Cyclohexanol

Other names:	1-Cyclohexanol
	ADRONAL
	ANOL
	Adronol
	CYCLOHEXYL ALCOHOL
	Cicloesanolo
	Cyclohexane, hydroxy-
	Cykloheksanol
	Hexahydrophenol
	Hexalin
	Hydralin
	Hydrophenol
	Hydroxycyclohexane
	NSC 403656
	Naxol
	Phenol, hexahydro-
Inchi:	InChI=1S/C6H12O/c7-6-4-2-1-3-5-6/h6-7H,1-5H2
InchiKey:	HPXRVTGHNJAIIH-UHFFFAOYSA-N
Formula:	C6H12O
SMILES:	OC1CCCCC1
Mol. weight [g/mol]:	100.16
CAS:	108-93-0

Physical Properties

Property code	Value	Unit	Source
af	0.5280		KDB
aigt	573.15	K	KDB
dm	1.70	debye	KDB
fpc	344.26	K	KDB
fpo	340.93	K	KDB
gf	-118.00	kJ/mol	KDB
gyrad	3.4340		KDB
hf	-294.80	kJ/mol	KDB
hfl	-352.00 ± 0.67	kJ/mol	NIST Webbook
hfl	-350.00 ± 2.00	kJ/mol	NIST Webbook
hfl	-359.20	kJ/mol	NIST Webbook
hfl	-349.20 ± 0.20	kJ/mol	NIST Webbook

hfl	-347.40 ± 2.20	kJ/mol	NIST Webbook
hfus	7.22	kJ/mol	Joback Method
hvap	46.06	kJ/mol	Joback Method
ie	10.00 ± 0.20	eV	NIST Webbook
ie	9.75	eV	NIST Webbook
ie	10.00	eV	NIST Webbook
log10ws	-0.44		Aqueous Solubility Prediction Method
log10ws	-0.44		Estimated Solubility Method
logp	1.311		Crippen Method
mcvol	90.410	ml/mol	McGowan Method
nfpaf	%!d(float64=2)		KDB
nfpah	%!d(float64=1)		KDB
pc	4260.00 ± 50.00	kPa	NIST Webbook
pc	4080.00 ± 150.00	kPa	NIST Webbook
pc	4260.00	kPa	KDB
pc	4401.00 ± 25.00	kPa	NIST Webbook
pc	4260.00 ± 50.00	kPa	NIST Webbook
pc	3749.02 ± 151.99	kPa	NIST Webbook
rhoc	300.48 ± 10.02	kg/m3	NIST Webbook
rinpol	880.00		NIST Webbook
rinpol	851.10		NIST Webbook
rinpol	921.10		NIST Webbook
rinpol	852.90		NIST Webbook
rinpol	855.00		NIST Webbook
rinpol	885.00		NIST Webbook
rinpol	891.00		NIST Webbook
rinpol	880.00		NIST Webbook
rinpol	900.00		NIST Webbook
rinpol	874.00		NIST Webbook
rinpol	874.00		NIST Webbook
rinpol	879.00		NIST Webbook
rinpol	883.00		NIST Webbook
rinpol	878.00		NIST Webbook
rinpol	863.00		NIST Webbook
rinpol	880.00		NIST Webbook
rinpol	898.00		NIST Webbook
rinpol	869.00		NIST Webbook
rinpol	872.00		NIST Webbook
rinpol	886.00		NIST Webbook
rinpol	856.00		NIST Webbook
rinpol	885.00		NIST Webbook
rinpol	881.00		NIST Webbook
rinpol	881.00		NIST Webbook

