-fructofuranosyl
Glucopyranoside, Â«betaÂ»-D-fructofuranosyl, Â«alphaÂ»-D-fructofuranosyl
Granulated sugar
Microse
NCI-C56597
NSC 406942
Rock candy
Saccharose
Saccharum
Sugar
Table sugar
White sugar
alpha,alpha-trehalose
lactose
«alpha»-D-Glucopyranoside, «beta»-D-fructofuranosyl
«alpha»-D-Glucopyranosyl «beta»-D-fructofuranoside
«beta»-D-Fructofuranoside, «alpha»-D-glucopyranosyl
«beta»-D-Fructofuranosyl «alpha»-D-glucopyranoside
Â«alphaÂ»-D-Glucopyranoside, Â«betaÂ»-D-fructofuranosyl
Â«alphaÂ»-D-Glucopyranosyl Â«betaÂ»-D-fructofuranoside
Â«betaÂ»-D-Fructofuranoside, Â«alphaÂ»-D-glucopyranosyl
Â«betaÂ»-D-Fructofuranosyl Â«alphaÂ»-D-glucopyranoside

Inchi:

InchiKey:

Formula:

SMILES:

Mol. weight [g/mol]:

CAS:

InChI=1S/C12H22O11/c13-1-4-6(16)8(18)9(19)11(21-4)23-12(3-15)10(20)7(17)5(2-14)22

CZMRCDWAGMRECN-SFOFJGFUSA-N

C12H22O11

OCC1OC(OC2(CO)OC(CO)C(O)C2O)C(O)C(O)C1O

342.30

57-50-1

Physical Properties

Property code	Value	Unit	Source
chs	-5664.38 ± 0.69	kJ/mol	NIST Webbook
chs	-5637.40 ± 1.70	kJ/mol	NIST Webbook
chs	-5644.17	kJ/mol	NIST Webbook
chs	-5643.40 ± 1.80	kJ/mol	NIST Webbook
gf	-1320.10	kJ/mol	Joback Method
hf	-1917.41	kJ/mol	Joback Method
hfs	-2221.20	kJ/mol	NIST Webbook
hfus	63.65	kJ/mol	Joback Method
hvap	184.54	kJ/mol	Joback Method
log10ws	0.79		Estimated Solubility Method
log10ws	0.79		Aqueous Solubility Prediction Method
logp	-5.396		Crippen Method
mcvol	222.790	ml/mol	McGowan Method
pc	4627.70	kPa	Joback Method
ss	392.40	J/molxK	NIST Webbook
ss	360.20	J/molxK	NIST Webbook
tb	1290.10	K	Joback Method
tc	1782.75	K	Joback Method
tf	462.00 ± 3.00	K	NIST Webbook
tf	424.40	K	Heat capacity and transition behavior of sucrose by standard, fast scanning and temperature-modulated calorimetry
tf	464.05	K	Artificial neural networks as a supporting tool for compatibility study based on thermogravimetric data
tf	461.00 ± 6.00	K	NIST Webbook
tf	458.65	K	Aqueous Solubility Prediction Method
vc	0.784	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1178.26	J/molxK	1782.75	Joback Method
cpg	970.71	J/molxK	1290.10	Joback Method
cpg	998.85	J/molxK	1372.21	Joback Method
cpg	1028.30	J/molxK	1454.32	Joback Method
cpg	1060.01	J/molxK	1536.42	Joback Method
cpg	1094.94	J/molxK	1618.53	Joback Method
cpg	1134.04	J/molxK	1700.64	Joback Method
cps	422.50	J/molxK	297.00	NIST Webbook
cps	408.50	J/molxK	288.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	416.60	J/molxK	293.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	424.30	J/molxK	298.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	429.40	J/molxK	303.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	437.50	J/molxK	308.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides

cps	445.50	J/mol×K	313.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	451.00	J/mol×K	318.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	466.20	J/mol×K	323.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	472.60	J/mol×K	328.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	482.50	J/mol×K	333.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	490.30	J/mol×K	338.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	498.80	J/mol×K	343.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	506.70	J/mol×K	348.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	513.90	J/mol×K	353.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides

cps	522.00	J/molxK	358.15	Temperature dependence of the heat capacities in the solid state of 18 mono-, di-, and poly-saccharides
cps	424.30	J/molxK	298.15	NIST Webbook
cps	430.00	J/molxK	300.00	NIST Webbook
cps	425.50	J/molxK	298.15	NIST Webbook
hfust	46.20	kJ/mol	459.00	NIST Webbook