rinpol	872.00	NIST Webbook
rinpol	866.00	NIST Webbook
rinpol	869.00	NIST Webbook
rinpol	870.00	NIST Webbook
rinpol	875.00	NIST Webbook
rinpol	881.00	NIST Webbook
rinpol	866.00	NIST Webbook
rinpol	861.00	NIST Webbook
rinpol	876.00	NIST Webbook
rinpol	891.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	870.00	NIST Webbook
rinpol	867.00	NIST Webbook
rinpol	849.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	928.00	NIST Webbook
rinpol	876.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	881.00	NIST Webbook
rinpol	874.00	NIST Webbook
rinpol	874.00	NIST Webbook
rinpol	899.00	NIST Webbook
rinpol	880.00	NIST Webbook
rinpol	913.00	NIST Webbook
rinpol	889.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1410.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1400.00	NIST Webbook
ripol	1396.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1394.00	NIST Webbook
ripol	1377.00	NIST Webbook
ripol	1386.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1430.00	NIST Webbook
ripol	1403.00	NIST Webbook
ripol	1386.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1375.00	NIST Webbook
ripol	1376.00	NIST Webbook

ripol	1414.00		NIST Webbook
ripol	1407.00		NIST Webbook
ripol	1408.00		NIST Webbook
ripol	1378.00		NIST Webbook
ripol	1410.00		NIST Webbook
ripol	1403.00		NIST Webbook
ripol	1403.00		NIST Webbook
ripol	1403.00		NIST Webbook
ripol	1393.00		NIST Webbook
ripol	1375.00		NIST Webbook
ripol	1392.00		NIST Webbook
ripol	1375.00		NIST Webbook
ripol	1392.00		NIST Webbook
sg	353.83	J/mol×K	NIST Webbook
sl	203.87	J/mol×K	NIST Webbook
sl	199.60	J/mol×K	NIST Webbook
tb	433.73	K	A New Test System for Distillation Efficiency Experiments at Elevated Liquid Viscosities: Vapor-Liquid Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol
tb	433.99	K	KDB
tb	433.89	K	Isobaric Vapor-Liquid Equilibria of the Ternary System Methylbutyl Ketone + Nonane + Cyclohexanol
tc	650.10	K	KDB
tf	296.62	K	Aqueous Solubility Prediction Method
tf	298.80	K	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols
tf	297.20	K	Vapour liquid equilibria, azeotropic data, excess enthalpies, activity coefficients at infinite dilution and solid liquid equilibria for binary alcohol ketone systems
tf	298.61	K	KDB
tt	299.09 ± 0.02	K	NIST Webbook
tt	297.00 ± 0.20	K	NIST Webbook
vc	0.324	m3/kmol	Joback Method
zra	0.24		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	201.24	J/mol×K	480.91	Joback Method
cpg	246.66	J/mol×K	610.89	Joback Method
cpg	213.50	J/mol×K	513.40	Joback Method
cpg	225.15	J/mol×K	545.90	Joback Method
cpg	236.19	J/mol×K	578.39	Joback Method
cpg	256.55	J/mol×K	643.39	Joback Method
cpg	188.35	J/mol×K	448.41	Joback Method
cpl	220.10	J/mol×K	298.00	NIST Webbook
cpl	212.00	J/mol×K	297.95	NIST Webbook
cpl	213.59	J/mol×K	300.00	NIST Webbook
cpl	209.99	J/mol×K	298.15	NIST Webbook
cpl	235.50	J/mol×K	320.00	Heat capacities of selected cycloalcohols
cpl	214.06	J/mol×K	298.15	NIST Webbook
cpl	274.70	J/mol×K	357.94	Heat capacities of selected cycloalcohols
cpl	202.50	J/mol×K	305.10	NIST Webbook
cpl	209.03	J/mol×K	298.15	NIST Webbook
cpl	218.30	J/mol×K	305.05	Heat capacities of selected cycloalcohols
cpl	218.60	J/mol×K	305.18	Heat capacities of selected cycloalcohols
cpl	223.80	J/mol×K	310.00	Heat capacities of selected cycloalcohols
cpl	224.00	J/mol×K	310.00	Heat capacities of selected cycloalcohols
cpl	229.60	J/mol×K	315.00	Heat capacities of selected cycloalcohols
cpl	229.70	J/mol×K	315.00	Heat capacities of selected cycloalcohols
cpl	174.90	J/mol×K	290.00	NIST Webbook
cpl	235.30	J/mol×K	320.00	Heat capacities of selected cycloalcohols

cpl	272.50	J/mol×K	355.00	Heat capacities of selected cycloalcohols
cpl	241.10	J/mol×K	325.00	Heat capacities of selected cycloalcohols
cpl	246.70	J/mol×K	330.00	Heat capacities of selected cycloalcohols
cpl	246.80	J/mol×K	330.00	Heat capacities of selected cycloalcohols
cpl	252.00	J/mol×K	335.00	Heat capacities of selected cycloalcohols
cpl	252.20	J/mol×K	335.00	Heat capacities of selected cycloalcohols
cpl	257.40	J/mol×K	340.00	Heat capacities of selected cycloalcohols
cpl	257.60	J/mol×K	340.00	Heat capacities of selected cycloalcohols
cpl	262.80	J/mol×K	345.00	Heat capacities of selected cycloalcohols
cpl	263.00	J/mol×K	345.00	Heat capacities of selected cycloalcohols
cpl	268.00	J/mol×K	350.00	Heat capacities of selected cycloalcohols
cpl	268.00	J/mol×K	350.00	Heat capacities of selected cycloalcohols
cpl	272.30	J/mol×K	355.00	Heat capacities of selected cycloalcohols
cpl	241.30	J/mol×K	325.00	Heat capacities of selected cycloalcohols
cpl	274.70	J/mol×K	357.82	Heat capacities of selected cycloalcohols
dvisc	0.0008063	Pa×s	374.13	Joback Method
dvisc	0.0004183	Pa×s	411.27	Joback Method
dvisc	0.0002419	Pa×s	448.41	Joback Method
dvisc	0.0966465	Pa×s	225.58	Joback Method
dvisc	0.0175844	Pa×s	262.72	Joback Method
dvisc	0.0017962	Pa×s	337.00	Joback Method
dvisc	0.0048796	Pa×s	299.86	Joback Method
hfust	1.70	kJ/mol	297.00	NIST Webbook
hfust	8.21	kJ/mol	263.50	NIST Webbook

hfust	1.73	kJ/mol	298.20	NIST Webbook
hfust	1.70	kJ/mol	297.00	NIST Webbook
hsubt	60.70	kJ/mol	285.00	NIST Webbook
hvapt	49.80	kJ/mol	410.00	NIST Webbook
hvapt	54.80	kJ/mol	364.50	NIST Webbook
hvapt	52.60	kJ/mol	400.00	NIST Webbook
hvapt	60.40	kJ/mol	309.00	NIST Webbook
hvapt	45.44	kJ/mol	431.70	NIST Webbook
hvapt	49.30	kJ/mol	418.00	NIST Webbook
hvapt	62.70	kJ/mol	367.00	NIST Webbook
hvapt	59.90	kJ/mol	376.00	NIST Webbook
hvapt	55.00	kJ/mol	403.00	NIST Webbook
hvapt	61.20 ± 0.60	kJ/mol	308.00	NIST Webbook
hvapt	58.40	kJ/mol	338.00	NIST Webbook
hvapt	60.10	kJ/mol	377.50	NIST Webbook
hvapt	45.48	kJ/mol	433.70	KDB
pvap	1.00	kPa	330.95	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	3.95	kPa	353.80	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	7.90	kPa	367.25	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	2.00	kPa	341.70	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	31.75	kPa	399.50	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol

pvap	0.50	kPa	322.25	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	63.75	kPa	419.05	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	100.05	kPa	433.30	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
pvap	99.00	kPa	433.38	Vapor-Liquid Equilibria of Cyclohexanone + 2-Cyclohexen-1-one and Cyclohexanol + 2-Cyclohexen-1-one, Validated in a Packed Column Distillation.
pvap	15.85	kPa	382.40	VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol
rfi	1.46480		293.20	Phase Equilibria for the Ternary Liquid Systems of (Water + Tetrahydrofurfuryl Alcohol + Cyclic Solvent) at 298.2 K
rfi	1.46320		303.15	Experimental study on the calorimetric data of cyclohexanol with alkanols (C1-C4) and correlation with Wilson, NRTL and UNIQUAC models at 300 K