Sources

Physicochemical study of solute-solute and solute-solvent interactions of Solving an effect of carbohydrates on the surface and viscosity of solutions at different temperatures. Studies of monosaccharides. Carbohydrate Synthesis of Imidazolium Ionic Liquids on the Interactions of Human Hemoglobin with H_2O , K^+ , Ca^{2+} , and Na^+ Ions: A Tool for Compatibility Study Based on Thermodynamic Analysis and the Effect of Organic Solvents on the Behavior of carbohydrate-surfactant system at partial molar volumes and viscosity B-coefficients of arginine in aqueous binary, mutual diffusion coefficients of solutions. Solutions of Sucrose, Estrose, Glucose, and Fructose in the Temperature Range from (298.15 to 328.15) K. Volumetric and ultrasonic properties of ternary (sucrose + water + protic ionic liquid) solutions: acetate on the volumetric behaviour of some mono-, di-, and tri-saccharides in aqueous solution. Solute and solvent range coefficients of glyglycine in sucrose, glucose, and fructose at different temperatures. Sucrose, Palmitate in Solvent. Based on ternary properties of ionic liquids, 1-Ethyl-3-methylimidazolium chloride, 1-butyl-3-methylimidazolium hexafluorophosphate, and Viscometric Studies on Carbohydrate Disodium Tartrate, Sucrose, and Mannitol in Aqueous Solutions of Poly(vinylpyrrolidone) and Poly(vinylalcohol) in Aqueous Solutions of Amino Acids in Aqueous Solutions: Osmotic properties of carbohydrate aqueous solutions: Partial molar volumes, expansibilities and compressibilities of glyglycine in aqueous binary and ternary solutions. Coefficients for Six Sugars at 0.1 MPa and Temperatures from 270.2 to 353.2 K. Different Natural Deep Eutectic Solvents at 278.15 to 328.15 K: activity and vapour pressure for ternary Properties of Sugars, Polyols and Polyols in Water-Ethanol or Ethanol-Water Systems and L-alanine, and L-valine in aqueous sucrose solutions at different temperatures. 1 M aqueous glucose / 1 M aqueous sucrose system at 278.15 to 328.15 K: Study of solute solute and solute solvent interactions of L-methionine in aqueous-sucrose solutions at different temperatures:

<https://www.doi.org/10.1016/j.jct.2013.09.008>

<https://www.doi.org/10.1016/j.jct.2017.08.027>

<https://www.doi.org/10.1021/acs.jced.7b00719>

<https://www.doi.org/10.1021/acs.jced.6b00695>

<https://www.doi.org/10.1016/j.tca.2017.12.015>

<https://www.doi.org/10.1016/j.fluid.2014.05.020>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

<https://www.doi.org/10.1016/j.jct.2004.07.030>

<https://www.doi.org/10.1021/je0602061>

http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt

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<https://www.doi.org/10.1016/j.tca.2014.05.014>

<https://www.doi.org/10.1021/acs.jced.8b01026>

<https://www.doi.org/10.1016/j.fluid.2016.11.001>

<https://www.doi.org/10.1016/j.fluid.2016.05.024>

<https://www.doi.org/10.1021/je400264a>

<https://www.doi.org/10.1016/j.jct.2014.05.015>

<https://www.doi.org/10.1021/je049582g>

<https://www.doi.org/10.1016/j.fluid.2016.02.030>

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<https://www.doi.org/10.1021/je0601816>

<https://www.doi.org/10.1021/acs.jced.6b00552>

<https://www.doi.org/10.1016/j.jct.2017.08.023>

<https://www.doi.org/10.1021/je700190m>

<https://www.doi.org/10.1016/j.jct.2004.12.015>

<https://www.doi.org/10.1016/j.tca.2011.03.024>

<http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa>

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Volumetric and Transport Behavior of
Different Carbohydrates in Aqueous
Solutions
Viscosity of Multicomponent Solutions
of Simple and Complex Sugars in
Water
Webbook:

<https://www.doi.org/10.1021/je2000099>

<https://www.doi.org/10.1021/je700062x>

<http://webbook.nist.gov/cgi/cbook.cgi?ID=C57501&Units=SI>

Legend

chs:	Standard solid enthalpy of combustion
cpg:	Ideal gas heat capacity
cps:	Solid phase heat capacity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfs:	Solid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
ss:	Solid phase molar entropy at standard conditions
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume

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