rfi	1.46390		293.15	Phase equilibria of water + 1-propanol + solvent (n-amyl acetate, cyclohexanol, and cyclohexyl acetate) at T = 298.2K
rhol	942.00	kg/m3	303.00	KDB
rhol	941.42	kg/m3	303.20	Optimization and modeling of extraction equilibria of the type 2 ternary systems containing (water + isovaleric acid + solvent)
rhol	941.40	kg/m3	293.10	Vapor-Liquid Equilibrium for Binary Systems of Cyclohexane + Cyclohexanone and + Cyclohexanol at Temperatures from (414.0 to 433.7) K
rhol	961.50	kg/m3	298.15	Determination and Correlation of Liquid-Liquid Equilibria Data for Water + Cyclohexanol + (Mesityl Oxide, Toluene, and p-Xylene) Ternary Systems at Different Temperatures
sfust	5.72	J/mol×K	297.00	NIST Webbook
sfust	31.14	J/mol×K	263.50	NIST Webbook
speedsl	1443.20	m/s	303.15	Densities, Viscosities, Sound Speed, and IR Studies of N-methyl-2-pyrrolidone with Cyclohexylamine, Cyclohexanol, and Cyclohexene at different Temperatures.

speedsl	1419.60	m/s	308.15	Densities, Viscosities, Sound Speed, and IR Studies of N-methyl-2-pyrrolidone with Cyclohexylamine, Cyclohexanol, and Cyclohexene at different Temperatures.
speedsl	1370.80	m/s	318.15	Densities, Viscosities, Sound Speed, and IR Studies of N-methyl-2-pyrrolidone with Cyclohexylamine, Cyclohexanol, and Cyclohexene at different Temperatures.
speedsl	1395.20	m/s	313.15	Densities, Viscosities, Sound Speed, and IR Studies of N-methyl-2-pyrrolidone with Cyclohexylamine, Cyclohexanol, and Cyclohexene at different Temperatures.
srf	0.03	N/m	323.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures
srf	0.03	N/m	318.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures
srf	0.03	N/m	308.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures
srf	0.03	N/m	298.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures

srf	0.03	N/m	293.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures
srf	0.03	N/m	313.20	KDB
srf	0.03	N/m	313.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures
srf	0.03	N/m	303.15	Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	341.68	K	2.00	A New Test System for Distillation Efficiency Experiments at Elevated Liquid Viscosities: Vapor-Liquid Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol
tbp	372.00	K	10.00	A New Test System for Distillation Efficiency Experiments at Elevated Liquid Viscosities: Vapor-Liquid Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol

tbp	387.67	K	20.00	A New Test System for Distillation Efficiency Experiments at Elevated Liquid Viscosities: Vapor-Liquid Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol
tbp	417.16	K	60.00	A New Test System for Distillation Efficiency Experiments at Elevated Liquid Viscosities: Vapor-Liquid Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.45470e+01
Coeff. B	-3.25877e+03
Coeff. C	-1.04780e+02
Temperature range (K), min.	333.32
Temperature range (K), max.	457.63

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	1.29104e+02
Coeff. B	-1.18780e+04
Coeff. C	-1.61598e+01
Coeff. D	5.55366e-06
Temperature range (K), min.	296.60
Temperature range (K), max.	625.15

Sources

Phase Equilibria for the Ternary Liquid Systems of (Water + Tetrahydrofurfuryl Alcohol + Cyclic Solvent) at 298.2 K:	https://www.doi.org/10.1021/je049605r
Estimated Solubility Method:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Surface Tension of Dilute Solutions of Alkanes in Cyclohexanol at Different Temperatures:	https://www.doi.org/10.1021/je8004112
Determination and Correlation for the Solubilities of Succinic Acid in Cyclohexanol + Cyclohexanone + Cyclohexane Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.7b00956
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 9. Activity Coefficients at Infinite Dilution of measurements and modeling for the Alcohols in + cyclohexanol binary	https://www.doi.org/10.1021/je050440b
Vapor-liquid equilibrium	https://www.doi.org/10.1016/j.fluid.2010.06.013
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Thermochemistry of Branched Ethers:	https://www.doi.org/10.1021/je034172y
Experimental Study of Chemical Equilibrium in the Reacting System of tert-Amyl Ether + Cyclohexanol + Water:	https://www.doi.org/10.1021/acs.jced.7b00683
Gas solubility of carbon dioxide and oxygen in cyclohexanol by experiment	https://www.doi.org/10.1016/j.jct.2012.01.017
Measurement and prediction of liquid-liquid equilibria in ternary systems containing water, an organic component, and cyclohexanol:	https://www.doi.org/10.1016/j.fluid.2016.12.001
Phase Behavior of the Reactant and Products of Cyclohexane Oxidation in Compressed CO2:	https://www.doi.org/10.1021/je800109s
Liquid-liquid equilibrium for the system water + 1,4-dioxane + cyclohexanol over the temperature range of 313.2 - 343.2 K:	https://www.doi.org/10.1016/j.fluid.2012.03.010
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Distribution of cyclohexanol and cyclohexanone between water and cyclohexane:	https://www.doi.org/10.1016/j.fluid.2015.02.029
Measurements and Correlation of Liquid-Liquid Equilibria for the Ternary System Cyclohexanol + Water + Cyclohexyl Formate at 298.2-338.2 K:	https://www.doi.org/10.1021/acs.jced.8b01104
KDB Pure (Korea Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=899
Isobaric Vapor-Liquid Equilibria of the Ternary System Methylbutyl Ketone + Nonane + Cyclohexanol:	https://www.doi.org/10.1021/je040002p
Solubilities of adipic acid in binary cyclohexanone + cyclohexanol, cyclohexane + cyclohexanol, and cyclohexane + cyclohexanone solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2016.01.006
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 4. LLE of Binary Mixtures of [C2MIM][NTf2] with	https://www.doi.org/10.1021/je049564z
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C108930&Units=SI
Propan-1-ol, Butan-1-ol, and Pentan-1-ol and [C4MIM][NTf2] with	https://www.doi.org/10.1016/j.tca.2014.10.002
Heat capacities of selected cycloalcohols:	http://link.springer.com/article/10.1007/BF02311772
Cyclopentanol and 1,2-Hexanediol Including Studies of the Influence of Small Amounts of Water:	https://www.doi.org/10.1016/j.fluid.2013.10.003
Measurement on the solubility of adipic acid in various solvents at high temperature and its thermodynamics parameters:	https://www.doi.org/10.1021/je900549r
(Liquid + Liquid) Equilibria of (Cyclohexane + Dimethyl Sulfoxide + Cyclohexanol) at T/K 303.2 K:	https://www.doi.org/10.1016/j.jct.2015.07.035
Optimization and modeling of extraction equilibria of the type 2 systems containing (water + cyclohexanol + alkanols (C1-C4) and correlation with Wilson,	https://www.doi.org/10.1016/j.jct.2012.02.037
Experimental study on the calorimetric data of cyclohexanol with alkanols	https://www.doi.org/10.1021/acs.jced.8b00929
A New Test System for Distillation Efficiency Experiments at Elevated	https://www.doi.org/10.1016/j.jct.2012.06.019
SRTL and UNIQUAC models at 300 K:	https://www.doi.org/10.1016/j.fluid.2008.02.021
Liquid Viscosities of Ternary Liquid Mixtures: (water + pyruvic acid + high boiling	https://www.doi.org/10.1016/j.tca.2014.06.026
Equilibrium and Liquid Viscosity Data for Cyclopentanol + Cyclohexanol +	https://www.doi.org/10.1021/acs.jced.6b00325
Vapour-liquid equilibria in azeotropic mixtures:	https://en.wikipedia.org/wiki/Joback_method
Optimization of activity coefficients at infinite dilution and solid	https://www.cheric.org/files/research/kdb/mol/mol899.mol
Densities, Viscosities, Sound Speed, and IR Studies of N-methyl-2-pyrrolidone (1) + Cyclohexylamine,	https://www.doi.org/10.1021/je900886s
Determination and Correlation for the Solubility of Glutaric Acid in Cyclohexanol, Cyclohexanone, and Cyclohexene at	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Joback Method:	https://www.doi.org/10.1021/acs.jced.8b00273
KDB:	https://www.doi.org/10.1021/acs.jced.5b00880
Liquid-Liquid Equilibrium for the System Water + Cyclohexene + Cyclohexanol over the Temperature Range of (303.2 to 403.2) K:	https://www.doi.org/10.1021/acs.jced.7b00674
Aqueous Solubility Prediction Method:	
Vapor-Liquid Phase Equilibrium of a Cyclohexene + Water + Cyclohexanol + Cyclohexanone Quaternary System at 500 kPa:	
Solubilities of Adipic Acid in Isophorone + Cyclohexanol + Cyclohexanone Mixtures and Correlation of Liquid-Liquid Equilibria Data for Water + Cyclohexanol + (Mesityl Oxide, Toluene, and p-Xylene) Ternary Systems at Different Temperatures:	

Measurement and Prediction of Vapor-Liquid Equilibria in Ternary Systems Containing an Organic Component, Cyclohexanone, and Cyclohexanol: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	https://www.doi.org/10.1021/acs.jced.7b00721
Vapor Pressure and Its Temperature Dependence of 26 Organic Compounds at 298.15 K	https://www.doi.org/10.1021/acs.jced.6b00576
Phase equilibria of water + 1-propanol + solvent in amyl acetate	https://www.doi.org/10.1016/j.fluid.2007.01.021
Vapor-Liquid Equilibria of Cyclohexanone + 2-Cyclohexen-1-one and Cyclohexanol + 2-Cyclohexen-1-one Containing Ionic Liquids: Activity Coefficients at Infinite Dilution and Selectivity at T = 298.2 K	https://www.doi.org/10.1021/je501135p
Thermodynamic Properties of Mixtures Containing Ionic Liquids. 6. Activity Coefficients at Infinite Dilution of Cyclohexane, Hex-1-ene, Benzene, Cyclohexanone, Ethanol, and Ethyl Acetate in the Ionic Liquid 4-Methyl-N-butylpyridinium Tetrafluoroborate from (298.15 to 328.15) K	https://www.doi.org/10.1021/je0503554
Experimental and Theoretical Study of the Phase Equilibria, Excess Properties, and Heat Capacities of Binary Mixtures of 1,4-Butanediol with Cyclohexanone at 298.15 K	https://www.doi.org/10.1021/je8008792
Vapor−Liquid Equilibrium for Binary Systems of Cyclohexane + Cyclohexanone and Cyclohexanol + Cyclohexanone at 13.33 and 101.33 kPa	https://www.doi.org/10.1021/je100028s
Measurement and Correlation of Vapor-Liquid Equilibrium for Binary Systems Containing Water, an Organic Component, and Cyclohexanol: Cyclohexanone + Cyclohexanol at 13.33 and 101.33 kPa	https://www.doi.org/10.1021/acs.jced.7b00100
Liquid Liquid Equilibria for the Ternary System Water + Cyclohexanol + Cyclohexanone:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=899
VLE and LLE Data for the System Cyclohexane + Cyclohexene + Water + Cyclohexanol:	https://www.doi.org/10.1021/je049902w
Measurements and Correlation of Liquid Liquid Equilibria for the Ternary System Water + Cyclohexanol + Cyclohexanone:	https://www.doi.org/10.1021/je500099u

Legend

af:	Acentric Factor
aigt:	Autoignition Temperature
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dm:	Dipole Moment
dvisc:	Dynamic viscosity
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hsubt:	Enthalpy of sublimation at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
nfpah:	NFPA Health Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density

rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zra:	Rackett Parameter